Outline: Super Heavy Elements

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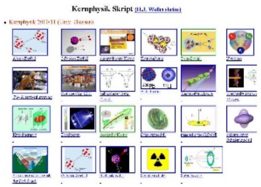
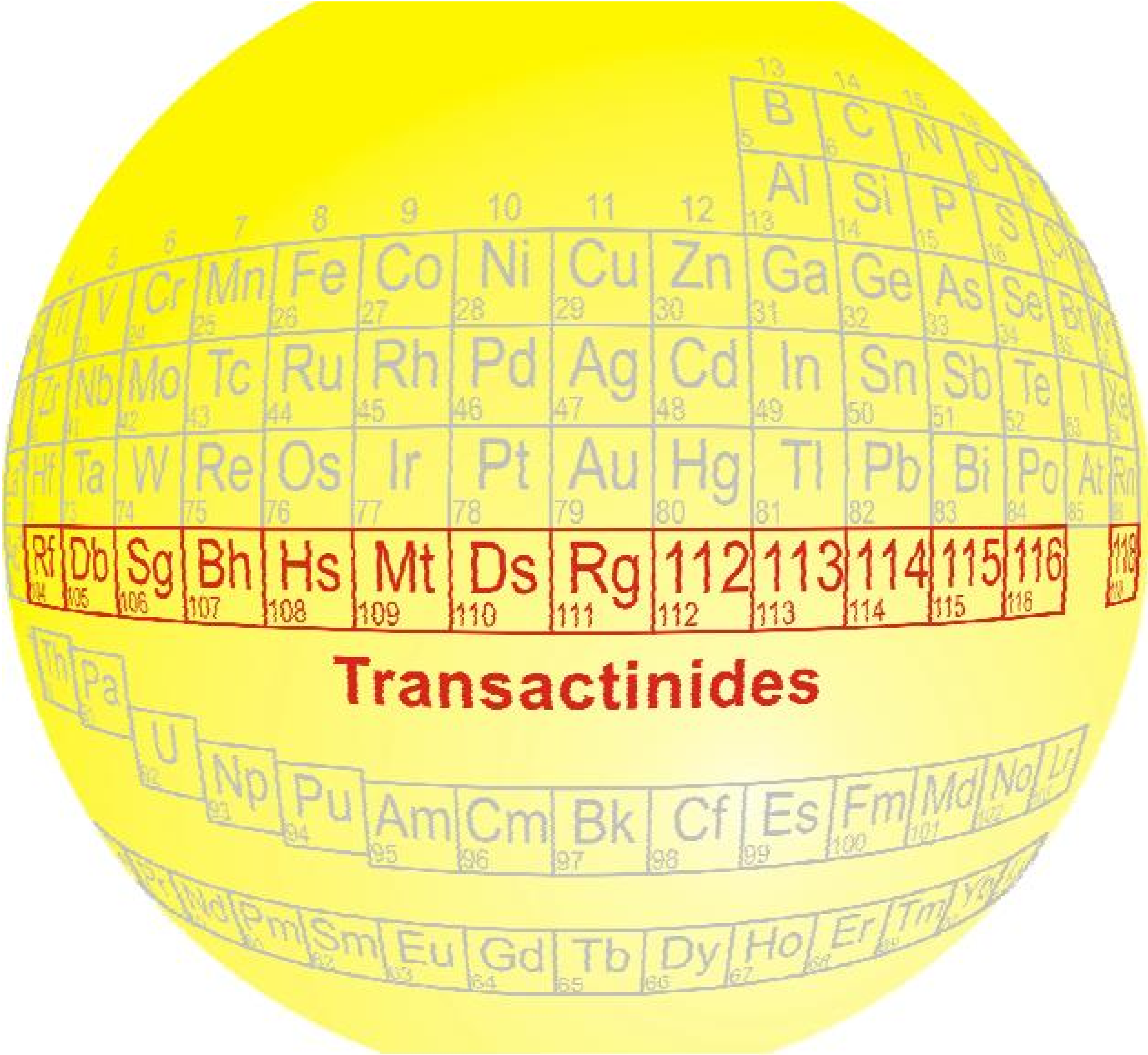
web-page: https://web-docs.gsi.de/~wolle/ and click on



1. history
2. cold and hot fusion
3. nuclear structure of SHE
4. chemistry of SHE

Super Heavy Elements



The Chart of Nuclides the “playground” for Nuclear Physics

chart of nuclides:

-representation of isotopes in the Z-N plane

-isotope: atom (nucleus) of an element with different number of neutrons

Pb (lead) and Bi (bismuth)

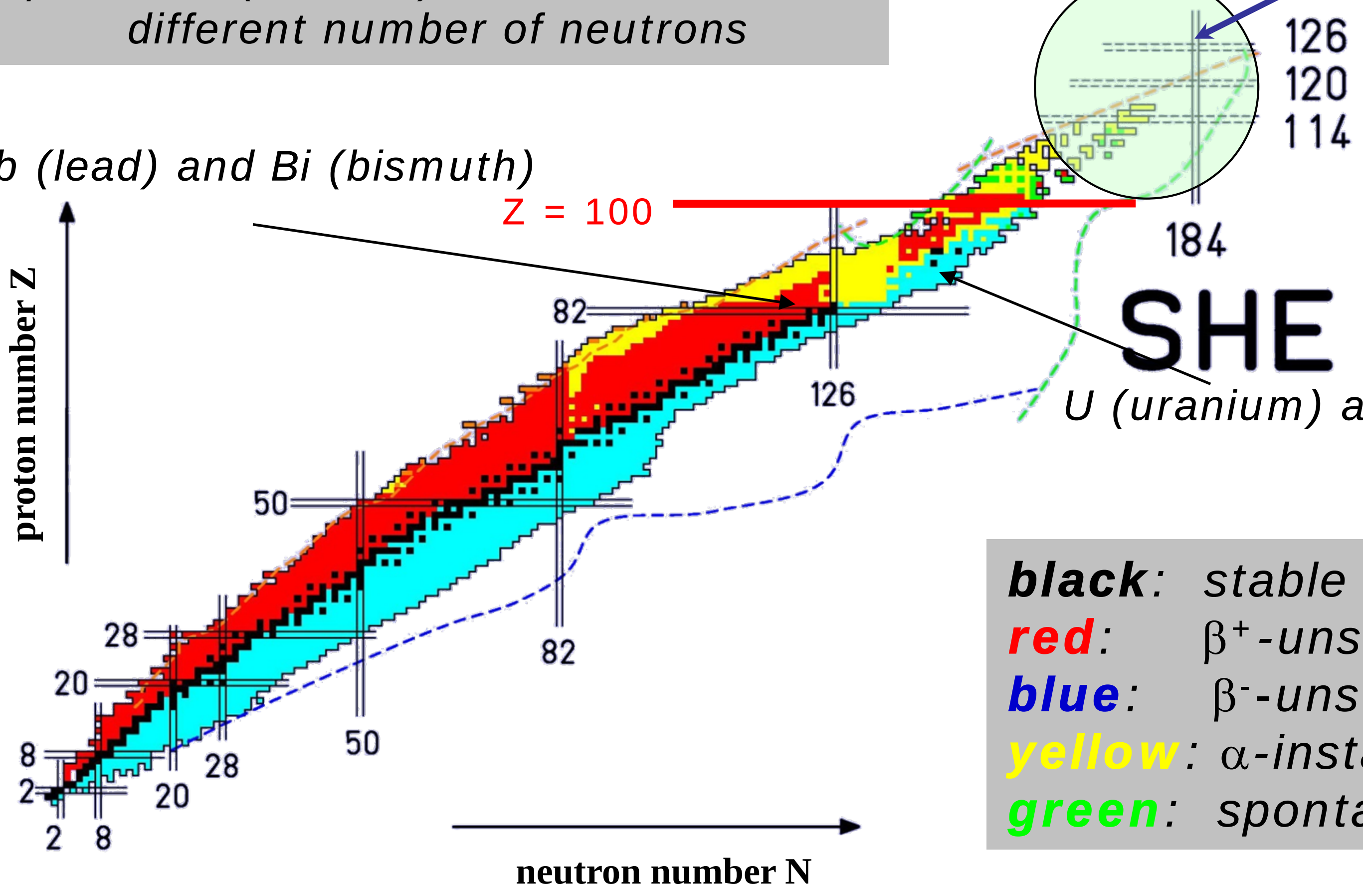
Z = 100

?

island of stability

stabilisation via shell effects

U (uranium) and Th (thorium)



***black**: stable isotope*

***red**: β^+-unstable isotope*

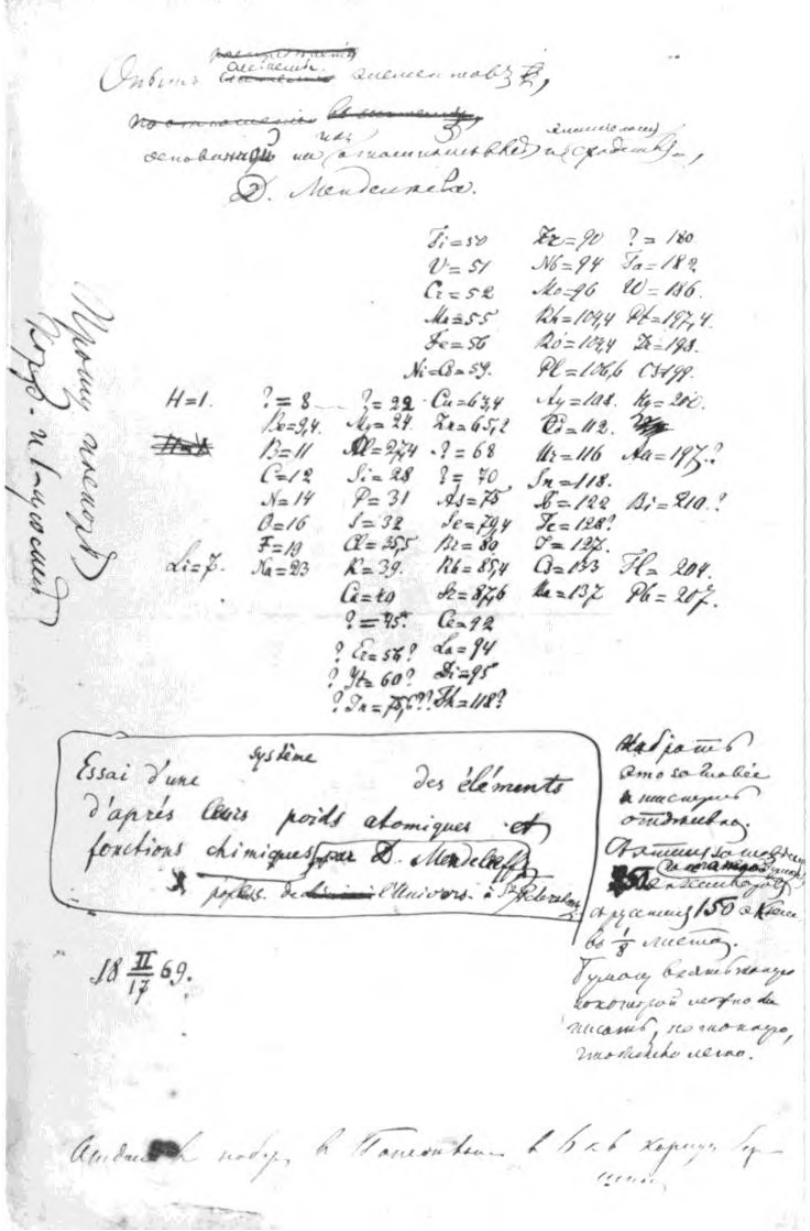
***blue**: β^--unstable isotope*

***yellow**: α-instable isotope*

***green**: spontaneous fission*

Periodic table of the elements

Dmitri Mendeleev (1869)



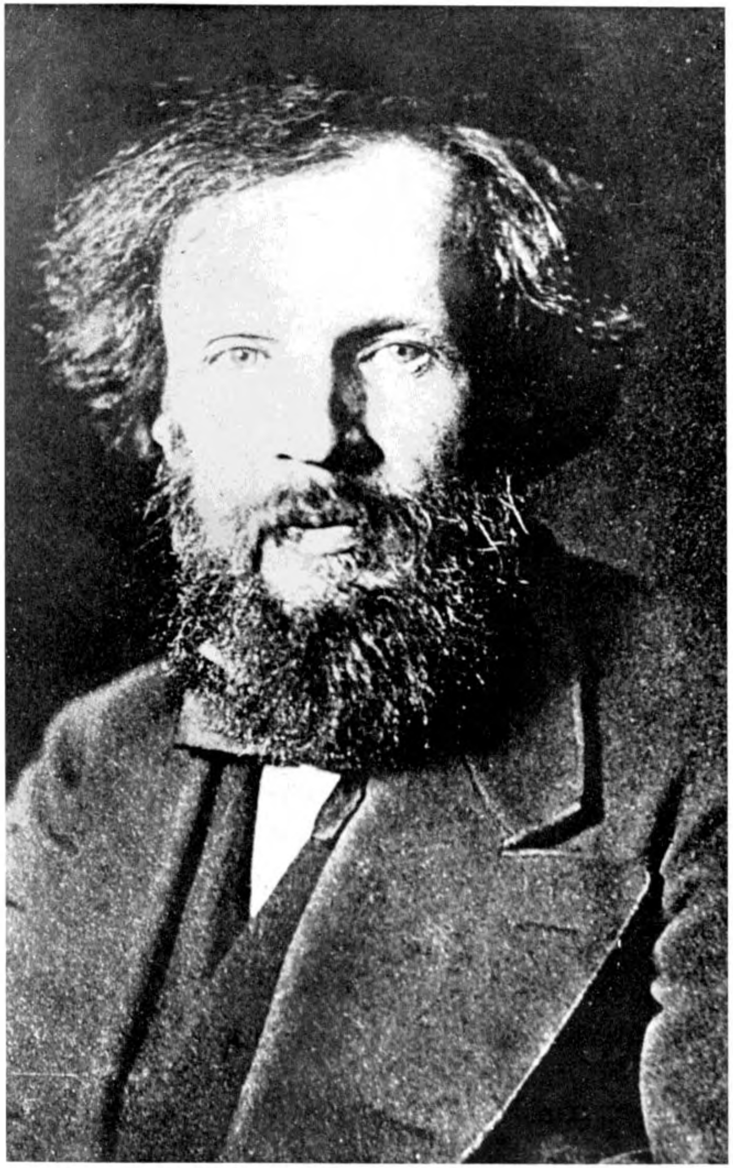
Д. Менделѣева.

			Ti=50	Zr=90	?=180
			V=51	Nb=94	Ta=182
			Cr=52	Mo=96	W=186
			Mn=55	Rh=104,4	Pt=197,4
			Fe=56	Ro=104,4	Ir=198
			Ni=Co=59	Pl=106,6	Os=199
H=1	?=8	?=22	Cu=63,4	Ag=108	Hg=200
	Be=9,4	Mg=24	Zn=65,2	Cd=112	
	B=11	Al=27,4	?=68	Ur=116	Au=197?
	C=12	Si=28	?=70	Sn=118	
	N=14	P=31	As=75	Sb=122	Bi=210?
	O=16	S=32	Se=79,4	Te=128?	
	F=19	Cl=35,5	Br=80	J=127	
Li=7	Na=23	K=39	Rb=85,4	Cs=133	Tl=204
		Ca=40	Sr=87,6	Ba=137	Pb=207
		?=45	Ce=92		
		?Er=56?	La=94		
		?Yt=60?	Di=95		
		?In=75,6?	Th=118?		

Essai d'une système des éléments d'après leurs poids atomiques et fonctions chimiques par D. Mendeleeff

prof. de l'Univers. à St Pétersbourg

18 $\frac{II}{17}$ 69.

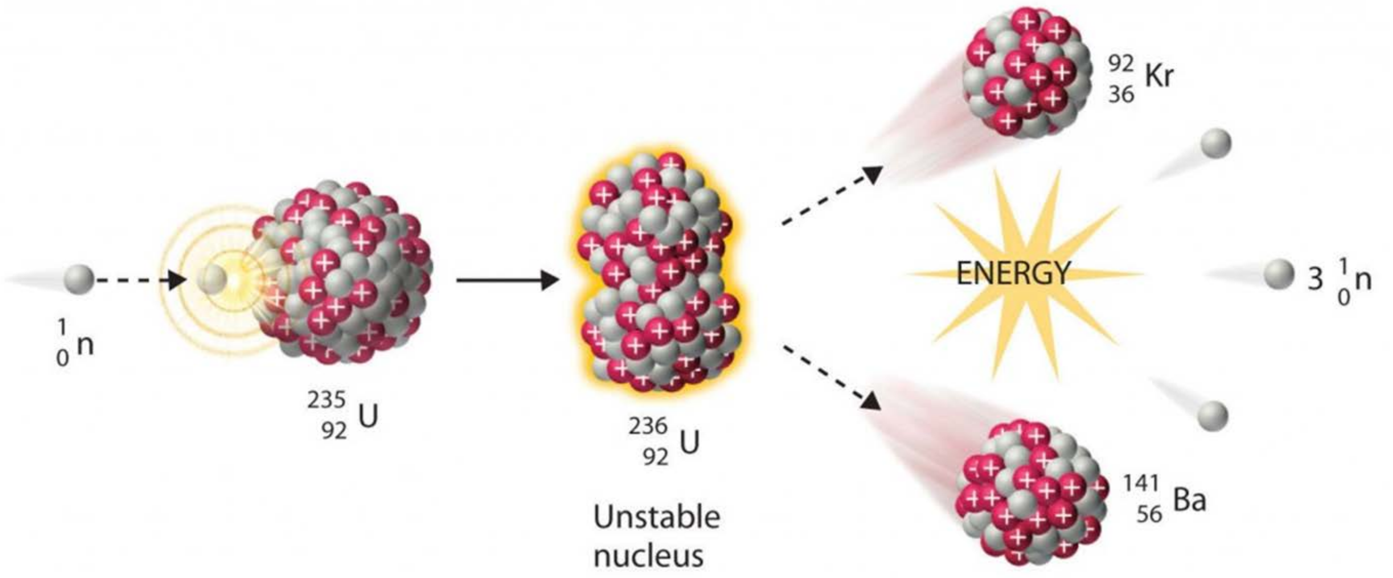


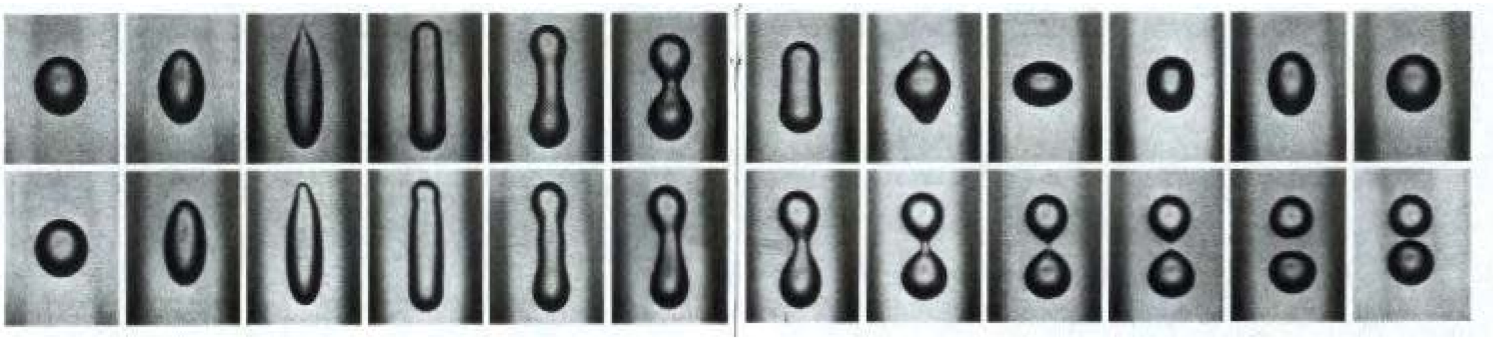
Search for transuranium elements (1934)

$${}^{238}_{92}U + n \rightarrow {}^{239}_{92}U^{*} \rightarrow {}^{239}_{93}Np + e^{-} + \bar{\nu}$$

Otto Hahn and Lise Meitner

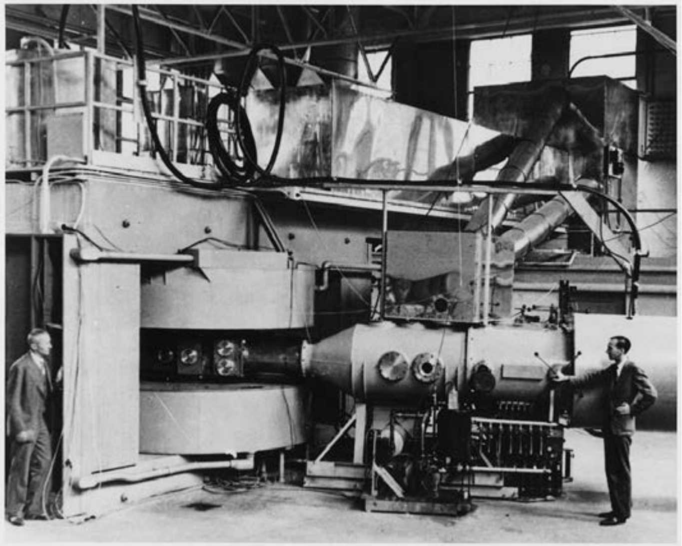
Discovery of nuclear fission (1938)





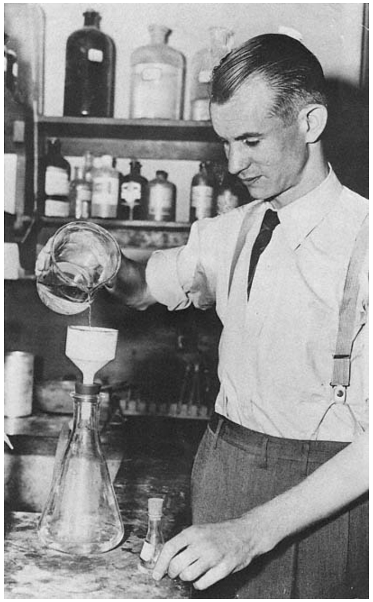
Liquid drop model: G. Gamow; Proc. Roy. Soc. London A123, 373 (1929)
N. Bohr & J.A. Wheeler; Phys. Rev. 56, 426 (1939)

Transuranium elements (1940)



60-inch cyclotron (Berkeley, 1939)

H_2^+, ^{2}H & ^{4}He beams (Q/A = ½)



Edwin M. McMillan

$$^{238}_{92}U + ^{1}_{0}n \rightarrow ^{239}_{92}U^{*} \xrightarrow[23\ min]{\beta^-} ^{239}_{93}Np \xrightarrow[2.355\ days]{\beta^-} ^{239}_{94}Pu$$

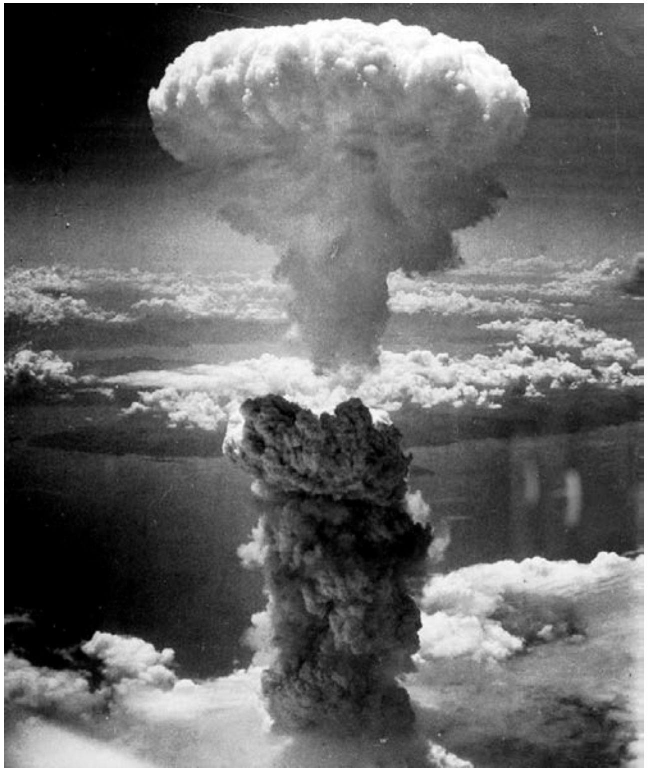
Philip H. Abelson

Neptunium sphere
with Ni clad in U shells

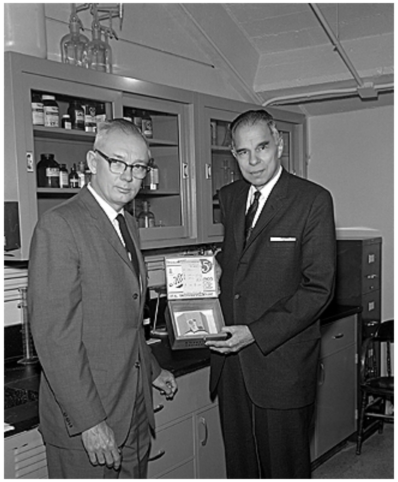
Making new elements by simple reactions

$$^{238}_{92}U + ^{2}_{1}H \rightarrow ^{238}_{93}Np + 2\,^{1}_{0}n$$

$$^{238}_{93}Np \xrightarrow[t_{1/2}=2.12\ days]{\beta^-} {}^{238}_{94}Pu(t_{1/2} = 87.7\ years)$$



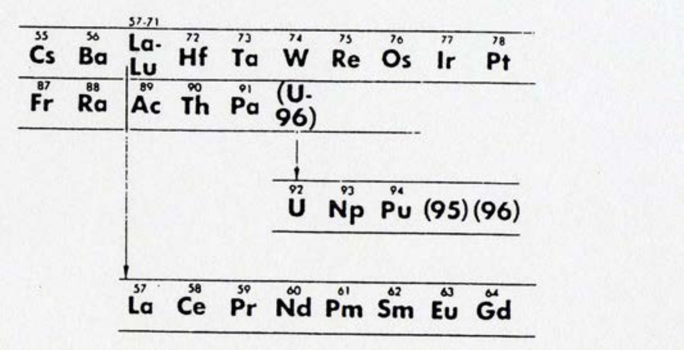
Joseph W. Kennedy 1940



Arthur C. Wahl and
Glenn T. Seaborg (1966)

Making new elements by simple reactions – the role of chemistry

- The discovery of elements 95 (Am) and 96 (Cm)



Glenn T. Seaborg

$$^{239}_{94}Pu + ^{1}_{0}n \rightarrow ^{240}_{94}Pu + \gamma$$

$$^{240}_{94}Pu + ^{1}_{0}n \rightarrow ^{241}_{94}Pu + \gamma$$

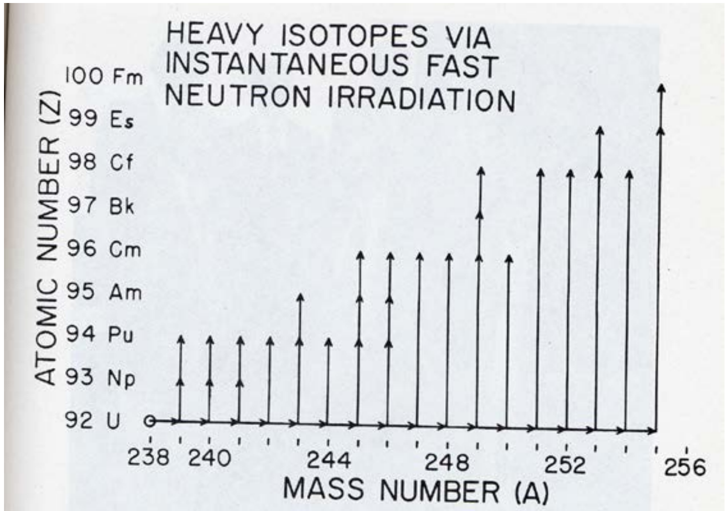
$$^{241}_{94}Pu \xrightarrow[t_{1/2}=14.4\ years]{\beta^-} {}^{241}_{95}Am\left(t_{1/2} = 432.7\ years\right)$$

$$^{241}_{95}Am + ^{1}_{0}n \rightarrow ^{242}_{95}Am + \gamma$$

$$^{242}_{95}Am \xrightarrow[t_{1/2}=16.0\ h]{\beta^-} {}^{242}_{96}Cm$$

Making new elements with nuclear weapons

- The discovery of elements 99 (Md) and 100 (Fm)



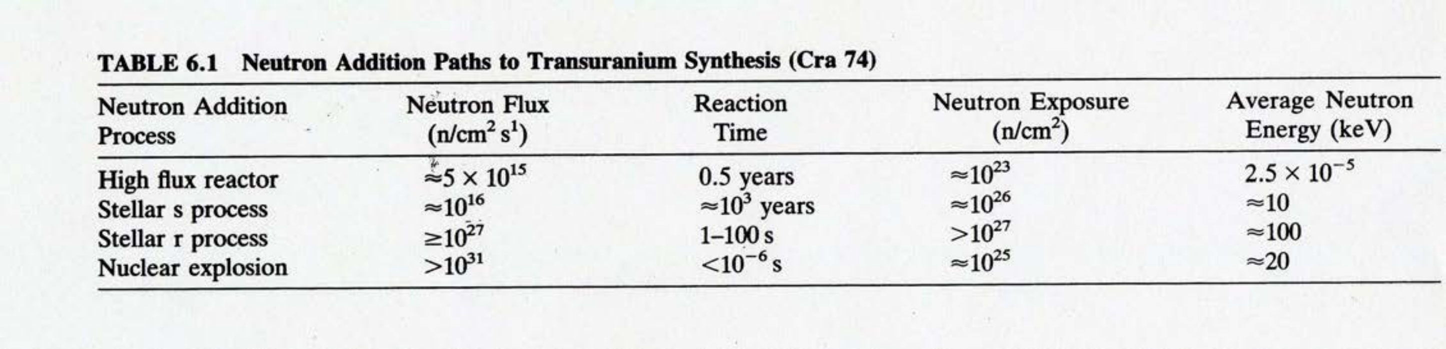


TABLE 6.1 Neutron Addition Paths to Transuranium Synthesis (Cra 74)

Neutron Addition Process	Neutron Flux ($n/cm^2\,s^1$)	Reaction Time	Neutron Exposure (n/cm^2)	Average Neutron Energy (keV)
High flux reactor	$\approx 5 \times 10^{15}$	0.5 years	$\approx 10^{23}$	2.5×10^{-5}
Stellar s process	$\approx 10^{16}$	$\approx 10^3$ years	$\approx 10^{26}$	≈ 10
Stellar r process	$\geq 10^{27}$	1–100 s	$> 10^{27}$	≈ 100
Nuclear explosion	$> 10^{31}$	$< 10^{-6}$ s	$\approx 10^{25}$	≈ 20

MIKE

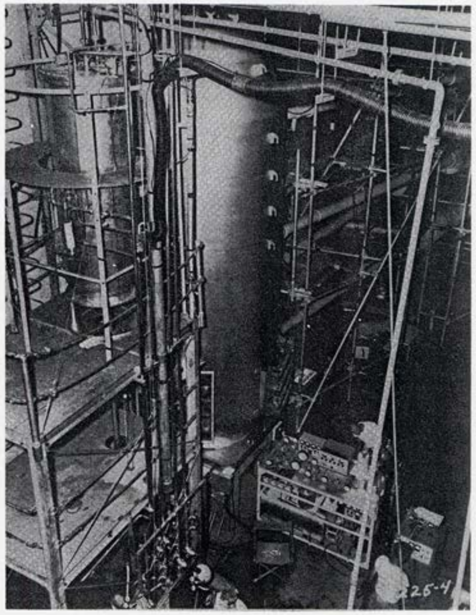
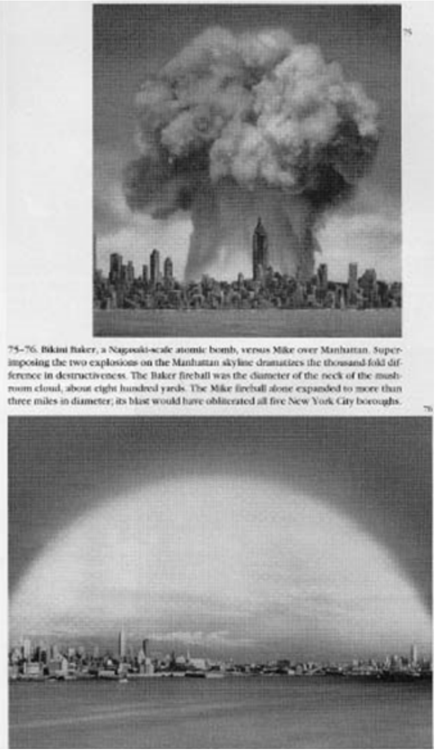


Figure 2.9 Closeup view of Mike device with its associated cryogenic equipment.

MIKE

73-74. Mike yielded 10.4 megatons, one thousand times the yield of the Hiroshima bomb, vaporizing the island of Elugelab. The crater the explosion left behind—two hundred feet deep and more than one mile across—is starkly visible in these before-and-after aerial photographs.

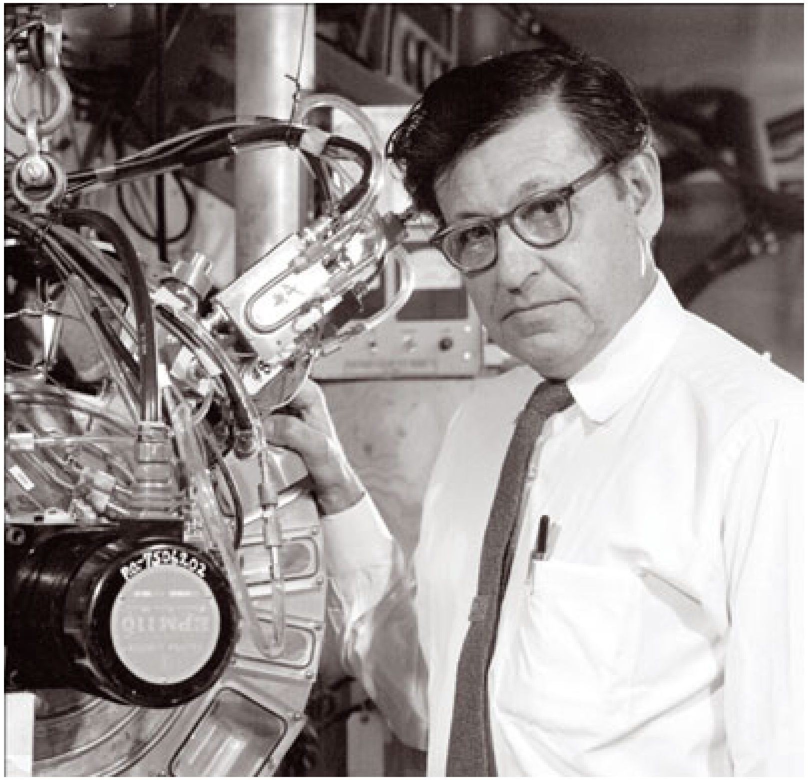
MIKE-2



75-76. Bikini Baker, a Nagasaki-scale atomic bomb, versus Mike over Manhattan. Superimposing the two explosions on the Manhattan skyline dramatizes the thousand-fold difference in destructiveness. The Baker fireball was the diameter of the neck of the mushroom cloud, about eight hundred yards. The Mike fireball alone expanded to more than three miles in diameter; its blast would have obliterated all five New York City boroughs.

Samples of the bomb debris were collected on filter papers by aircraft flying through the mushroom cloud

Using heavy ion reactions to make new elements – the Berkeley area

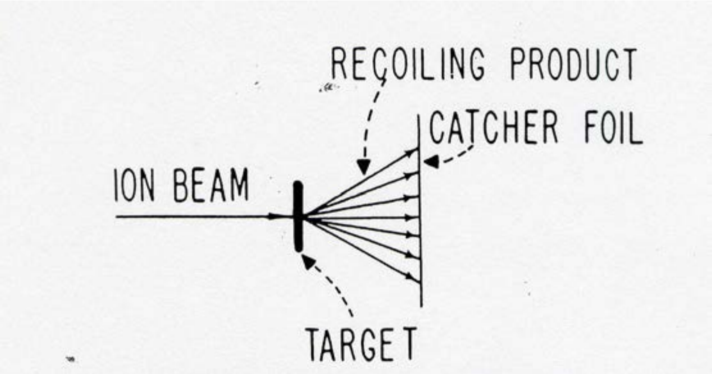


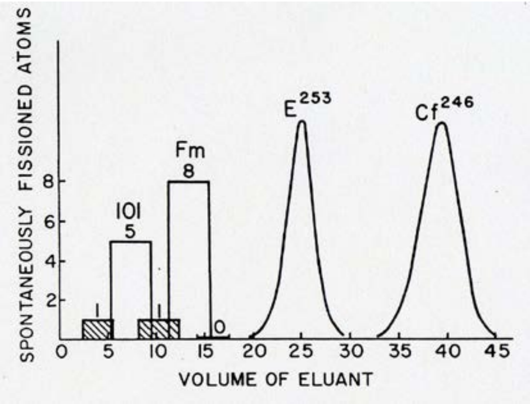
Albert Ghiorso

Glenn Seaborg

Synthesis of elements 101 - 106

- Making elements one atom at a time
- $^{254}Es + ^{4}He \rightarrow ^{256}Md + n$





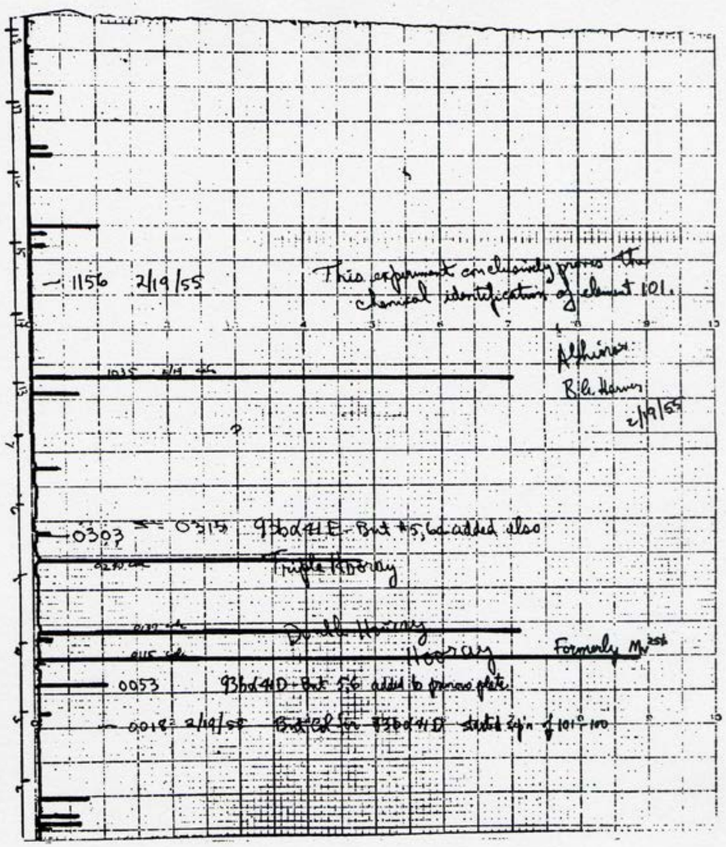


Figure 2.20 The ionization recording chart showing the first four events of the disintegration of mendelevium. The ordinate is the event time while the abscissa is the intensity of the ionization. The four pulses occurred at 1:15 a.m., 1:37 a.m., 2:40 a.m., and 10:35 a.m. on February 19, 1955. At 11:56 a.m., Ghiorso and Harvey made a note directly in the chart: "This experiment conclusively proves the chemical identification of element 101."

Hot fusion (1961-1974) successful up to element 106 (Seaborgium)

- **Coulomb barrier V_C** between projectile and target nucleus has to be exceeded

$$V_C = \frac{Z_p \cdot Z_t \cdot e^2}{R_{int}} = 126.2\ MeV \quad (^{26}Mg + {}^{248}Cm)$$

- reaction: $a + A \rightarrow C^* \rightarrow B + b$

$$\Delta m = m_a + m_A - m_{CN}$$

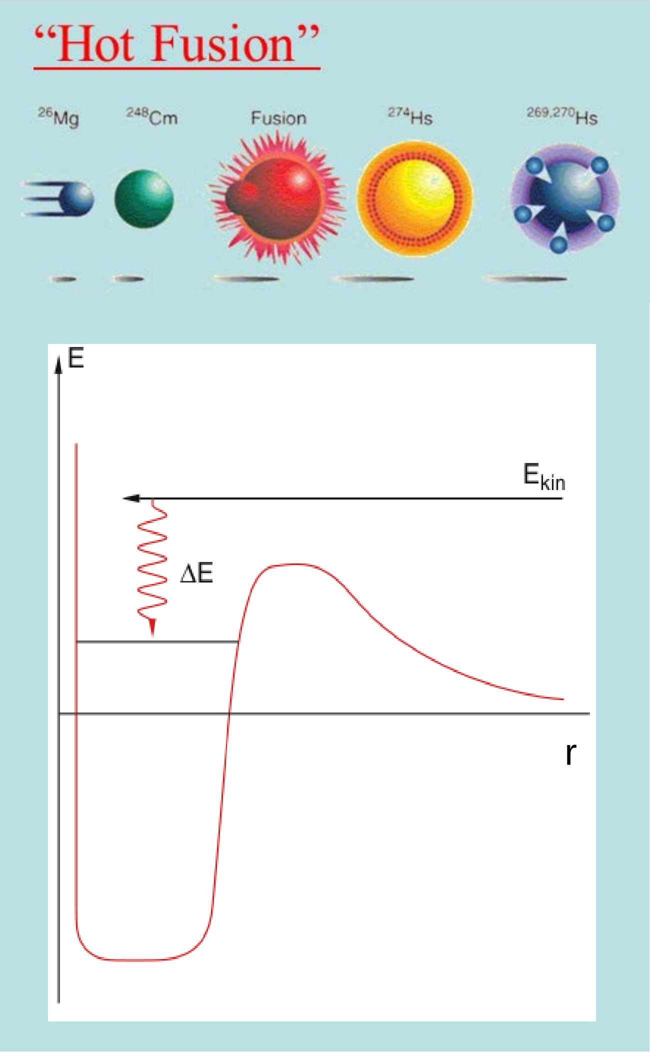
$$\Delta m = (25.983 + 248.072 - 274.143) * 931.478\ \mathrm{MeV/c^2}$$

$$= -82.153\ \mathrm{MeV/c^2}$$

- excitation energy of compound nucleus

$$E^* = E_{kin} + \Delta m \cdot c^2 \quad = 126.2\ \mathrm{MeV} - 82.2\ \mathrm{MeV} \quad \mathbf{= 44.0\ MeV}$$

- approximate 4 neutrons will be evaporated to avoid fission



http://nuclear.lu.se/database/masses/

The problem

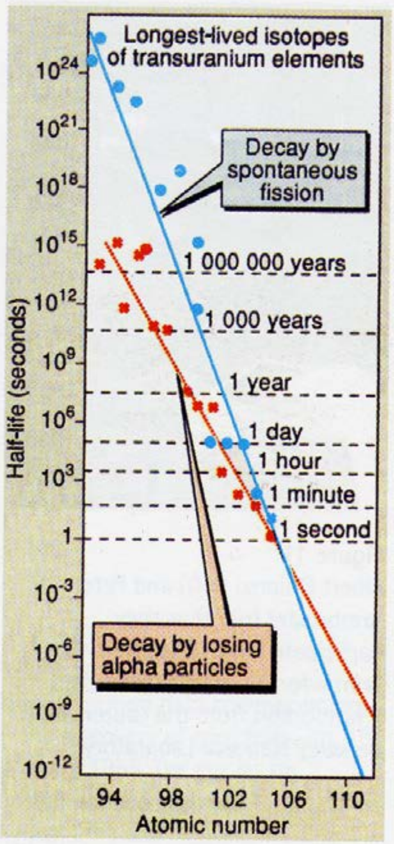
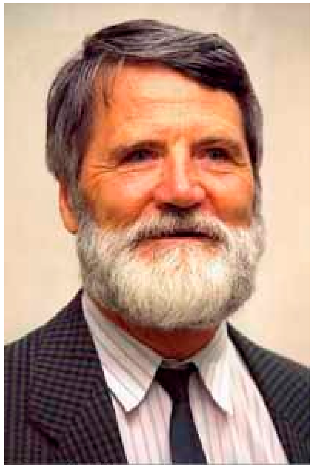


Figure 13
The half-lives of the longest-lived isotope of each element versus atomic number *Z*, *circa* 1970.

element	name	longest-lived isotope	half-life [s]
95	Americium	^{243}Am	$2.3\cdot10^{11}$
96	Curium	^{247}Cm	$5.0\cdot10^{14}$
97	Berkelium	^{247}Bk	$3.2\cdot10^{10}$
98	Californium	^{251}Cf	$2.8\cdot10^{10}$
99	Einsteinium	^{252}Es	$4.1\cdot10^{7}$
100	Fermium	^{257}Fm	$8.7\cdot10^{6}$
101	Mendelevium	^{258}Md	$4.4\cdot10^{6}$
102	Nobelium	^{259}No	$3.5\cdot10^{3}$
103	Lawrencium	^{266}Lr	$3.6\cdot10^{4}$
104	Rutherfordium	^{267}Rf	$4.7\cdot10^{3}$
105	Dubnium	^{268}Db	$1.1\cdot10^{5}$
106	Seaborgium	^{269}Sg	$1.8\cdot10^{2}$
107	Bohrium	^{270}Bh	$6.0\cdot10^{1}$
108	Hassium	^{277}Hs	$3.0\cdot10^{1}$
109	Meitnerium	^{278}Mt	$4.0\cdot10^{0}$
110	Darmstadtium	^{281}Ds	$1.4\cdot10^{1}$
111	Roentgenium	^{282}Rg	$1.2\cdot10^{2}$
112	Copernicium	^{285}Cn	$3.0\cdot10^{1}$
113	Nihonium	^{286}Nh	$8.0\cdot10^{0}$
114	Flerovium	^{289}Fl	$2.0\cdot10^{0}$
115	Moscovium	^{290}Mc	$8.0\cdot10^{-1}$
116	Livermorium	^{293}Lv	$6.0\cdot10^{-2}$
117	Tennessine	^{294}Ts	$5.0\cdot10^{-2}$
118	Organesson	^{294}Og	$7.0\cdot10^{-4}$

The solution – The Darmstadt Area

- “Cold fusion” reactions
- Bombarding Pb or Bi with heavy ions – the resulting species are borne “cold” - with low excitation energies – they survive better



Peter Armbruster

Gottfried Münzenberg

Sigurd Hofmann

Cold fusion (1981-1996)

- **Coulomb barrier V_C** between projectile and target nucleus has to be exceeded

$$V_C = \frac{Z_p \cdot Z_t \cdot e^2}{R_{int}} = 223.3\ MeV \quad (^{58}Fe + ^{208}Pb)$$

- reaction: $a + A \rightarrow C^* \rightarrow B + b$

$$\Delta m = m_a + m_A - m_{CN}$$

$$\Delta m = (57.933 + 207.977 - 266.130) * 931.478\ MeV/c^2$$

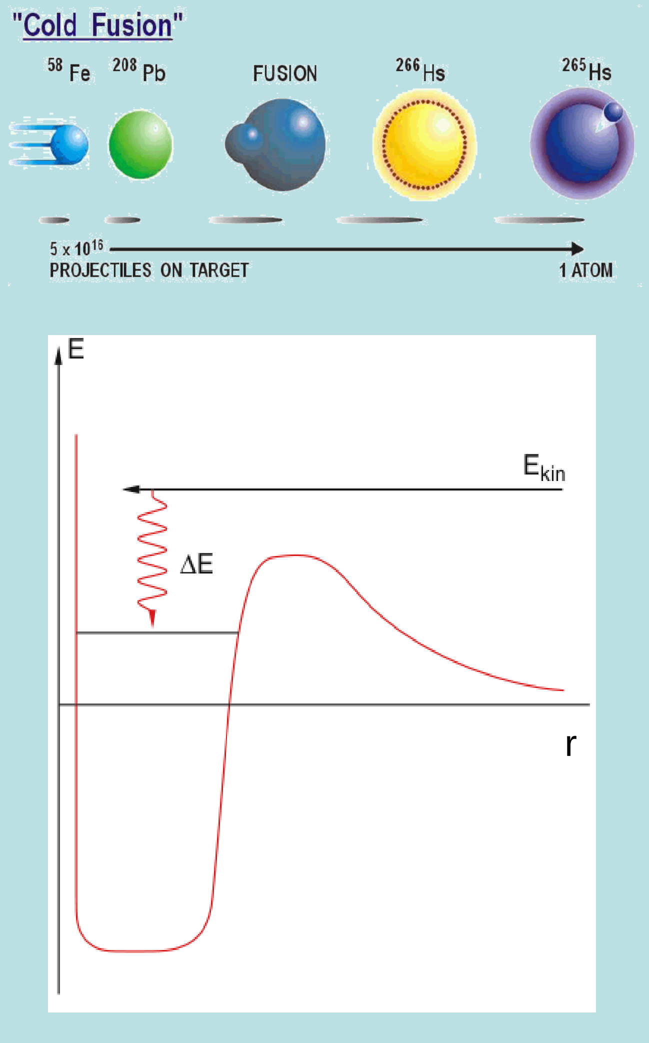
$$= -205.045\ MeV/c^2$$

- excitation energy of compound nucleus

$$E^* = E_{kin} + \Delta m \cdot c^2 \quad = 223.3\ MeV - 205.0\ MeV \quad \mathbf{= 18.2\ MeV}$$

- approximate 1-2 neutrons will be evaporated to avoid fission

reaction	Q_{gg} (MeV)	$V_C(R_{int})$ (MeV)	E* (MeV)
$^{48}_{20}Ca + ^{208}_{82}Pb \rightarrow ^{256}_{102}No$	-153.796	175.05	21.25
$^{54}_{24}Cr + ^{209}_{83}Bi \rightarrow ^{263}_{107}Bh$	-189.911	210.00	20.09
$^{58}_{26}Fe + ^{208}_{82}Pb \rightarrow ^{266}_{108}Hs$	-205.045	223.31	18.27
$^{58}_{26}Fe + ^{209}_{83}Bi \rightarrow ^{267}_{109}Mt$	-208.526	225.86	17.33
$^{62}_{28}Ni + ^{208}_{82}Pb \rightarrow ^{270}_{110}Ds$	-223.225	238.84	15.62
$^{64}_{28}Ni + ^{209}_{83}Bi \rightarrow ^{273}_{111}Rg$	-228.52	240.78	12.26
$^{70}_{30}Zn + ^{208}_{82}Pb \rightarrow ^{278}_{112}Cn$	-244.38	252.67	8.29



http://nuclear.lu.se/database/masses/

G. Audi et al. Nucl. Phys. A729 (2003) 337

Excitation functions

G. Münzenberg / Nuclear Physics A 693 (2001) 207–218 209

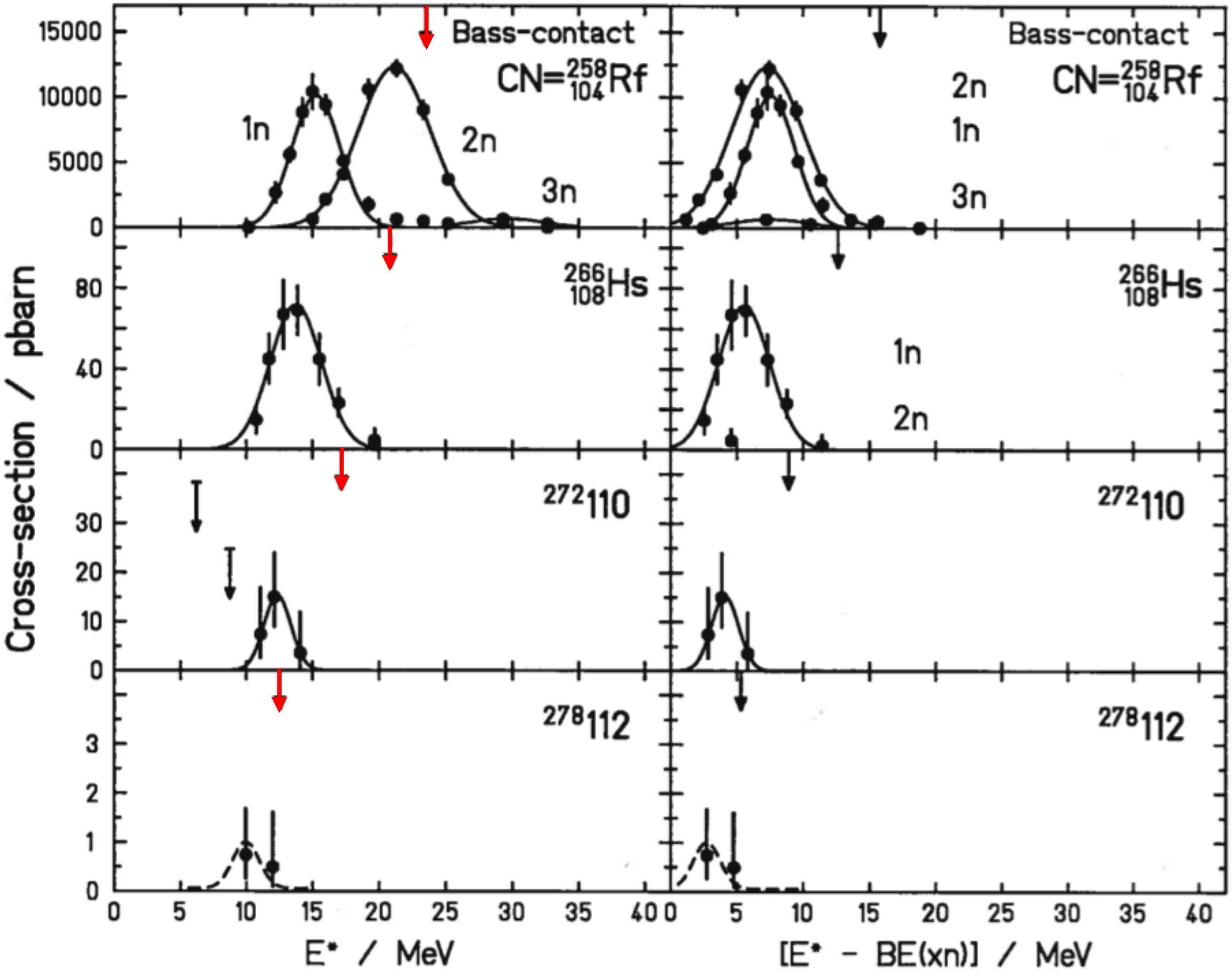


Fig. 1. Excitation functions for the production of elements 104, 108, 110, and 112 plotted versus the excitation energy of the compound system (left-hand side) and versus the excitation energy corrected by subtraction of the neutron binding energy [6] (right-hand side).

$${}^{50}_{22}Ti + {}^{208}_{82}Pb \rightarrow {}^{258}_{104}Rf$$

Q_{gg} = -169.5 MeV

V_C = 194.2 MeV $_{Bass}$

$${}^{58}_{26}Fe + {}^{208}_{82}Pb \rightarrow {}^{266}_{108}Hs$$

Q_{gg} = -205.0 MeV

V_C = 227.1 MeV $_{Bass}$

$${}^{64}_{28}Ti + {}^{208}_{82}Pb \rightarrow {}^{272}_{110}Ds$$

Q_{gg} = -224.9 MeV

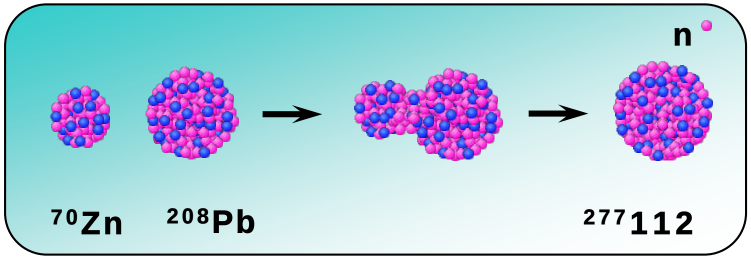
V_C = 242.2 MeV $_{Bass}$

$${}^{70}_{30}Zn + {}^{208}_{82}Pb \rightarrow {}^{278}_{112}Cn$$

Q_{gg} = -244.2 MeV

V_C = 257.2 MeV $_{Bass}$

Synthesis of heavy elements

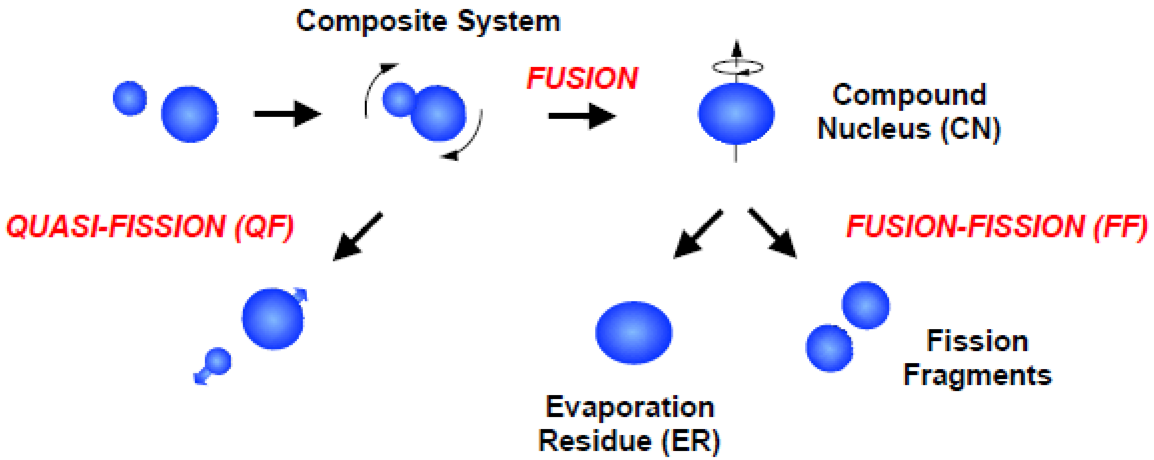


Fusion

Compound Fission

$$\frac{1}{10^{12}}$$

Fusion / fission competition



Composite System

FUSION

Compound Nucleus (CN)

QUASI-FISSION (QF)

FUSION-FISSION (FF)

Evaporation Residue (ER)

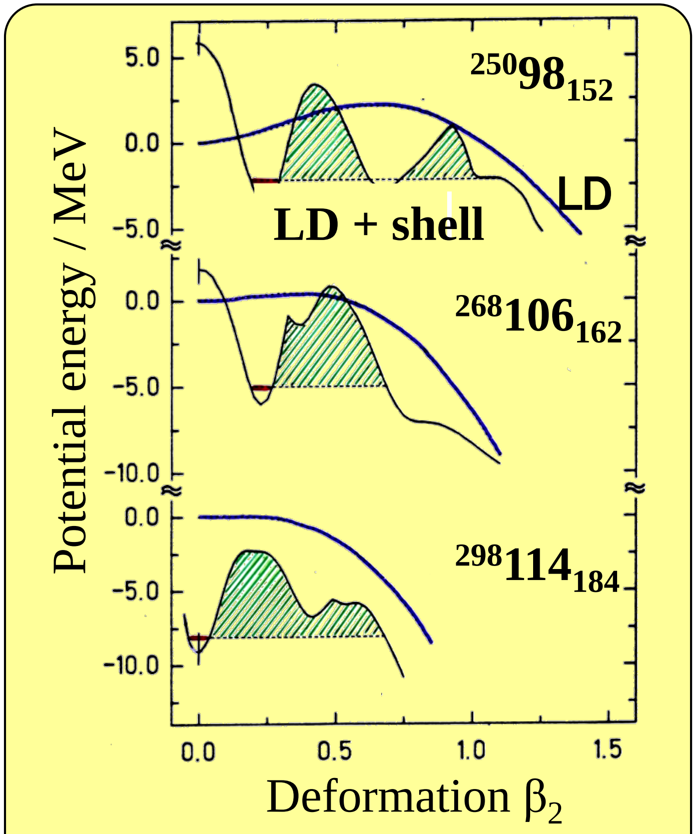
Fission Fragments

$$\boxed{\sigma_{ER} = \sigma_{capture} \cdot P_{CN} \cdot P_{survival}}$$

$$\sigma_{ER}^{xn}(E) = \frac{\pi}{k^2}\sum_{\ell=0}^{\infty}(2\ell+1)\cdot \underbrace{P_{cont}(E,\ell)}_{\text{capture}} \cdot \underbrace{P_{CN}(E^*,\ell)}_{\text{formation}} \cdot \underbrace{P_{xn}(E^*,\ell)}_{\text{survival}}$$

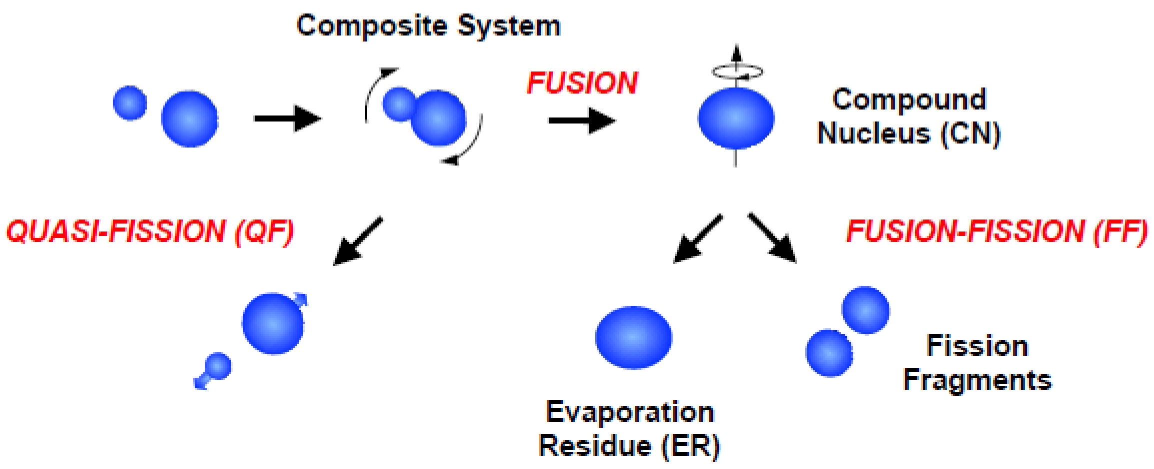
competing: quasi-elastic quasi-fission fission

Superheavy system: $\sigma_{ER} \ll \sigma_{capture}$



A. Sobiczewski et al.

Fusion / fission competition



Composite System

FUSION

Compound Nucleus (CN)

QUASI-FISSION (QF)

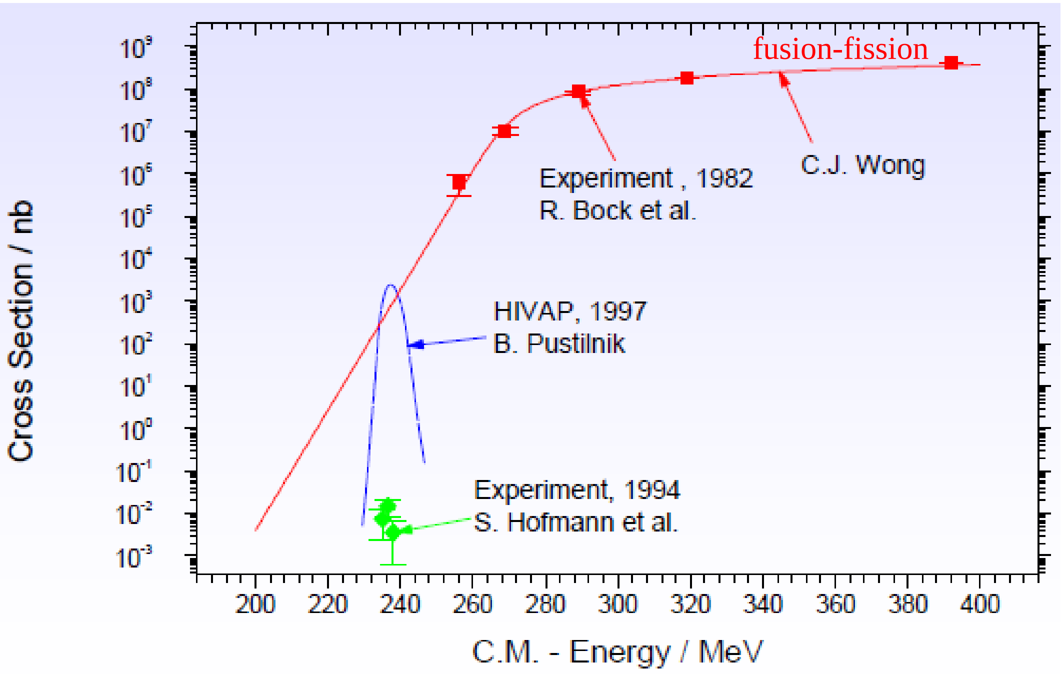
FUSION-FISSION (FF)

Fission Fragments

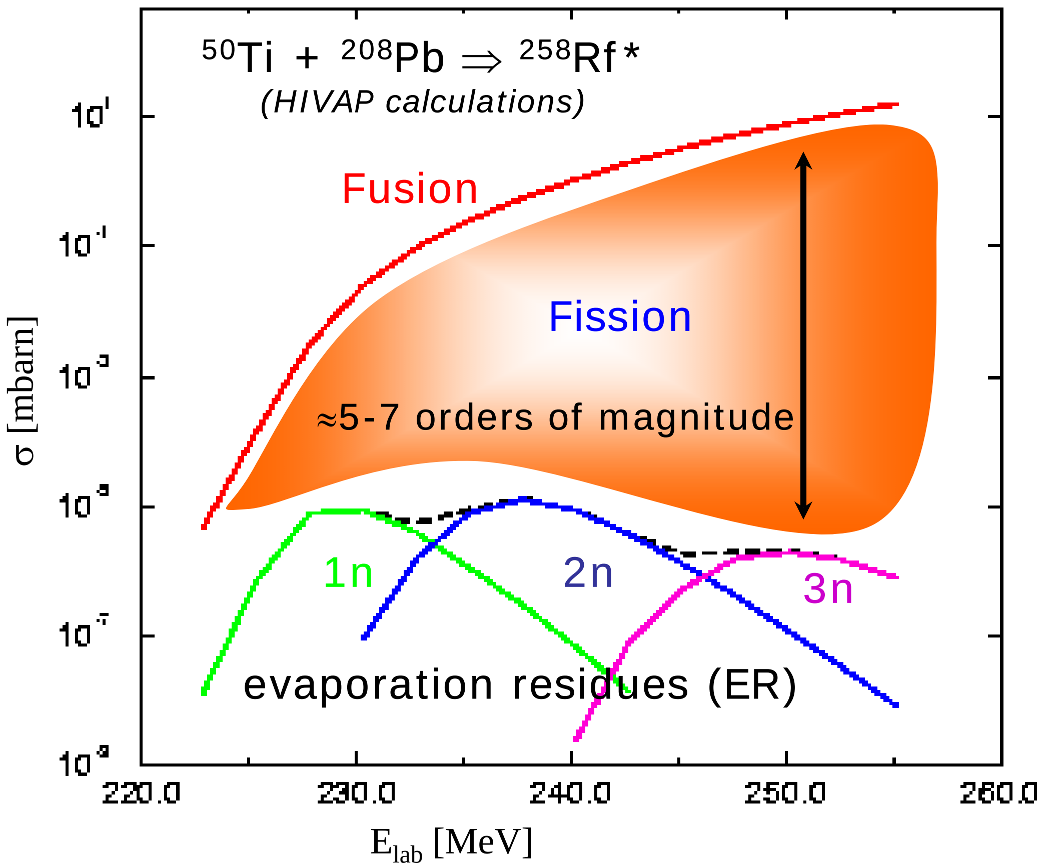
Evaporation Residue (ER)

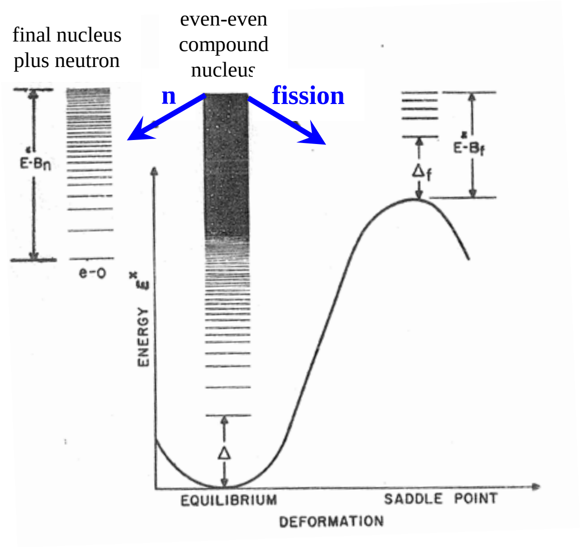
$$\sigma_{ER} = \sigma_{capture} \cdot P_{CN} \cdot P_{survival}$$

$^{64}\text{Ni} + {}^{208}\text{Pb} \rightarrow {}^{271}\text{Ds} + 1\text{n}$ $V_C(R_{int}) = 238\ MeV$



Fusion and evaporation





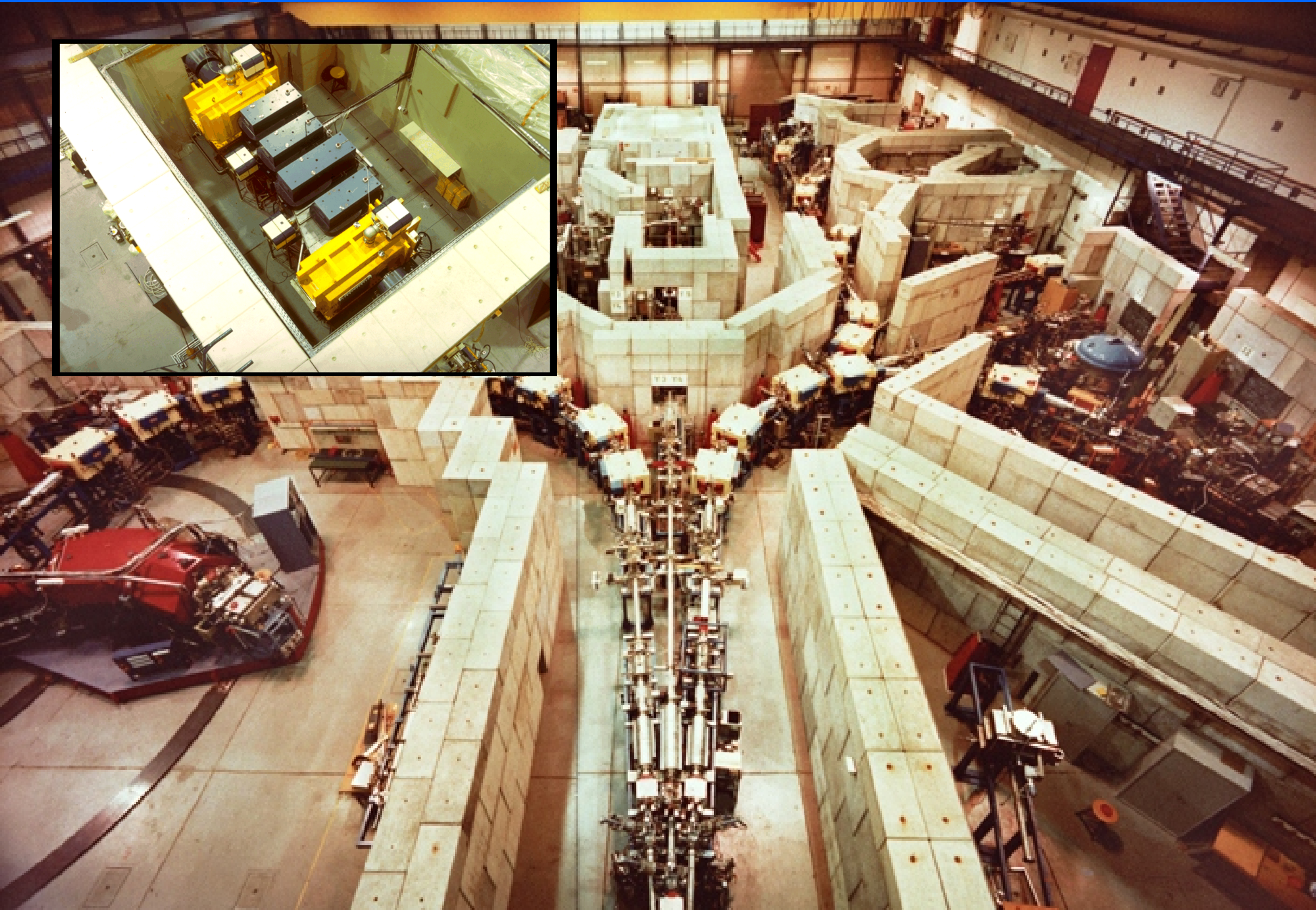
Both decay processes are determined by the level density, either from the residual nucleus or at the saddle point.

level density: $\rho(E^*) = const \cdot \exp(E^*/T)$

$$\frac{\Gamma_n}{\Gamma_f} = \frac{2 \cdot T \cdot A_{CN}^{2/3}}{K_0} \cdot \exp\left[(B_f - B_n)/T\right]$$

$K_0 = \hbar^2/2 \cdot m \cdot r_0^2 \approx 11.4\ MeV$ $\qquad T = \sqrt{8 \cdot E^* / A_{CN}}$

Separator for Heavy Ion Products (SHIP)



Separator for Heavy Ion Products (SHIP)

- Fusion products are slower than scattered or transfer particles

$$v_{CN} = [m_p/(m_p + m_t)] \cdot v_p$$

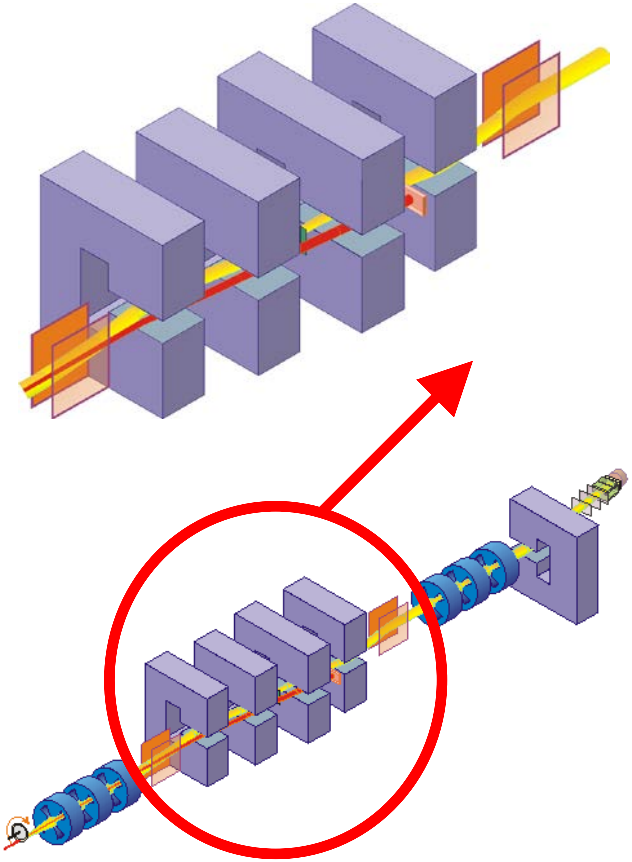
$$e.q.\, v_p \approx 10.3\% \quad \rightarrow \quad v_{CN} \approx 2.2\%$$

- E- and B-field are perpendicular to each other

$$B \cdot \rho = \frac{m \cdot v}{e \cdot q}$$

$$E \cdot \rho = \frac{m \cdot v^2}{e \cdot q}$$

$$F_{mag} = F_{el} \Rightarrow F_{tot} = 0$$



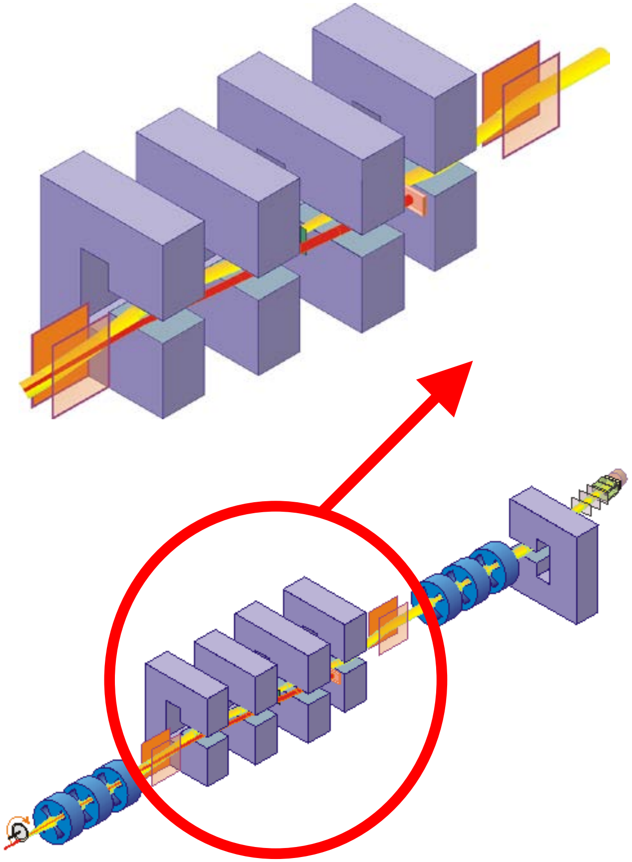
electric deflectors: ±330 kV dipole magnets: 0.7 T max

Separator for Heavy Ion Products (SHIP)

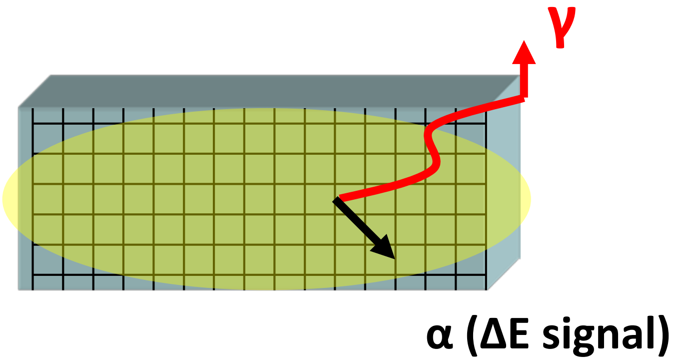
- The choice of E and B determines the transmitted velocity

$$v = \frac{E}{B}$$

- The rejected beam will be stopped on a cooled Cu plate



SHIP – stop detector



- position sensitive Silicon detector determines the position an energy of SHE and α, β, ...

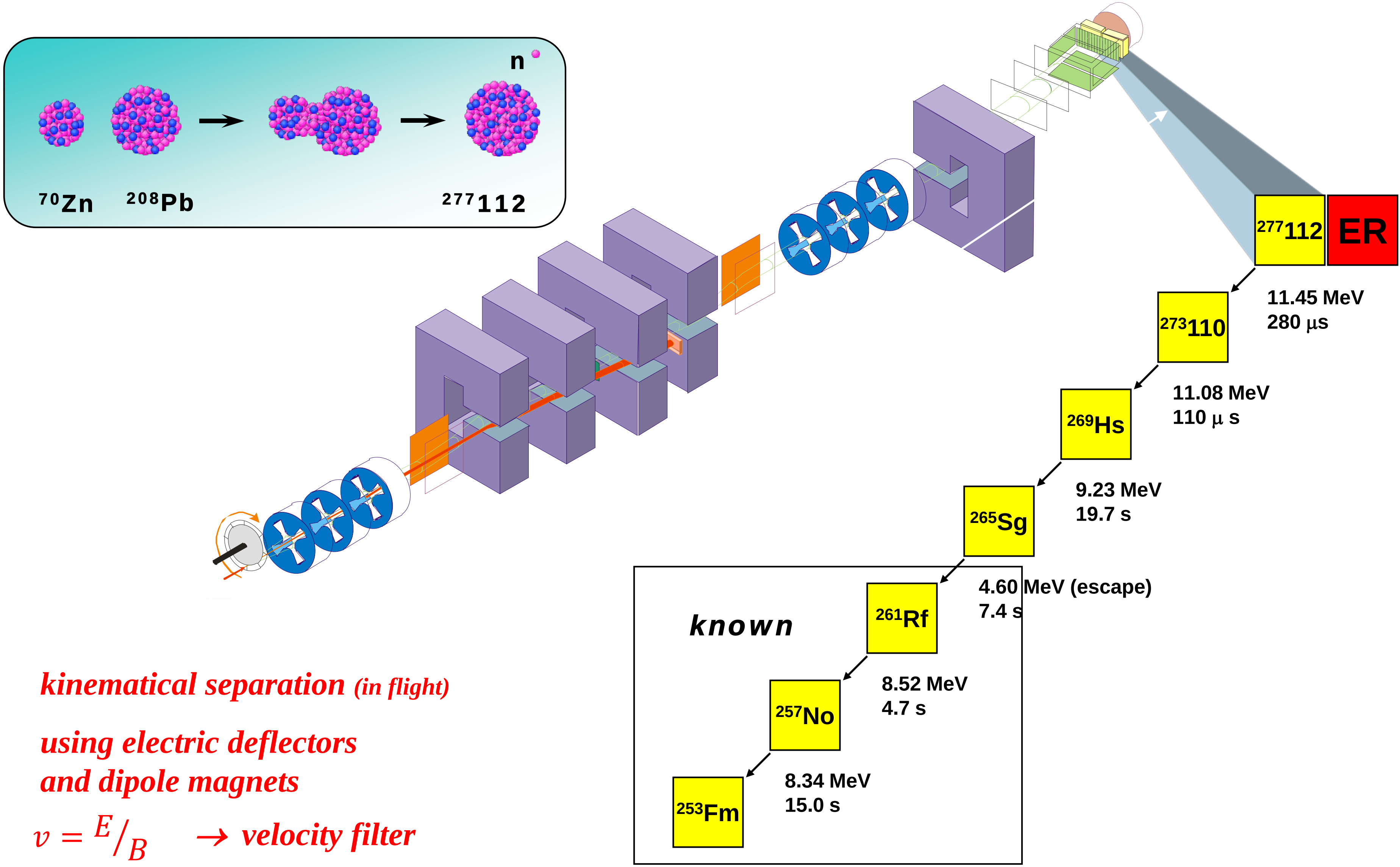
area: 27*87mm^2, thickness: 0.3mm, 16 strips
energy resolution ΔE=18-20 keV @ E_α> 6MeV (cooling 260K)
position resolution Δx=0.3mm (FWHM)

SHE will be measured in a pixel

Wait for the emission of an α-particle
(or β-particle)

correlation method: implantation and decay event in the same pixel

Synthesis and identification of heavy elements with SHIP



^{277}112 ER

11.45 MeV
280 μs

^{273}110

11.08 MeV
110 μ s

^{269}Hs

9.23 MeV
19.7 s

^{265}Sg

4.60 MeV (escape)
7.4 s

known

^{261}Rf

8.52 MeV
4.7 s

^{257}No

8.34 MeV
15.0 s

^{253}Fm

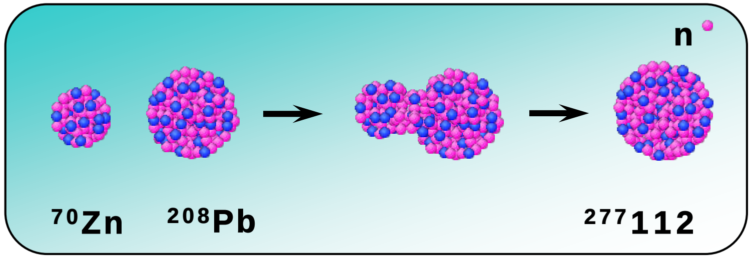
Date: 09-Feb-1996
Time: 22:37 h

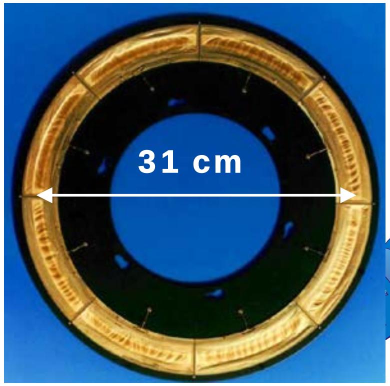
kinematical separation (in flight)

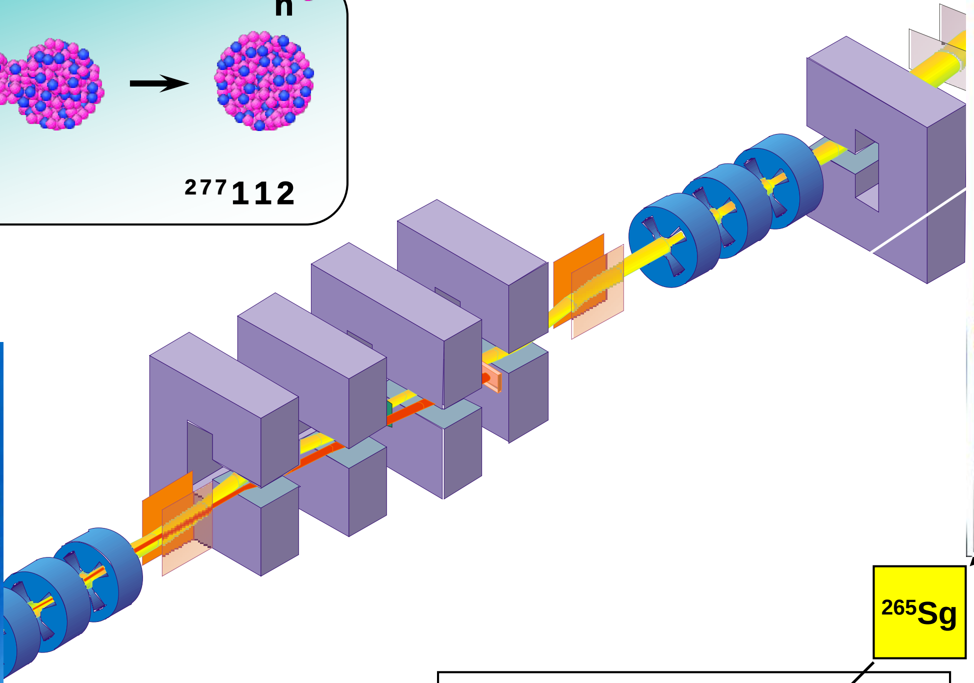
using electric deflectors
and dipole magnets

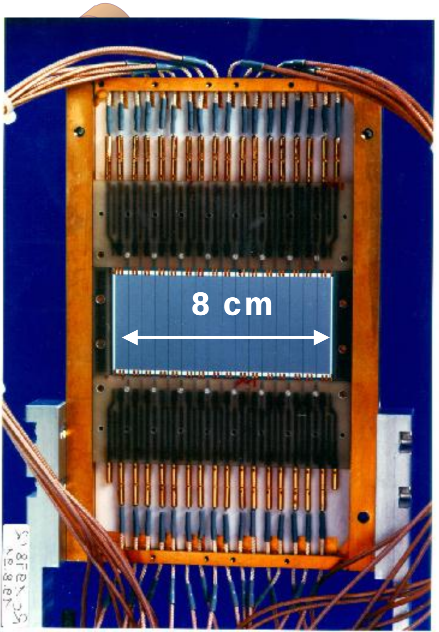
$v = E/B$ → *velocity filter*

Synthesis and identification of heavy elements with SHIP









^{265}Sg

4.60 MeV (
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known

^{261}Rf

8.52 MeV
4.7 s

^{257}No

8.34 MeV
15.0 s

^{253}Fm

Date: 09-Feb-1996
Time: 22:37 h

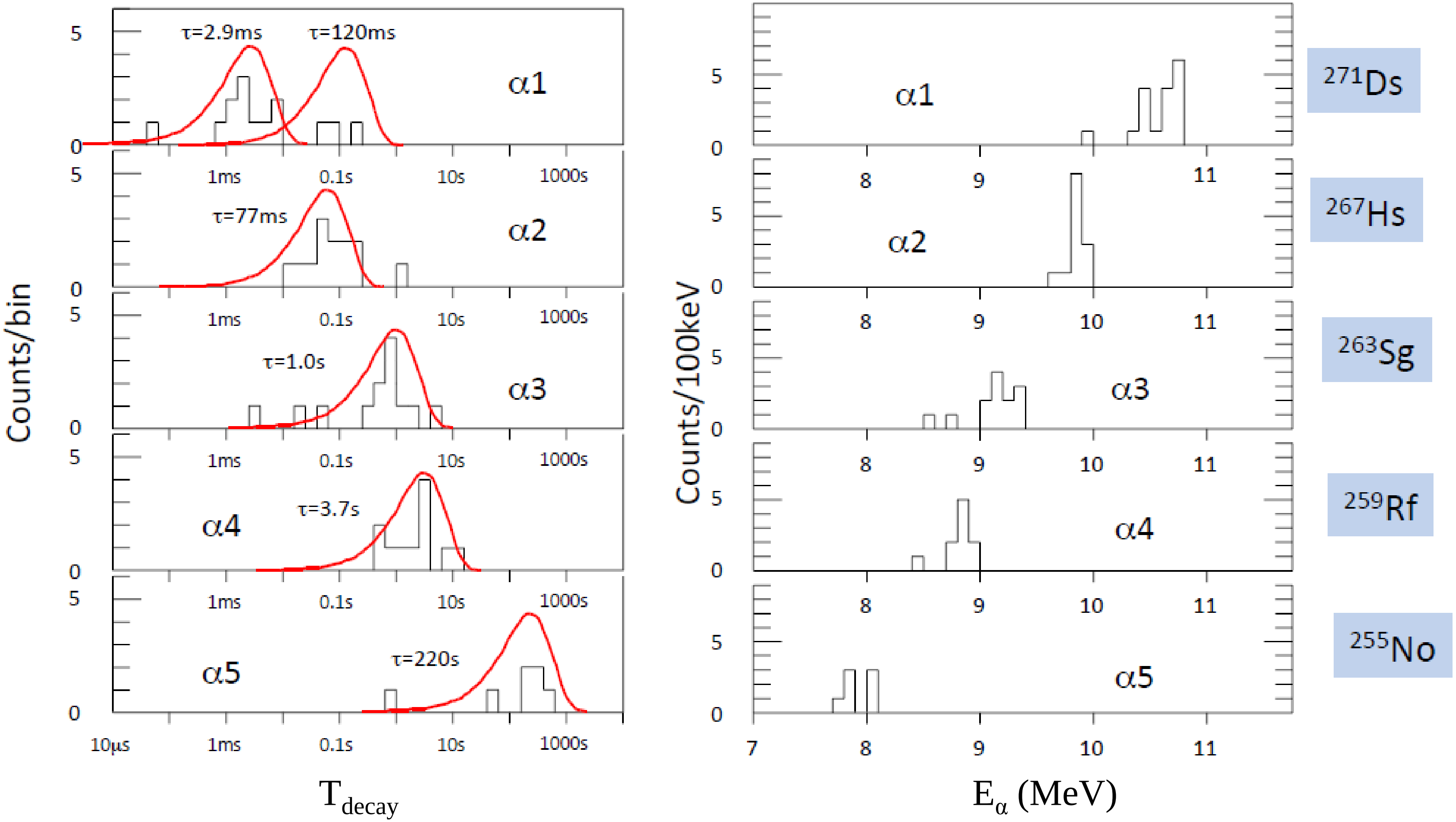
Identification by α-α correlations down to __known__ isotopes

kinematical separation (in flight)

using electric deflectors
and dipole magnets

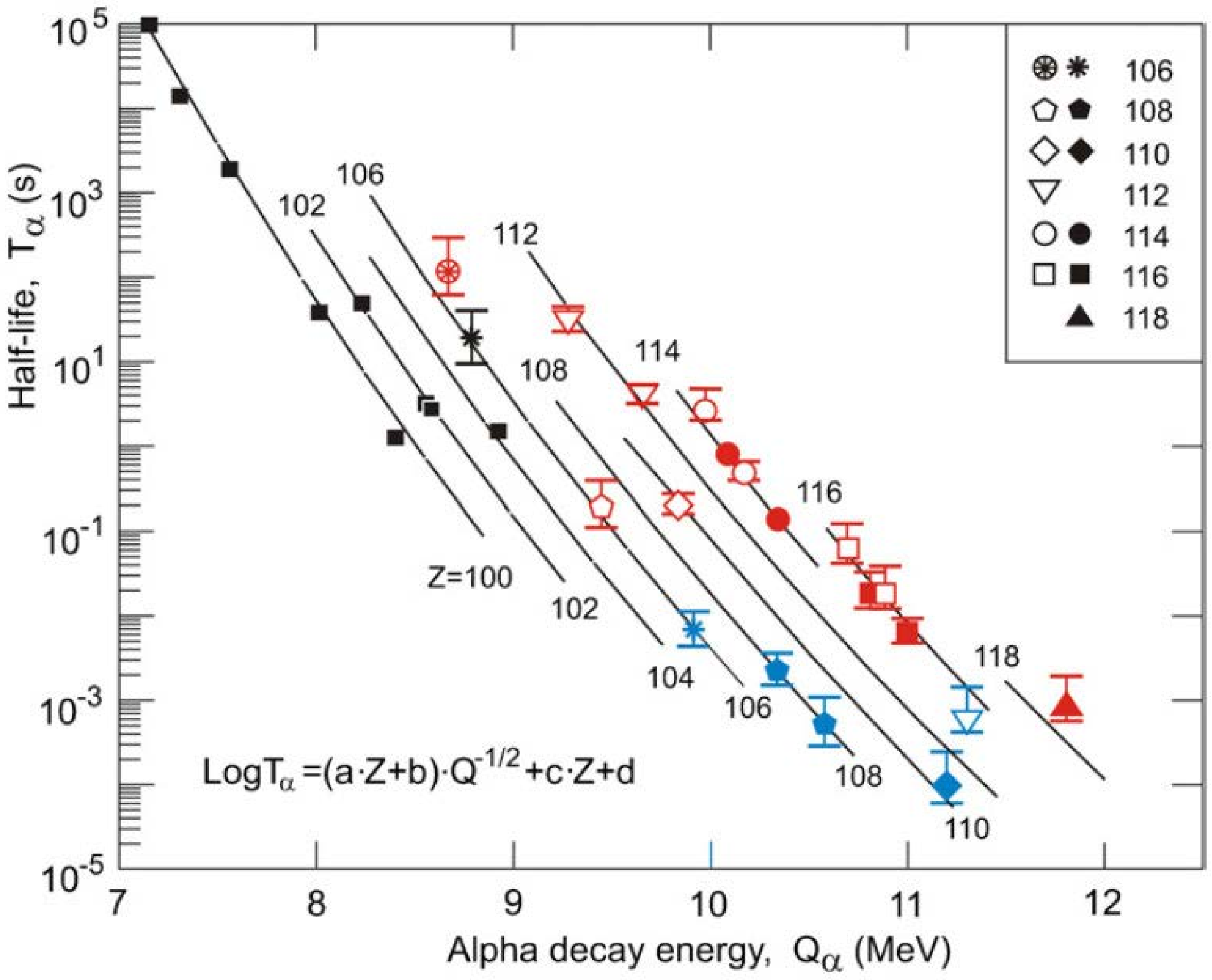
$v = {}^{E}/_{B}$ → *velocity filter*

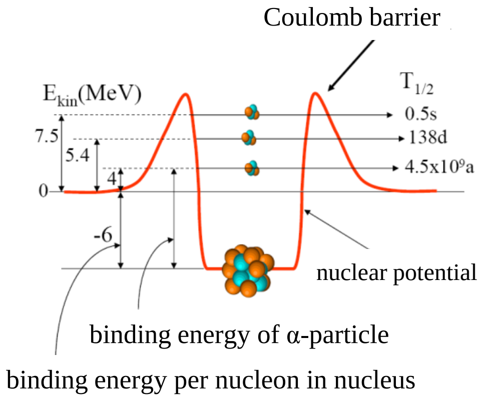
^{208}Pb + ^{64}Ni → ^{272}Ds*



Geiger-Nuttall relationship

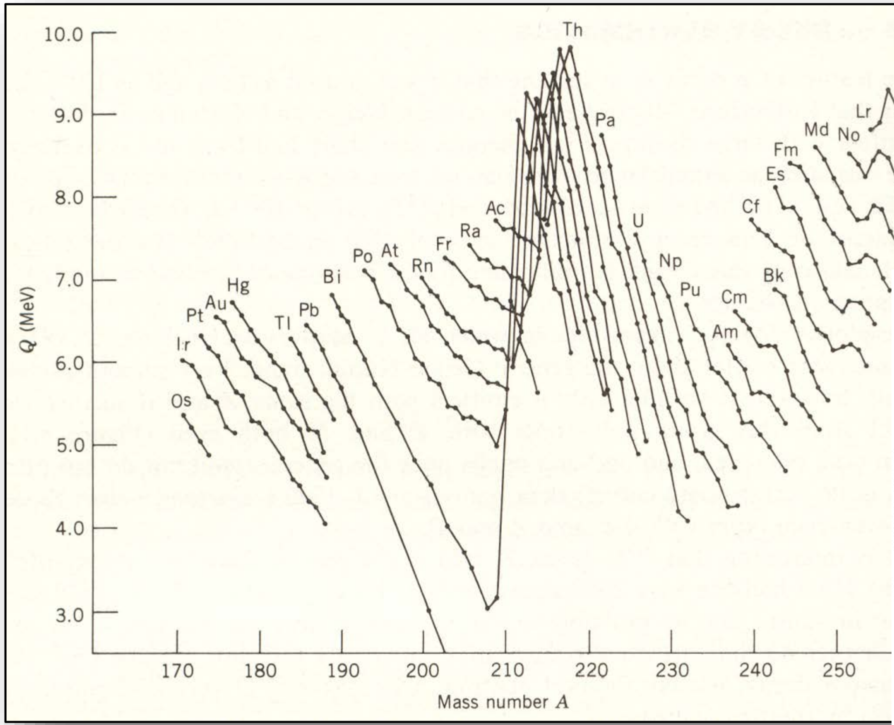
- ❖ The average decay properties of even mass decay chains match the Geiger-Nuttall relationship



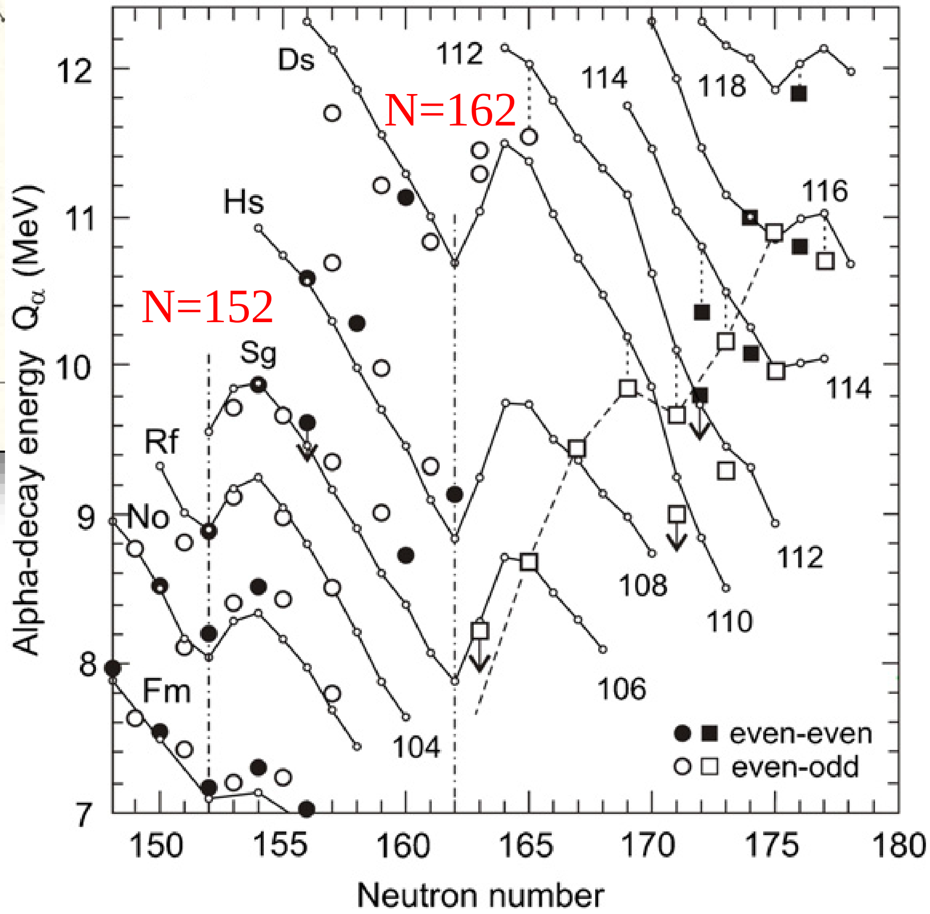


Y. Oganessian; J. Phys. G: Nucl. Part. Phys. 34 (2007) E165

Increase of Q_α values for the isotopes with Z = 112-118



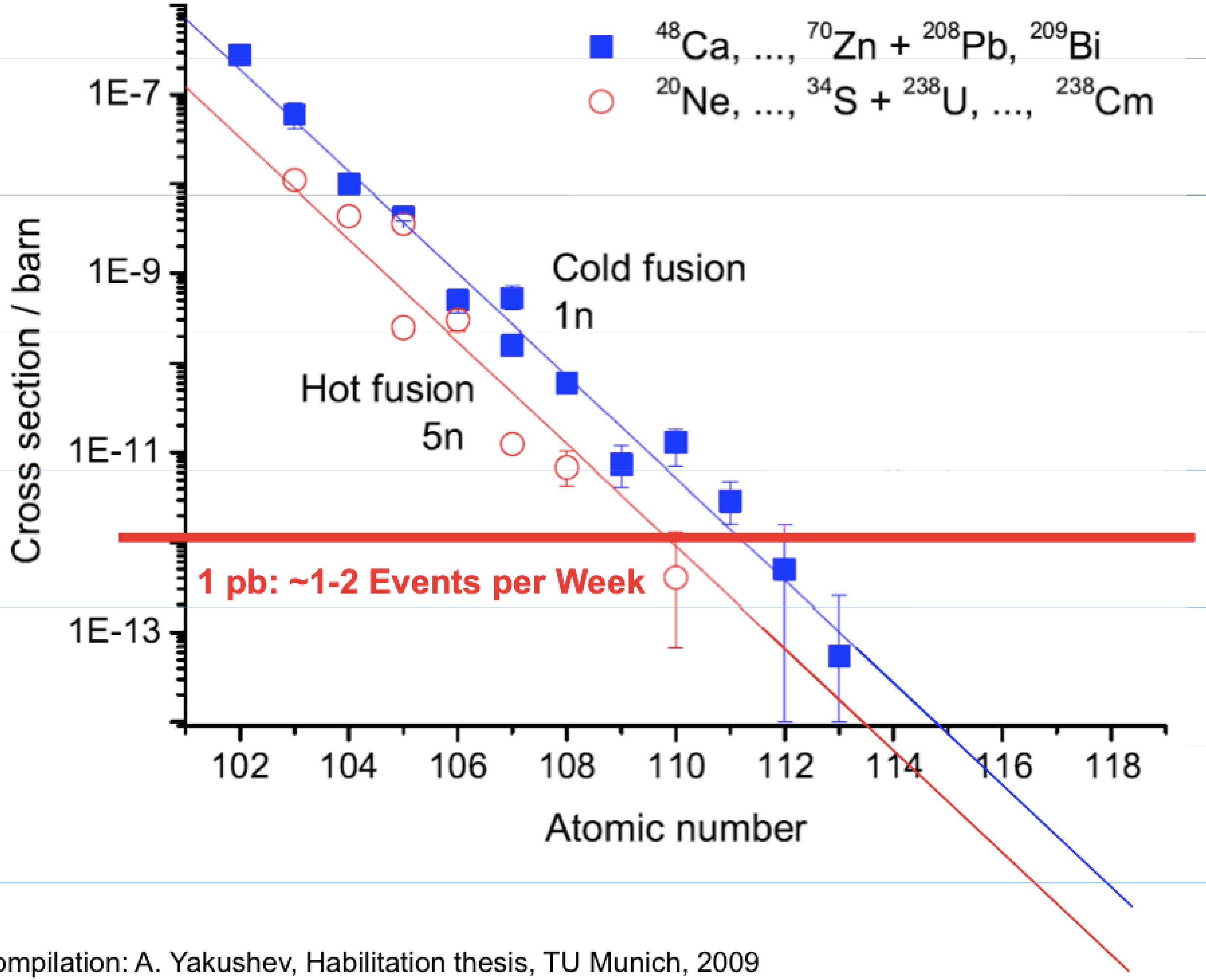
deformed shell closure



Micro-macro predictions:
nuclei with N ≥ 175 are close to spherical ($\beta_2 \approx 0.09$)

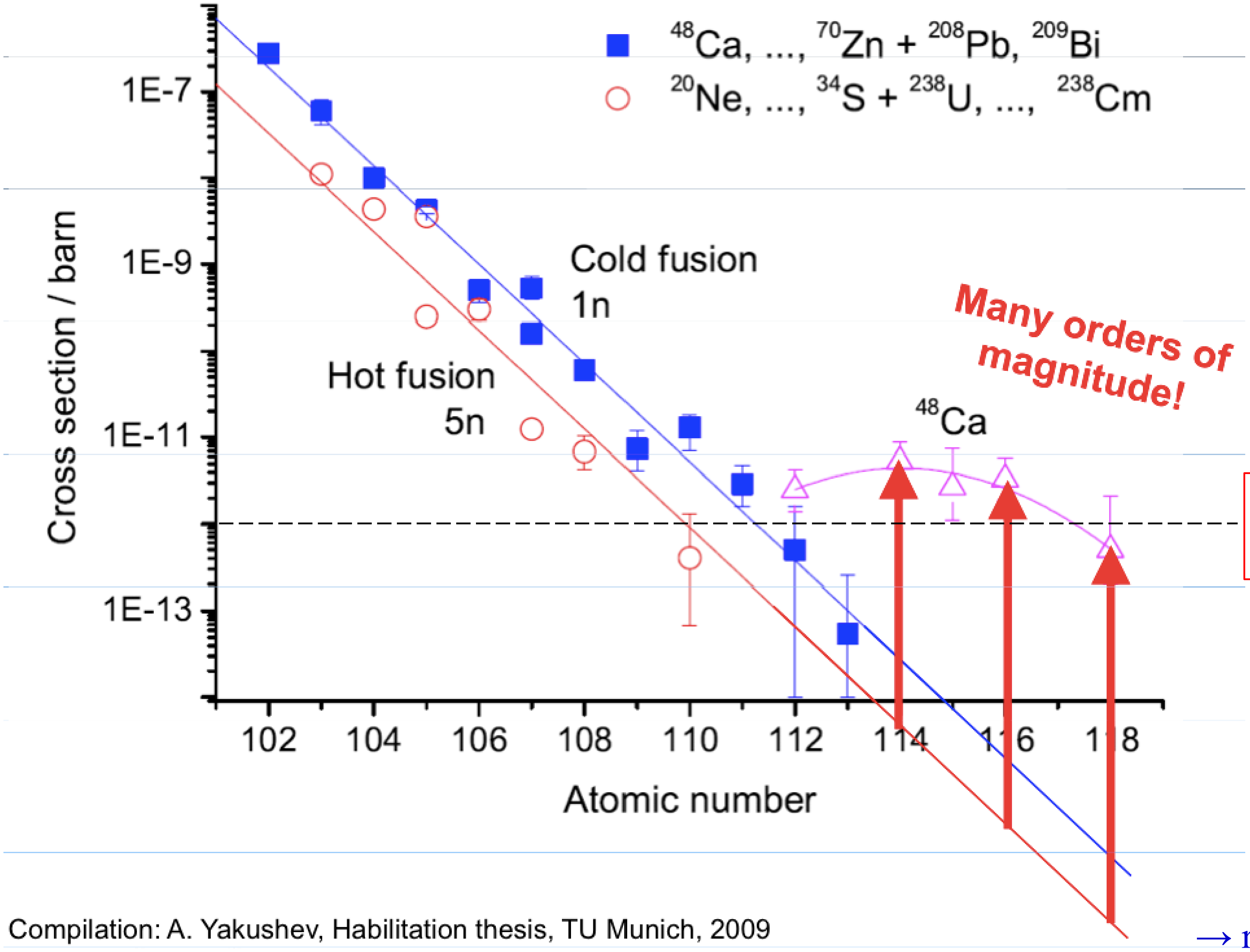
Y. Oganessian; J. Phys. G: Nucl. Part. Phys. 34 (2007) E165

The end of the “cold fusion” path?



Compilation: A. Yakushev, Habilitation thesis, TU Munich, 2009

"Hot fusion" – The Dubna Era



Compilation: A. Yakushev, Habilitation thesis, TU Munich, 2009

Yuri Oganessian

1 pico barn →
1 nucleus / week

→ requires very efficient separation and detection techniques

Using ^{48}Ca beams with actinide targets

Joint Institute of Nuclear Research (Dubna) has extended the periodic table to Z=118

- ^{48}Ca projectiles produced by U400 heavy ion accelerator

energy: 235 -250 MeV
beam intensity: 1.0 – 1.5 pμA
consumption: 0.5 – 0.8 mg/h
beam dose: $(0.3 – 3.0)\cdot 10^{19}$

year	element	reaction	number of atoms
2000	114	$^{48}Ca \rightarrow {}^{244}Pu$	50
2004	113	decay product of Z=115	8
2004	115	$^{48}Ca \rightarrow {}^{243}Am$	30
2005	116	$^{48}Ca \rightarrow {}^{248}Cm$	30
2006	118	$^{48}Ca \rightarrow {}^{249}Cf$	3 - 4
2010	117	$^{48}Ca \rightarrow {}^{249}Bk$	6

prices per 1 mg

$^{197}Au \approx 0.03$ \$
$^{239}Pu \approx 4$ \$
$^{48}Ca \approx 80$ \$
$^{249}Cf \approx 60\,000$ \$

Using ^{48}Ca beams with actinide targets

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energy: 235 -250 MeV
beam intensity: 1.0 – 1.5 pμA
consumption: 0.5 – 0.8 mg/h
beam dose: (0.3 – 3.0)·10^{19}

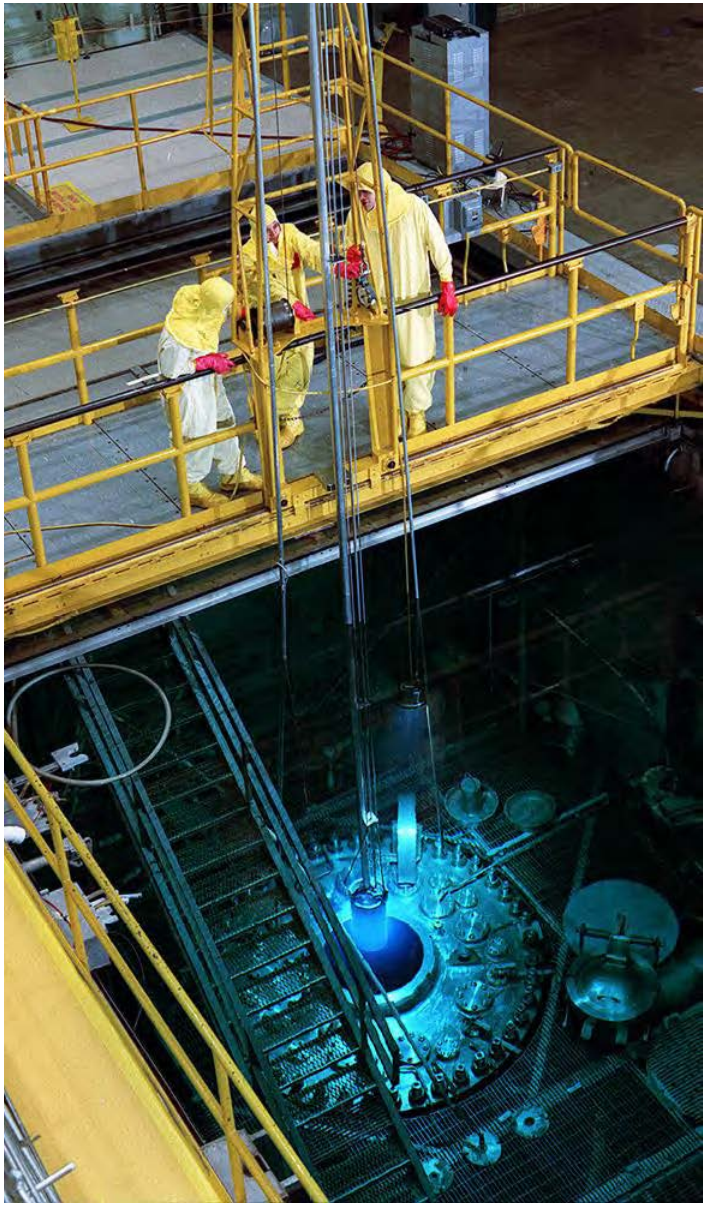
reaction	Q_{gg} (MeV)	$V_C(R_{int})$ (MeV)	E^* (MeV)
${}^{48}_{20}Ca + {}^{244}_{94}Pu \rightarrow {}^{292}_{114}Fl$	-160.5	197.3	36.8
${}^{48}_{20}Ca + {}^{243}_{95}Am \rightarrow {}^{291}_{115}Mc$	-170.6	199.7	29.1
${}^{48}_{20}Ca + {}^{248}_{96}Cm \rightarrow {}^{296}_{116}Lv$	-166.6	201.1	34.5
${}^{48}_{20}Ca + {}^{249}_{97}Bk \rightarrow {}^{297}_{117}Ts$	-170.1	203.2	33.1
${}^{48}_{20}Ca + {}^{249}_{98}Cf \rightarrow {}^{297}_{118}Og$	-174.3	205.4	31.1

Q_{gg} and V_C(Bass) from http://nrv.jinr.ru/nrv/webnrv/qcalc/

Irradiation of targets at HFIR reactor (Oak Ridge)

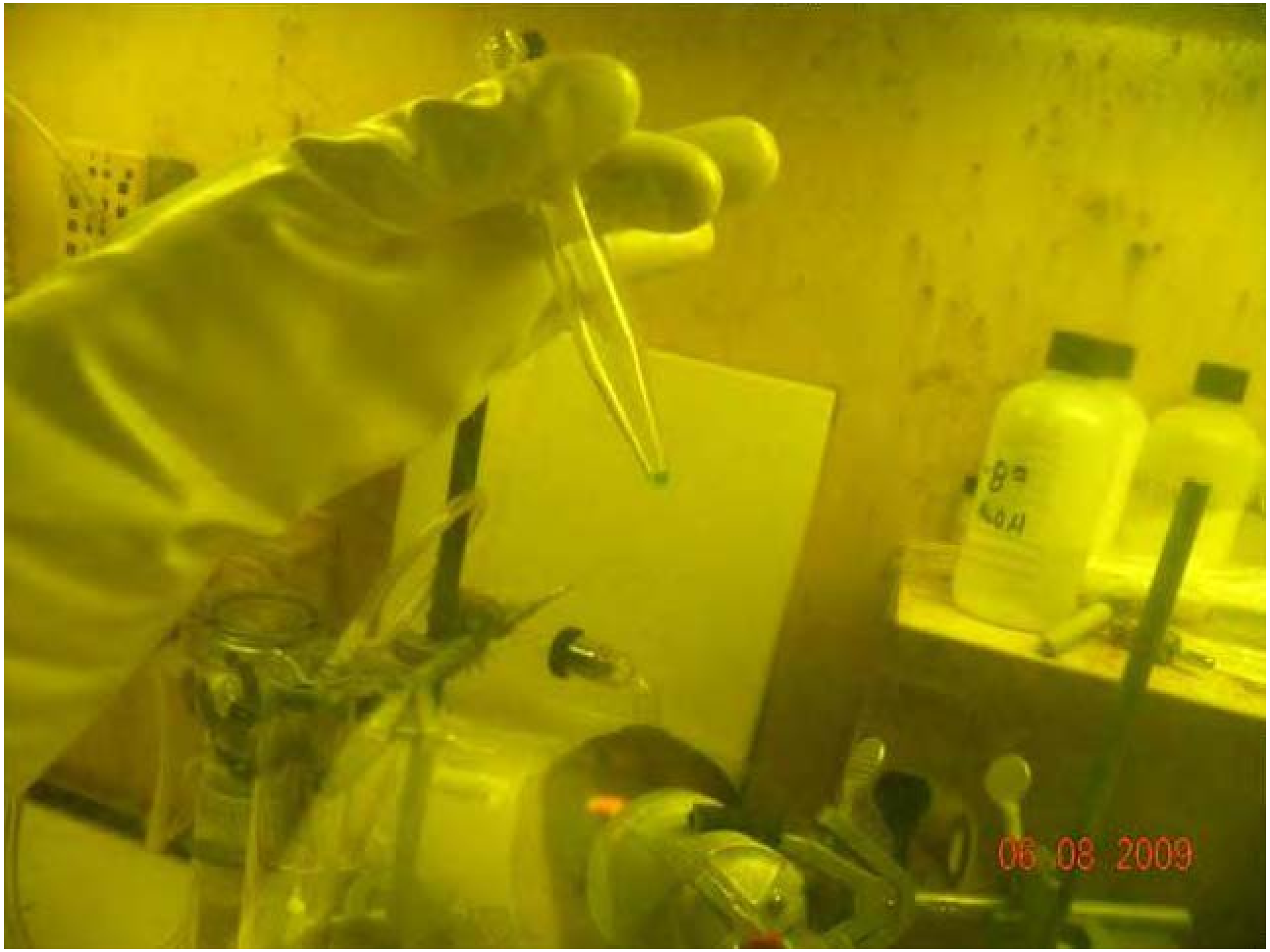
$$ {}^{238}_{92}U + {}^{1}_{0}n \longrightarrow {}^{239}_{92}U \xrightarrow[23.5\ min]{\beta^-} {}^{239}_{93}Np \xrightarrow[2.357\ d]{\beta^-} {}^{239}_{94}Pu $$

$$ {}^{239}_{94}Pu \xrightarrow{(n,\gamma)} {}^{240}_{94}Pu \xrightarrow{(n,\gamma)} {}^{241}_{94}Pu \xrightarrow[14.35\ a]{\beta^-} {}^{241}_{95}Am \left(\xrightarrow[432.2\ a]{\alpha} {}^{237}_{93}Np \right) $$



- Irradiation in the HFIR flux trap (18 month)
  - thermal-neutron flux of $2.5 \cdot 10^{15}$ neutrons/(s·cm^2)
  - 31 target positions (10 – 13 targets typically irradiated)
  - produces ~35 mg ^{252}Cf per target (smaller quantities of Bk, Es, Fm)

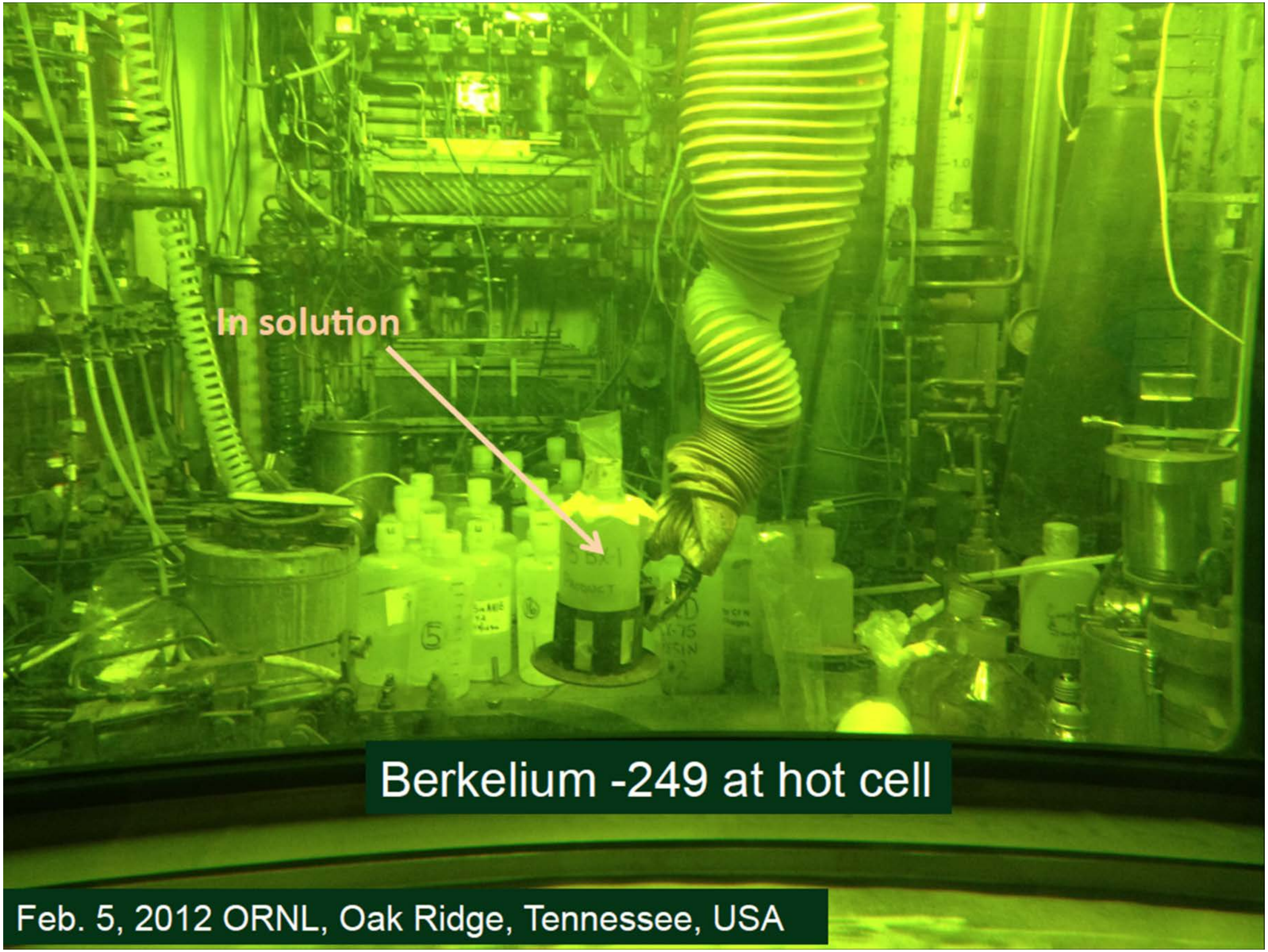
22 mg of ^{249}Bk ≈ 1 M$, 250 day irradiation in HIFR (ORNL)



$Bk(NO_3)_3$ product

- The two year experimental campaign began with a 250 day irradiation in HFIR, producing 22 milligram of ^{249}Bk, which has a 320 day half-life. The irradiation was followed by 90 days of processing at radiochemical Engineering Development Center (REDC) to separate and purify the Berkelium. The ^{249}Bk target was prepared at Dimitrovgrad and then bombarded for 150 days at the Dubna facility.

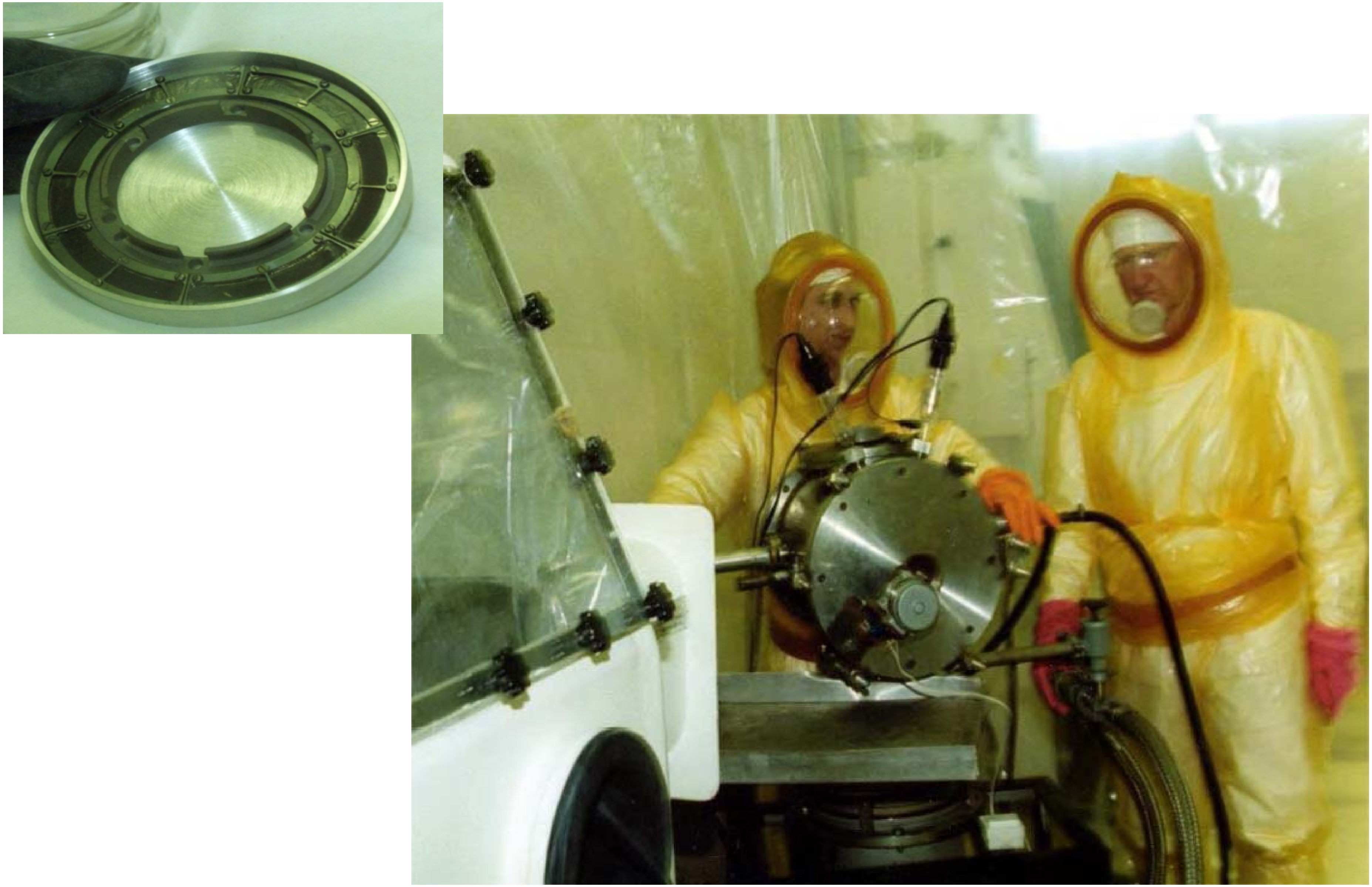
22 mg of ^{249}Bk ≈ 1 M$, 250 day irradiation in HIFR (ORNL)



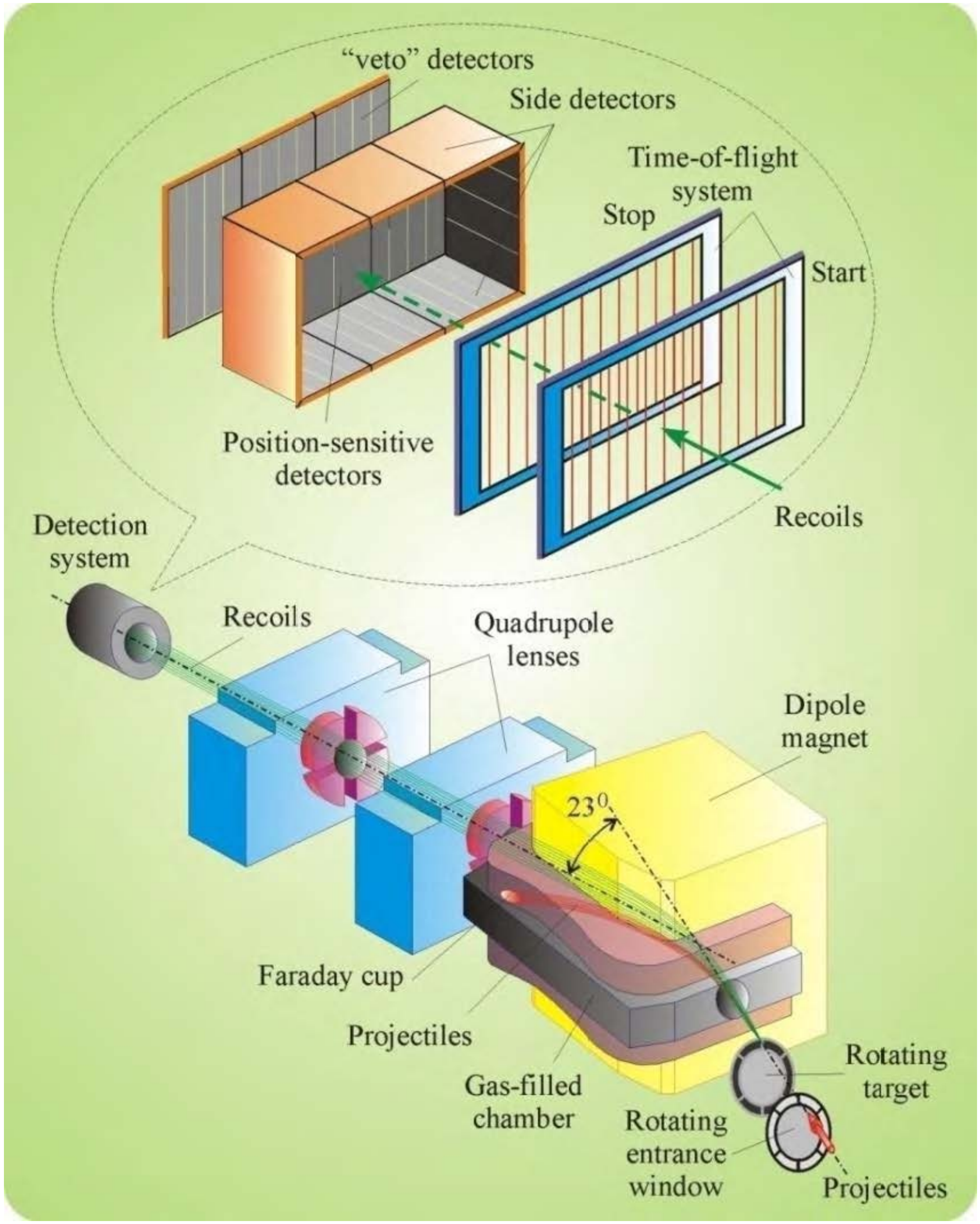
Berkelium -249 at hot cell

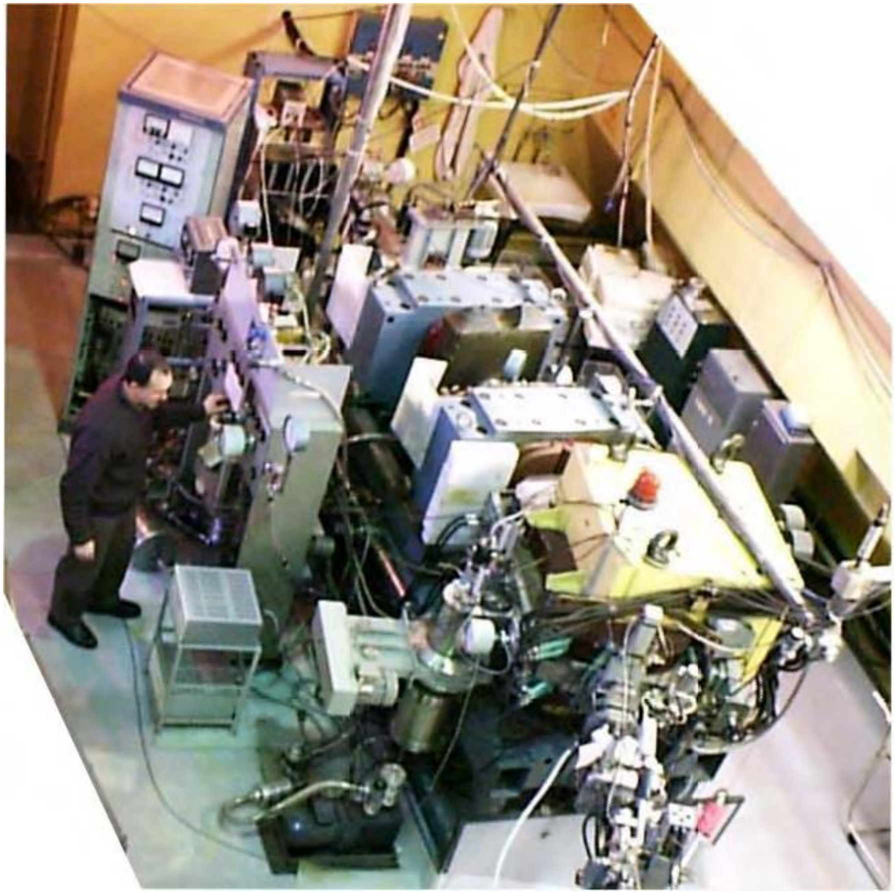
Feb. 5, 2012 ORNL, Oak Ridge, Tennessee, USA

^{249}Cf target



Dubna gas filled recoil separator (DGFRS)

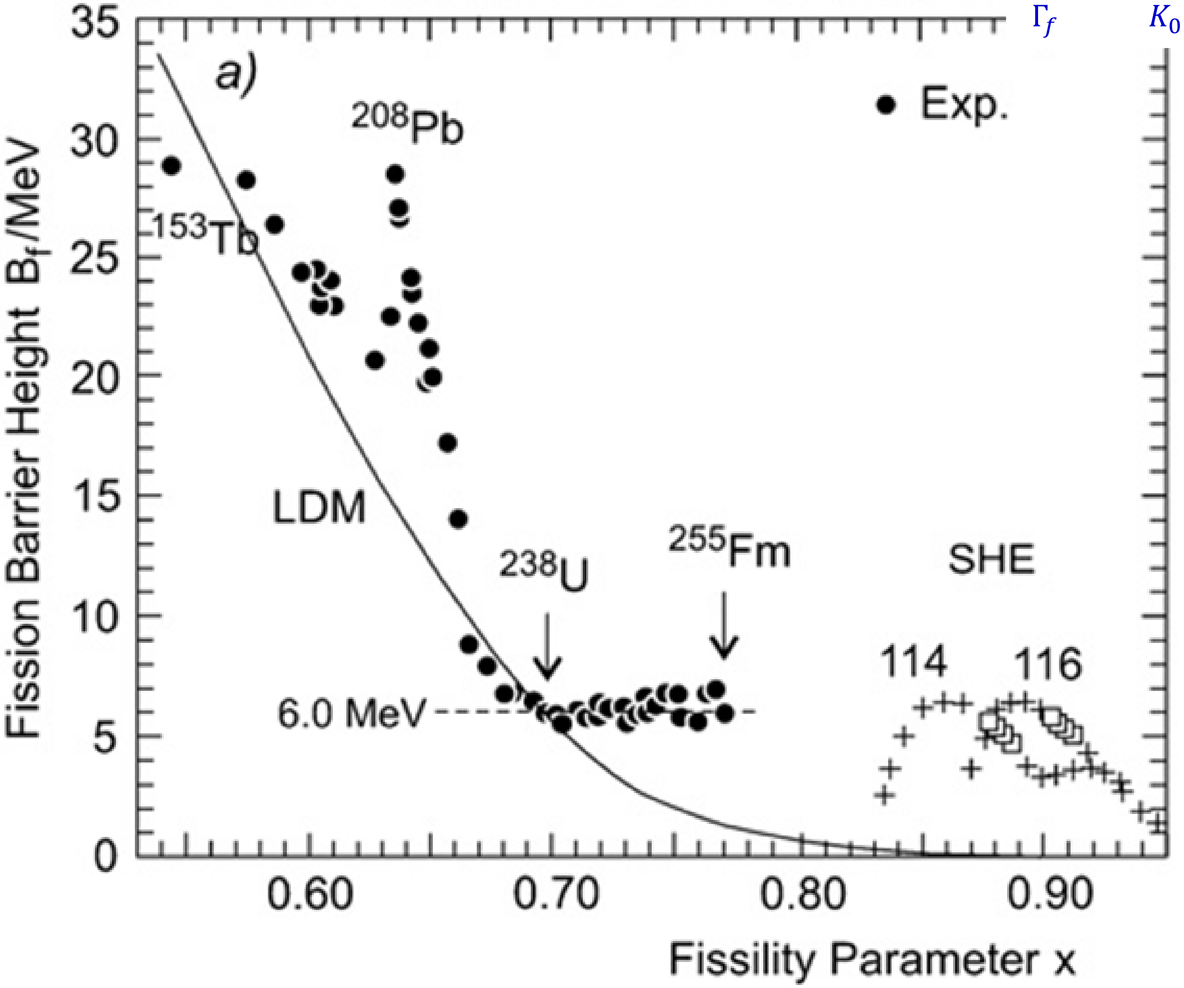




Cross sections – fission barriers

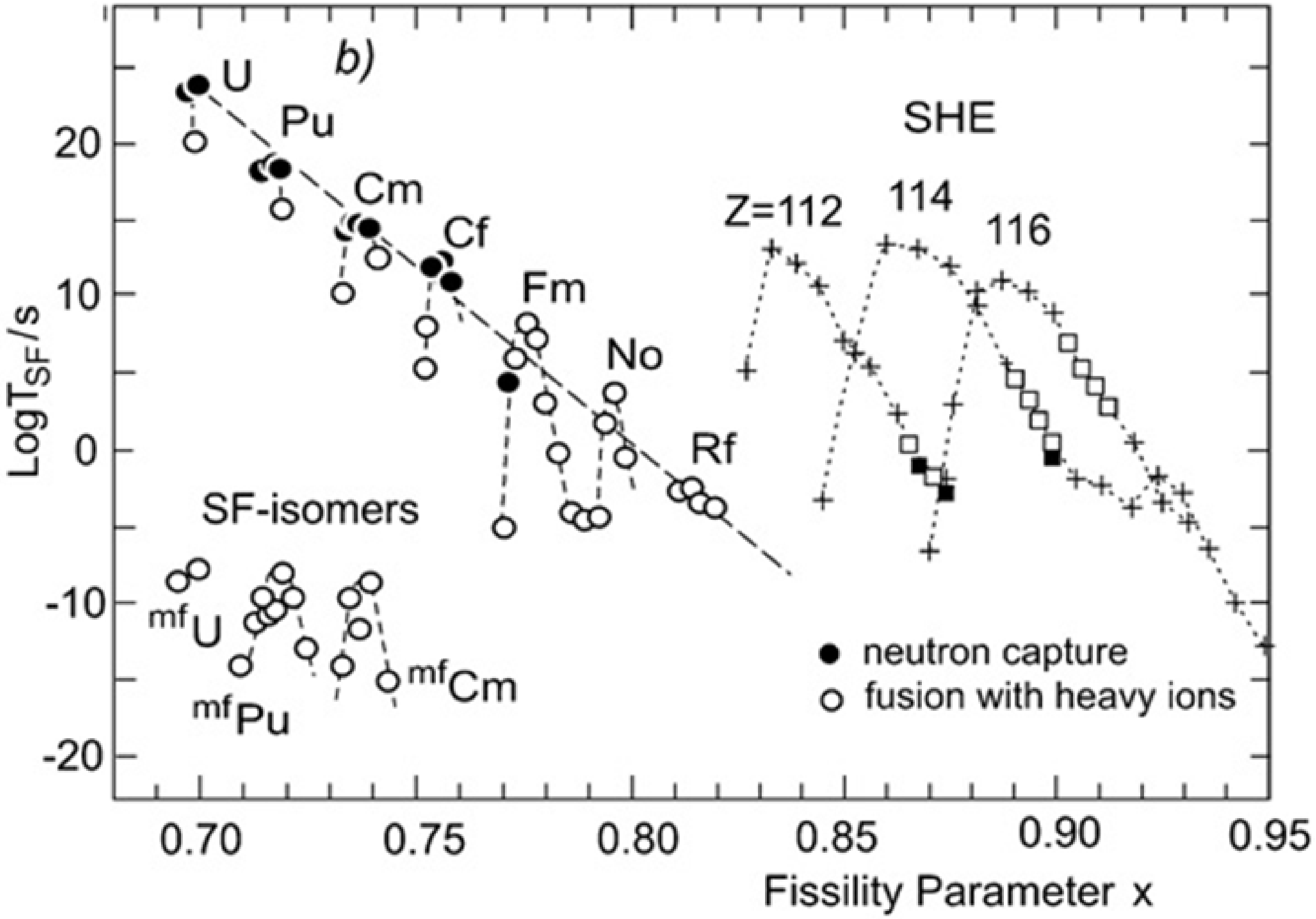
$$\sigma_{ER} = \sigma_{capture} \cdot P_{CN} \cdot P_{survival}$$

$$\frac{\Gamma_n}{\Gamma_f} \sim \frac{2 \cdot T \cdot A_{CN}^{2/3}}{K_0} \cdot exp\left[\left(B_f - B_n\right)/T\right]$$



Y. Oganessian; J. Phys. G: Nucl. Part. Phys. 34 (2007) E165

Spontaneous fission half-lives of actinides



Y. Oganessian; J. Phys. G: Nucl. Part. Phys. 34 (2007) E165

Chart of nuclides: the domain of heavy and super heavy elements

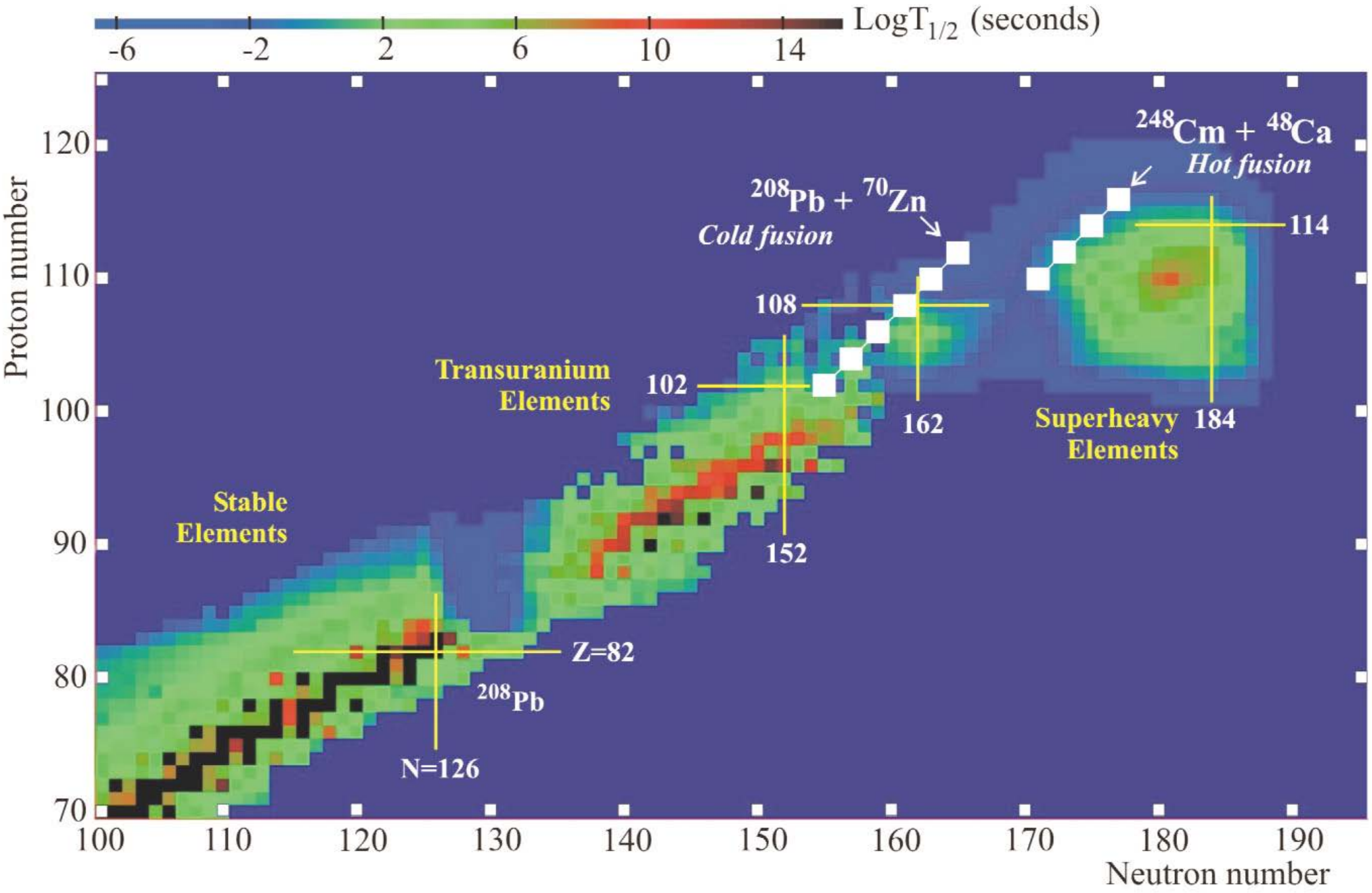
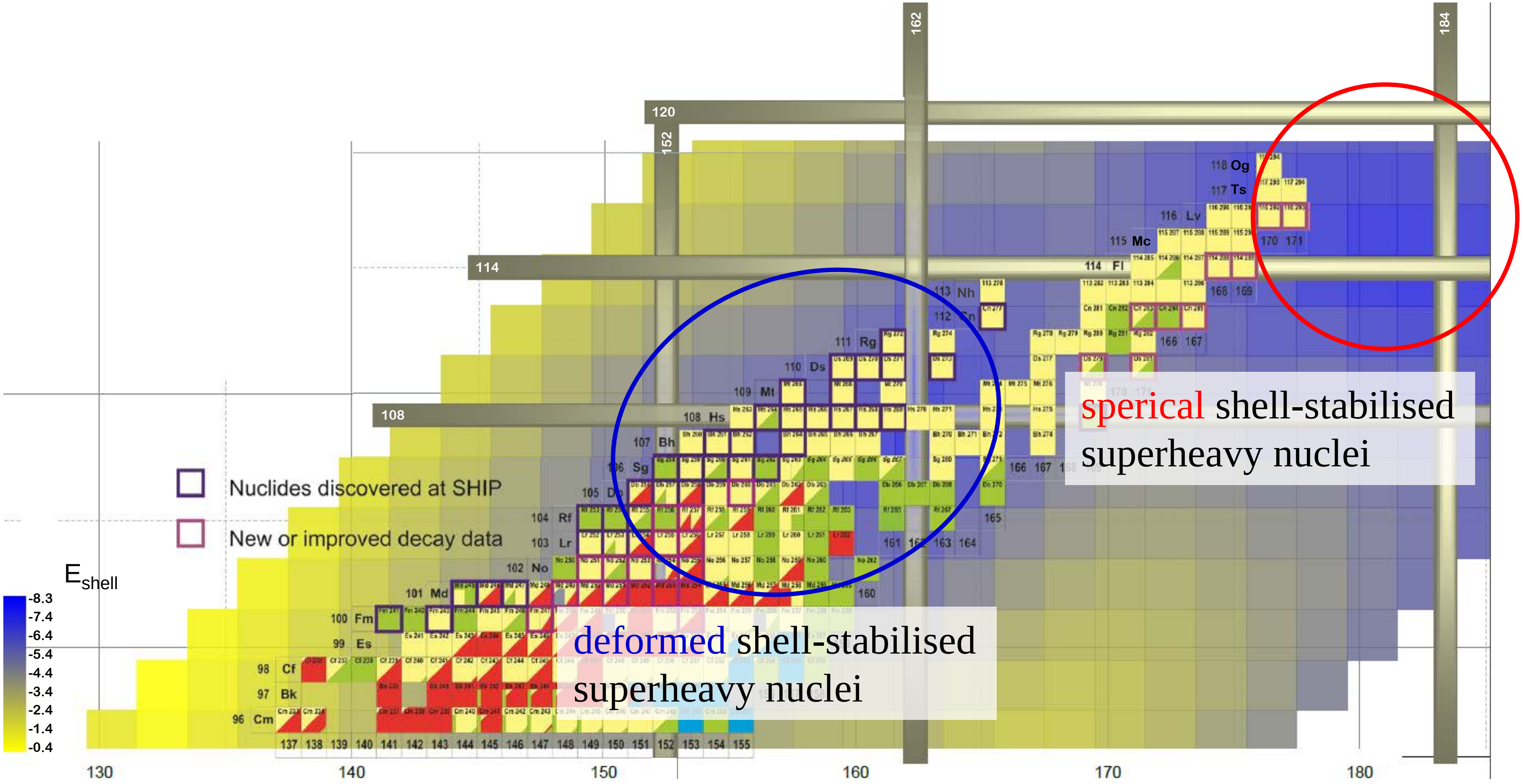


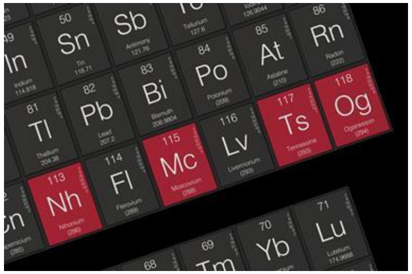
Chart of nuclides: the domain of heavy and superheavy elements



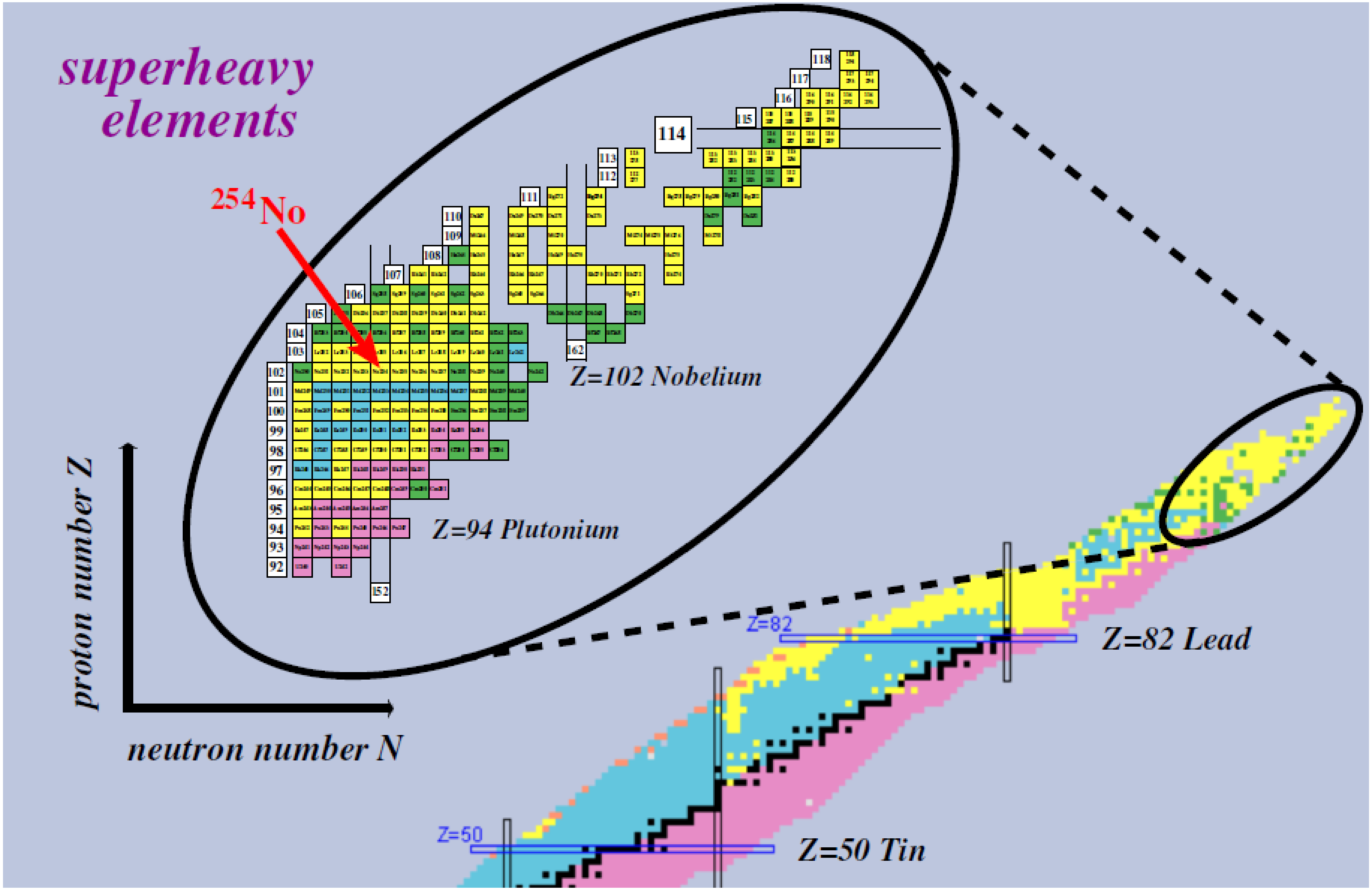
sperical shell-stabilised superheavy nuclei

deformed shell-stabilised superheavy nuclei

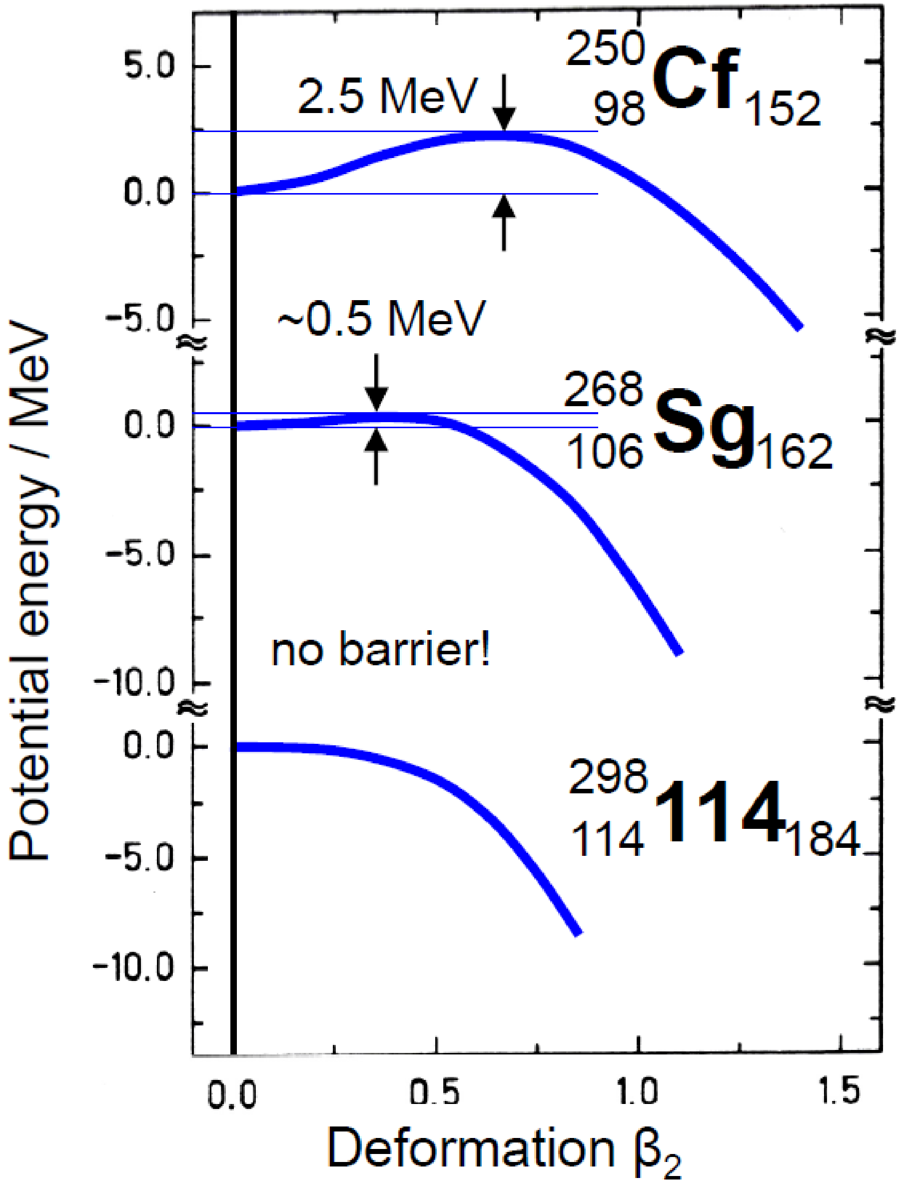
Calc.: A. Sobiczewski



What is the structure of SHE?



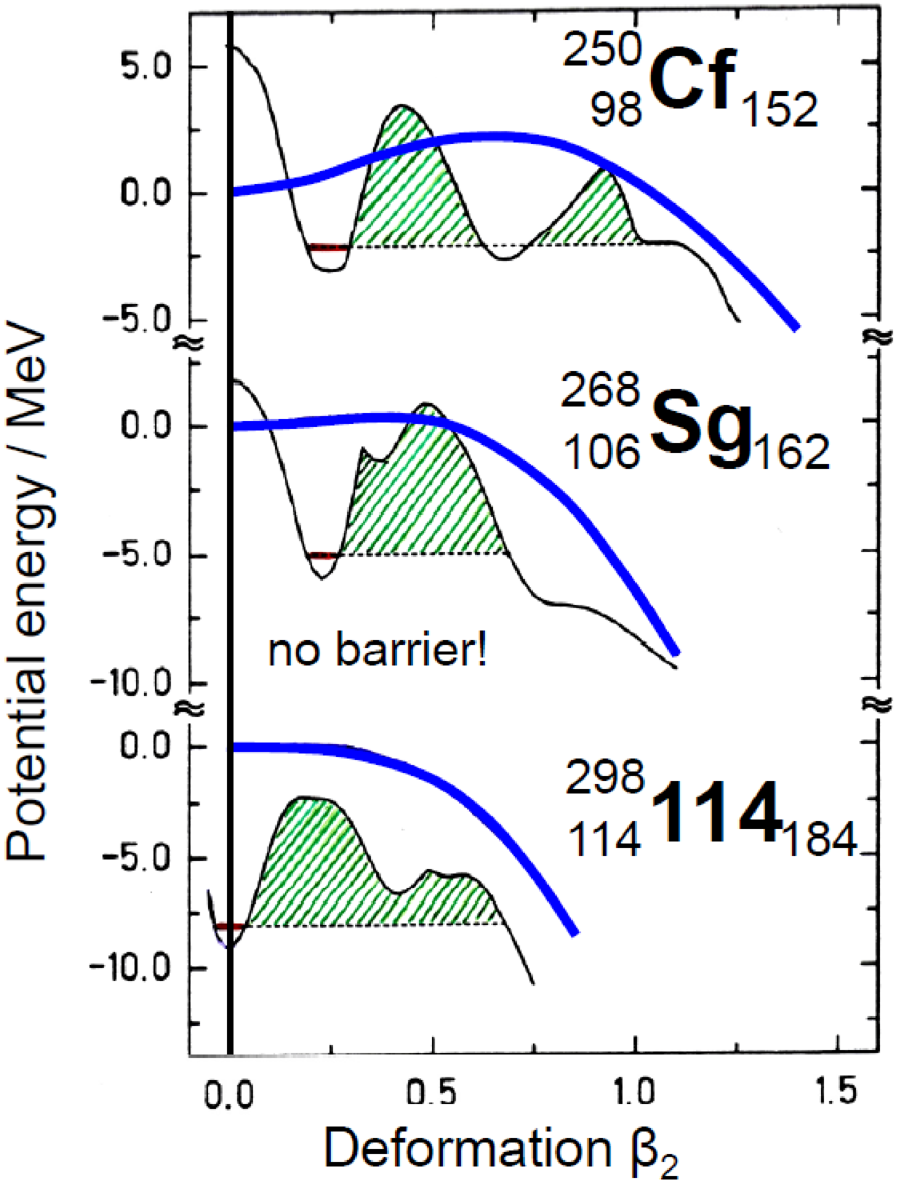
Influence of shell effects on fission barrier



▬ macroscopic barrier:

disappears at Z ~ 104!

Influence of shell effects on fission barrier



▬▬ macroscopic barrier:

disappears at Z ~ 104!

—— with shell structure:

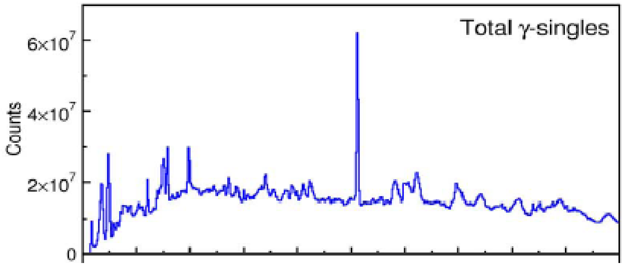
spherical ↔ deformed ground state

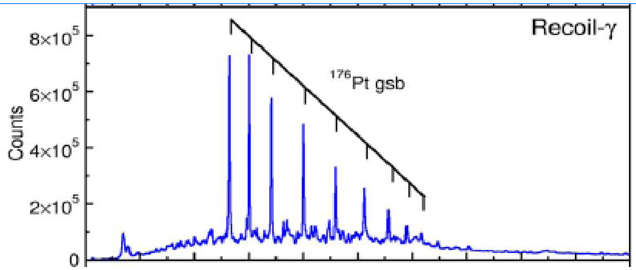
fission barrier is also > 0 for Z ≥ 104

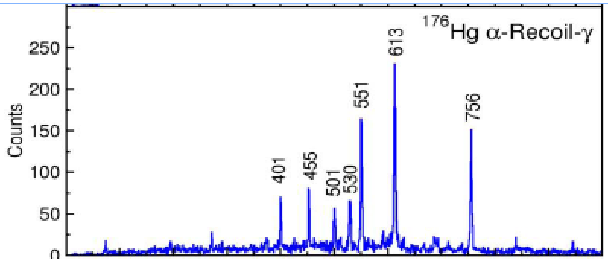
some fission barriers have complicated shapes, multi-humped structure

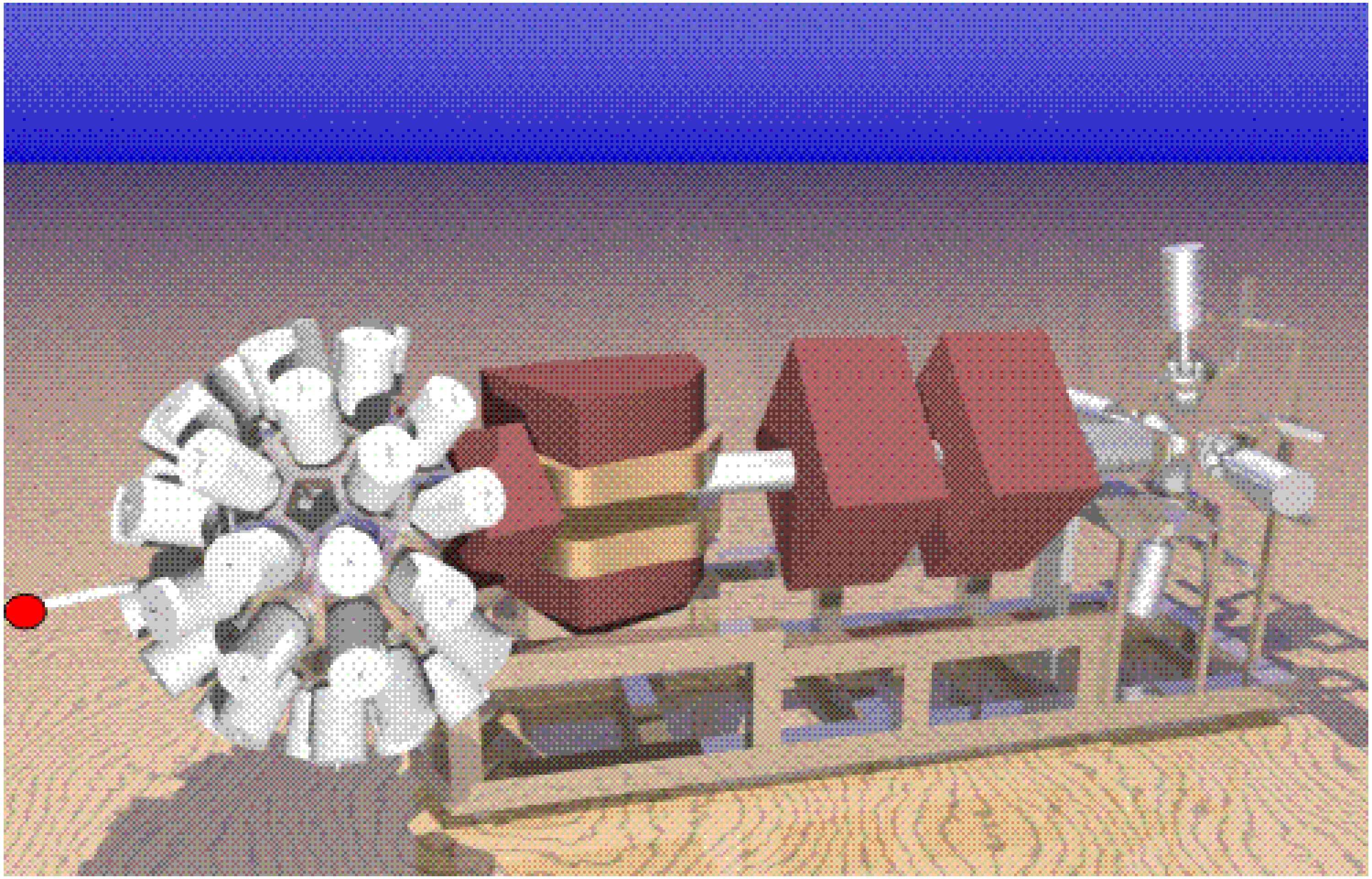
elements exist only due to shell effects:

superheavy elements



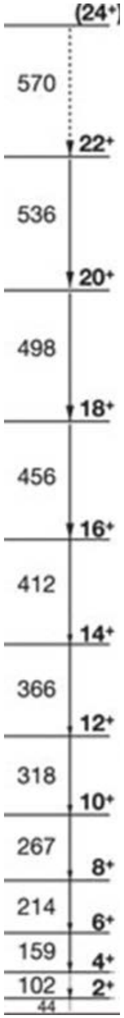


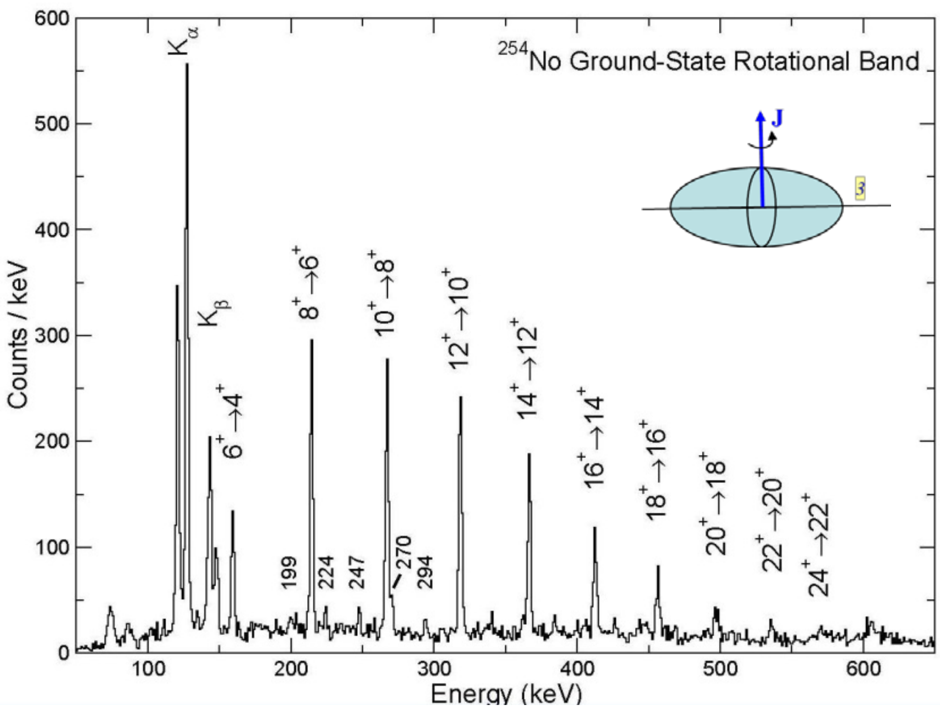




tagging instrumentation at JYFL (Jyväskylä) JUROGAM, RITU, GREAT

Spinning the heaviest elements

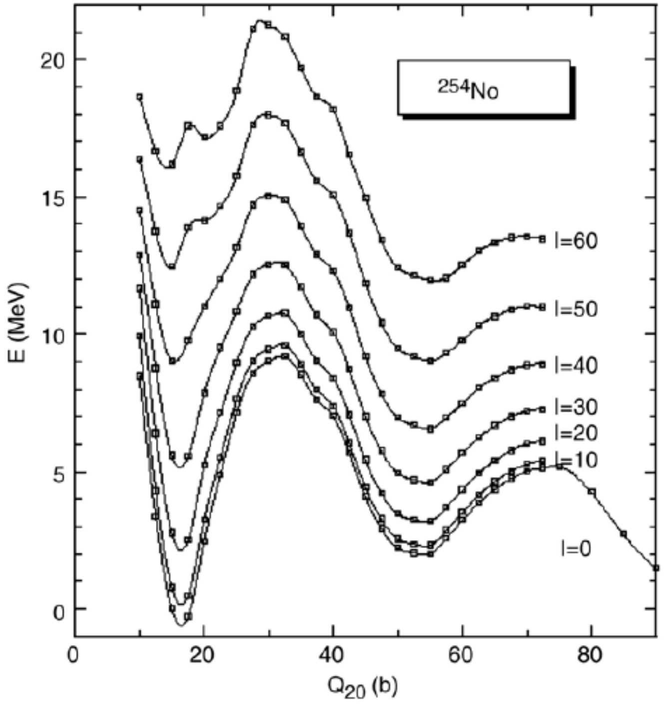




rotational energy: $E_I = \frac{\hbar^2}{2\Im} \cdot I(I+1)$

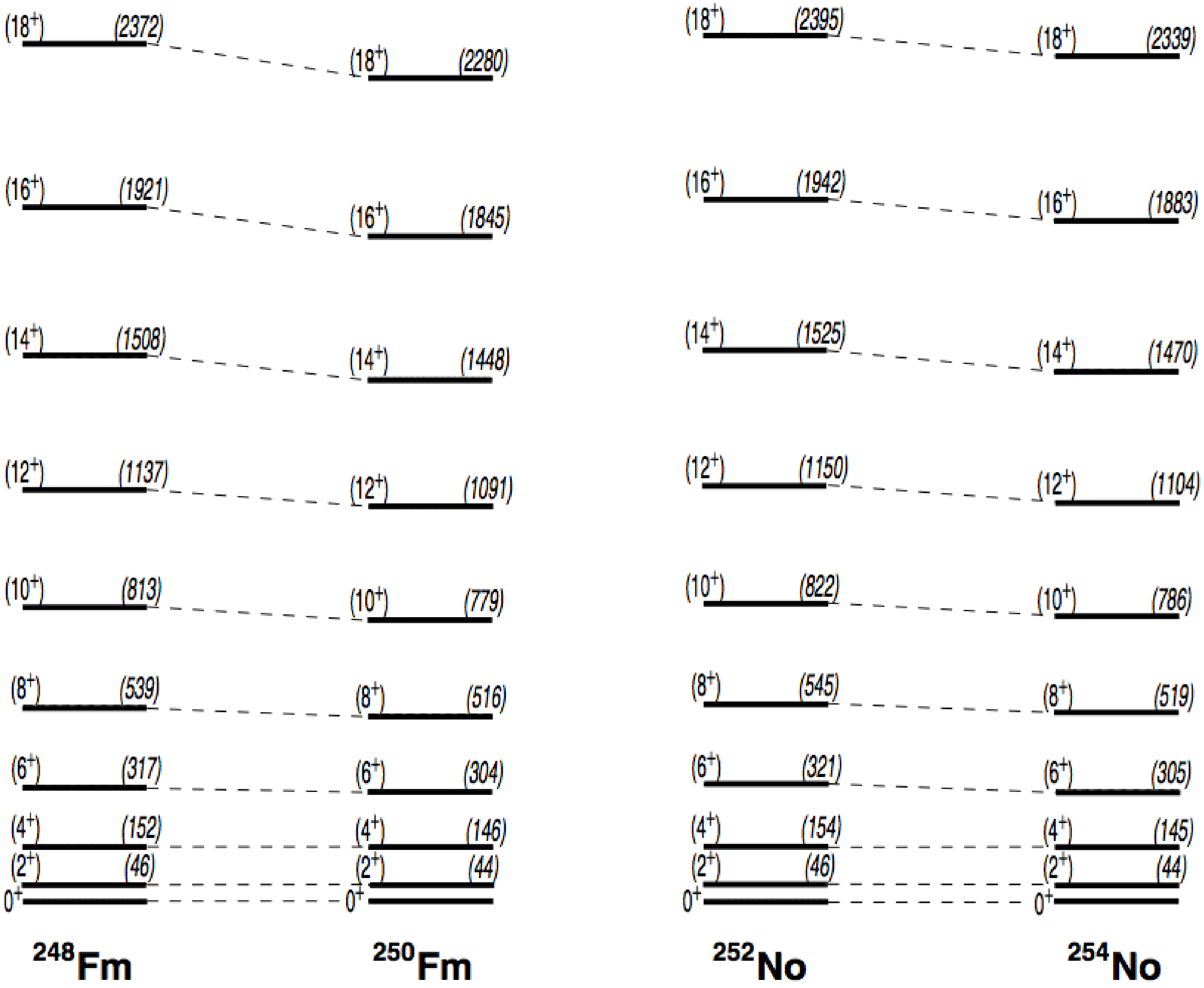
γ-ray energy: $E_I - E_{I-2} = \frac{\hbar^2}{2\Im} \cdot (4I-2)$

dependence of the fission barrier on spin I



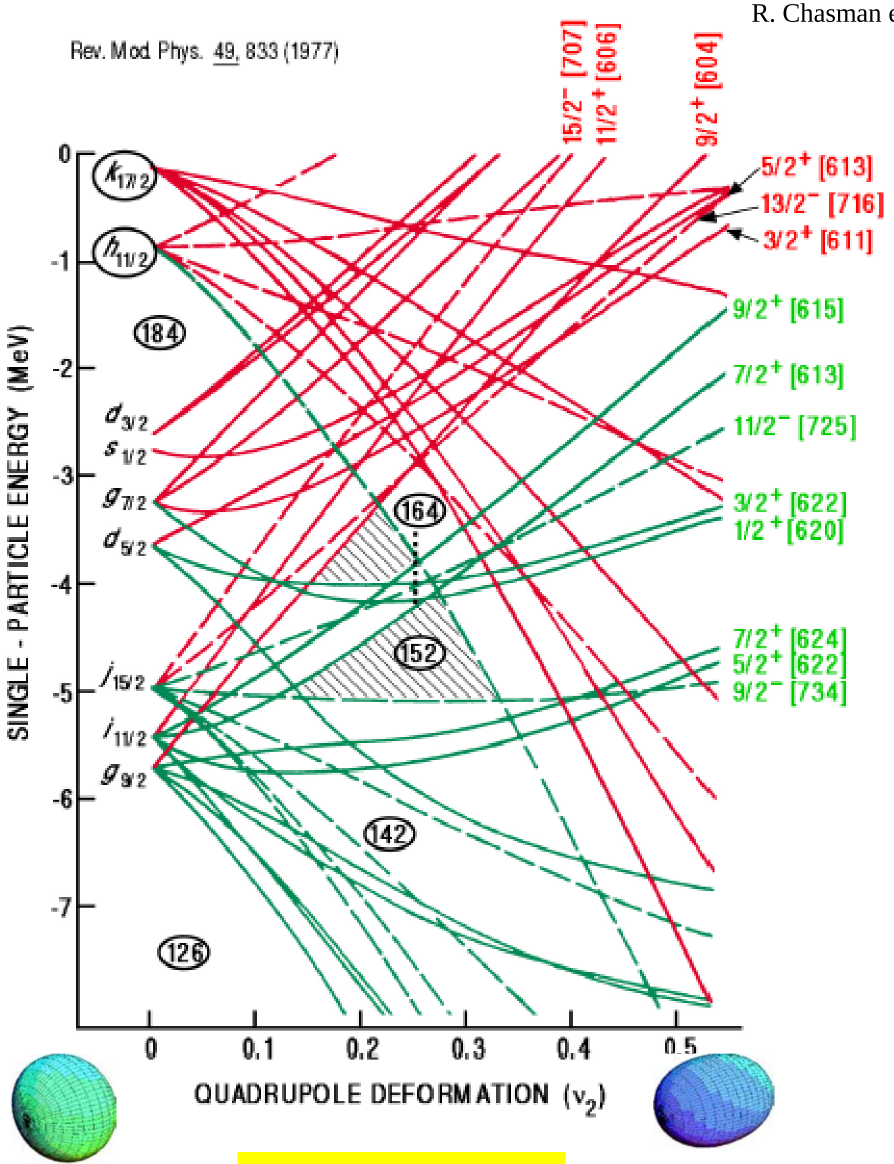
S. Eeckhaupt et al.; EPJA 26 (2005), 227

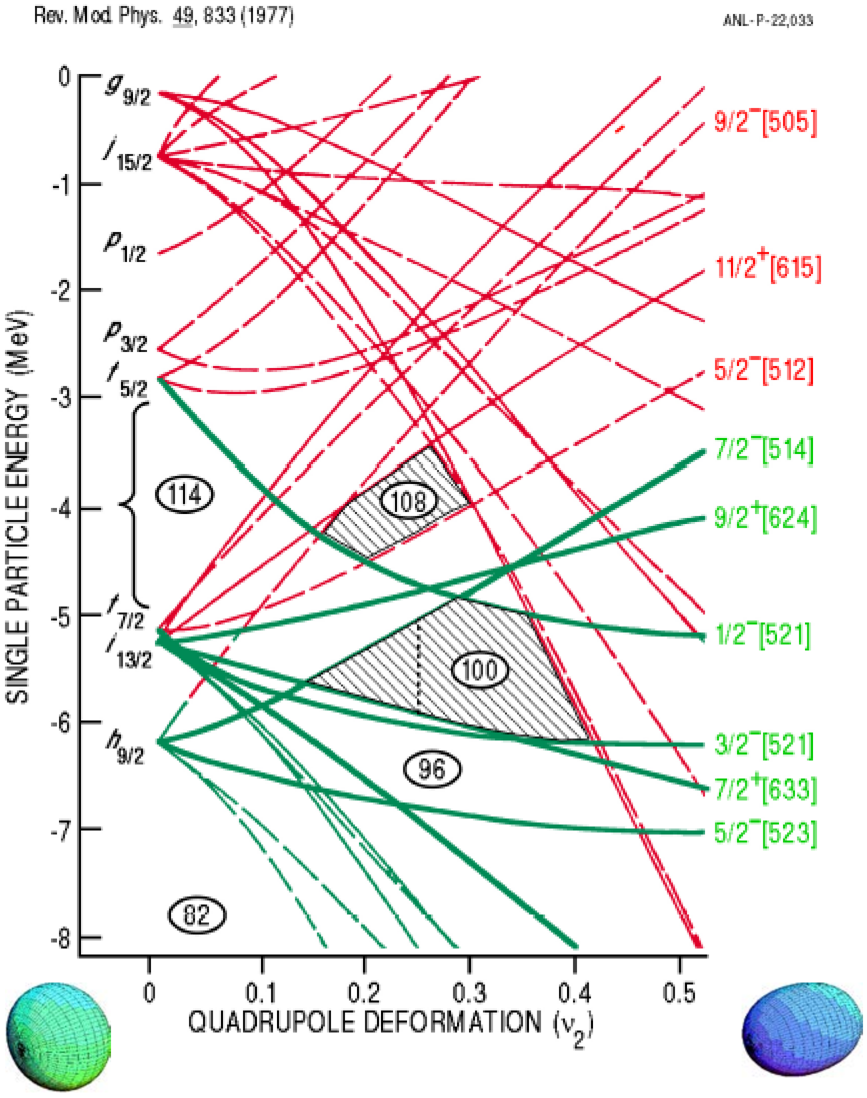
Rotational Bands



Single particle orbitals

R. Chasman et al. Rev. Mod. Phys. 49 (1977), 833



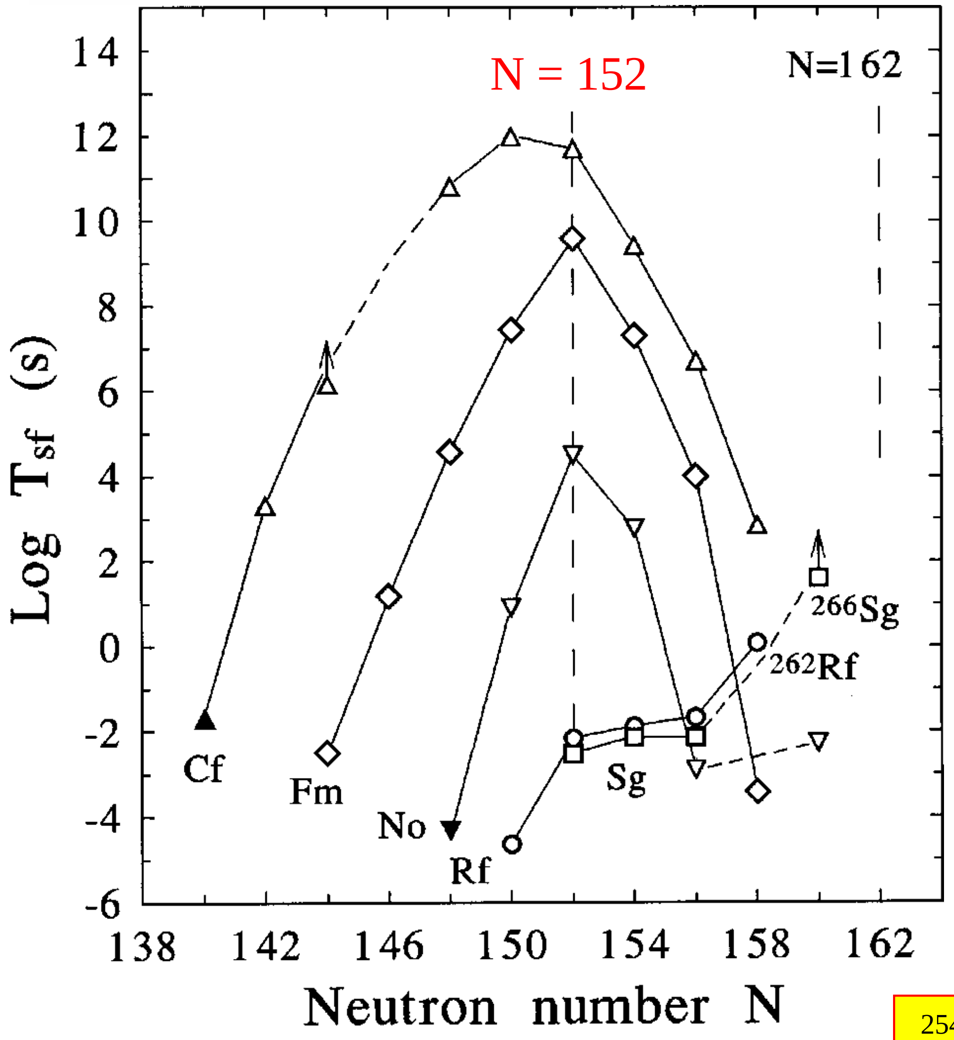


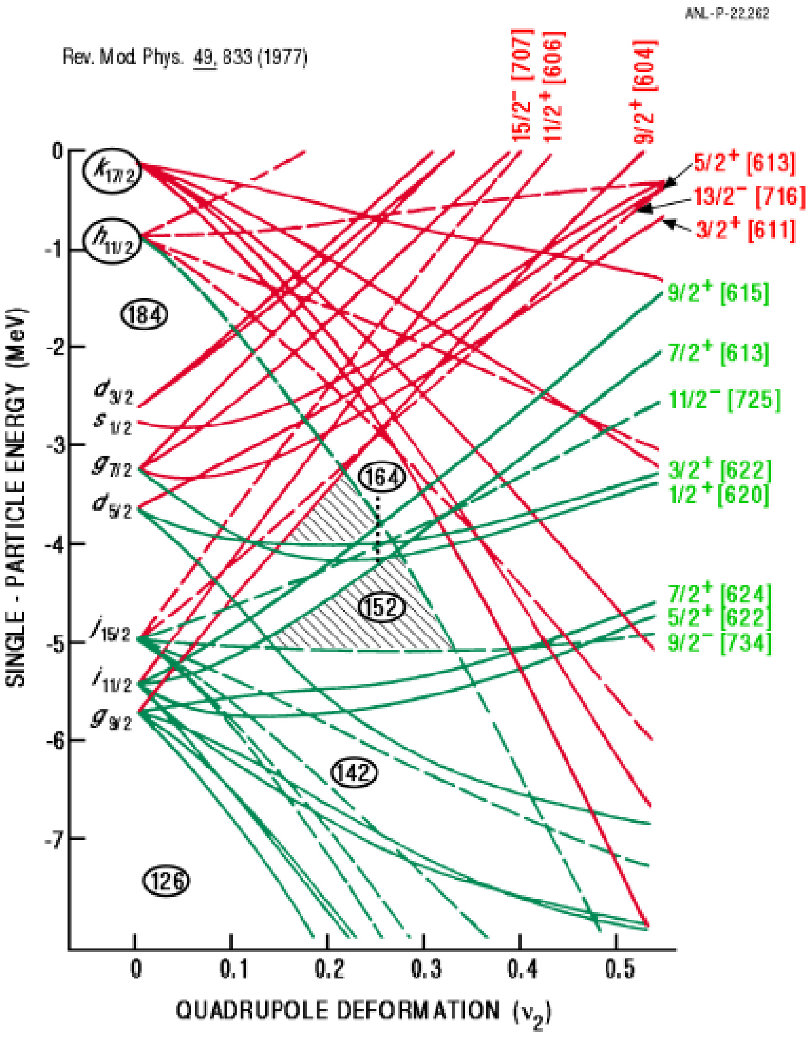
oblate

$^{254}_{102}No_{152}$ β_2~0.28

prolate

Stability of heavy elements – Nilsson level energy gap



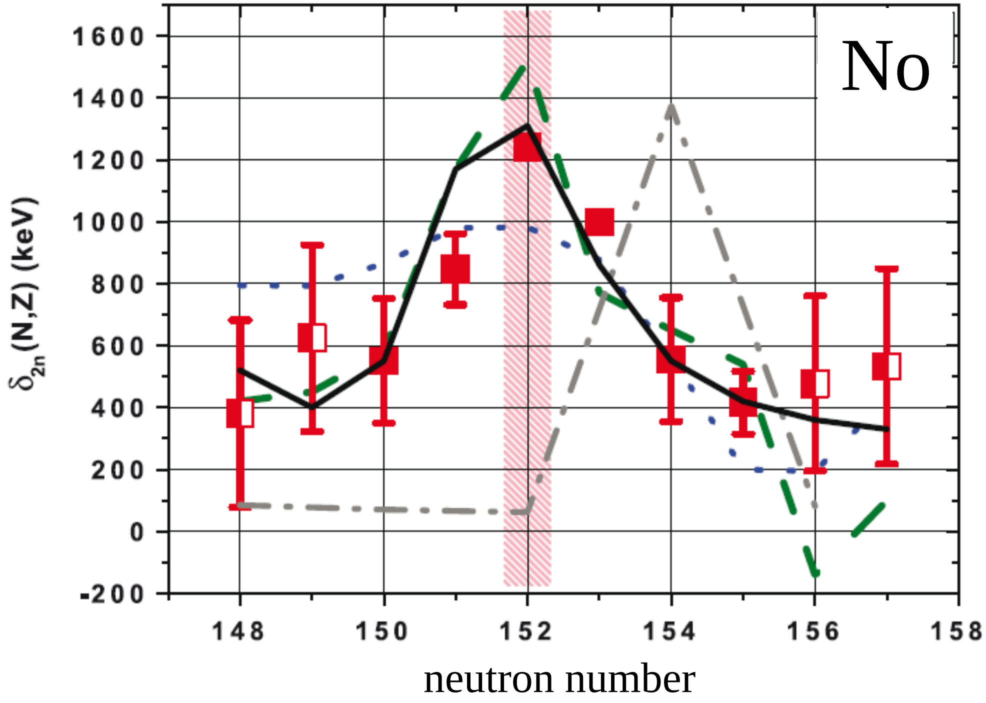


^{254}No (Z=102), ^{252}Fm (Z=100) and ^{250}Cf (Z=98)
with N=152
seem to be more stable than their neighbors

Shiptrap: Probing the strength of shell effects

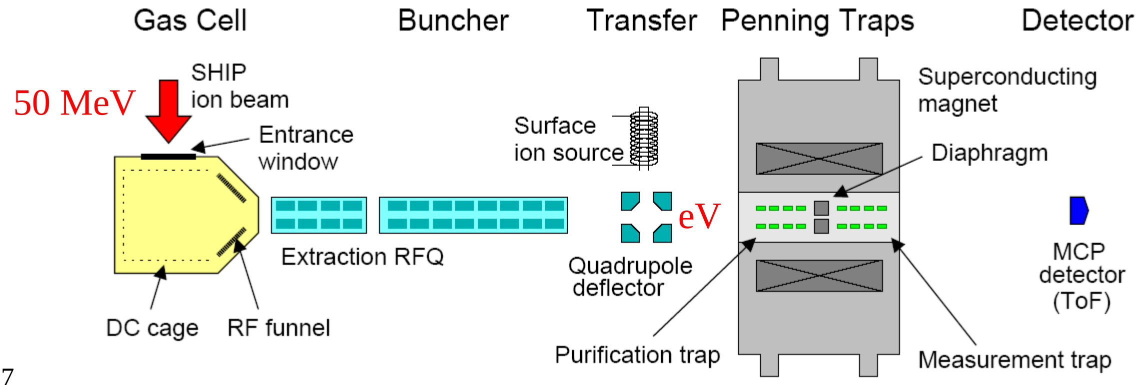
$$\delta_{2n}(N,Z) = 2 \cdot B(N,Z) - B(N-2,Z) - B(N+2,Z)$$

$$\delta_{2n}(N,Z) = S_{2n}(N,Z) - S_{2n}(N+2,Z)$$



Muntian (mic-mac), Z=114 N=184
Möller FRDM, Z=114 N=184
TW-99, Z=120 N=172
SkM* Z=126 N=184

^{206}Pb(^{48}Ca,2n)^{252}No
^{207}Pb(^{48}Ca,2n)^{253}No
^{208}Pb(^{48}Ca,2n)^{254}No
^{208}Pb(^{48}Ca,1n)^{255}No

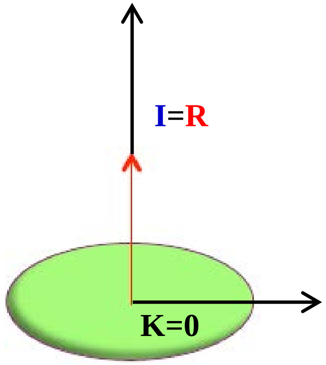


E. Minaya Ramirez et al.; Science 337 (2012), 1207

Isomeric states in even-even nuclei

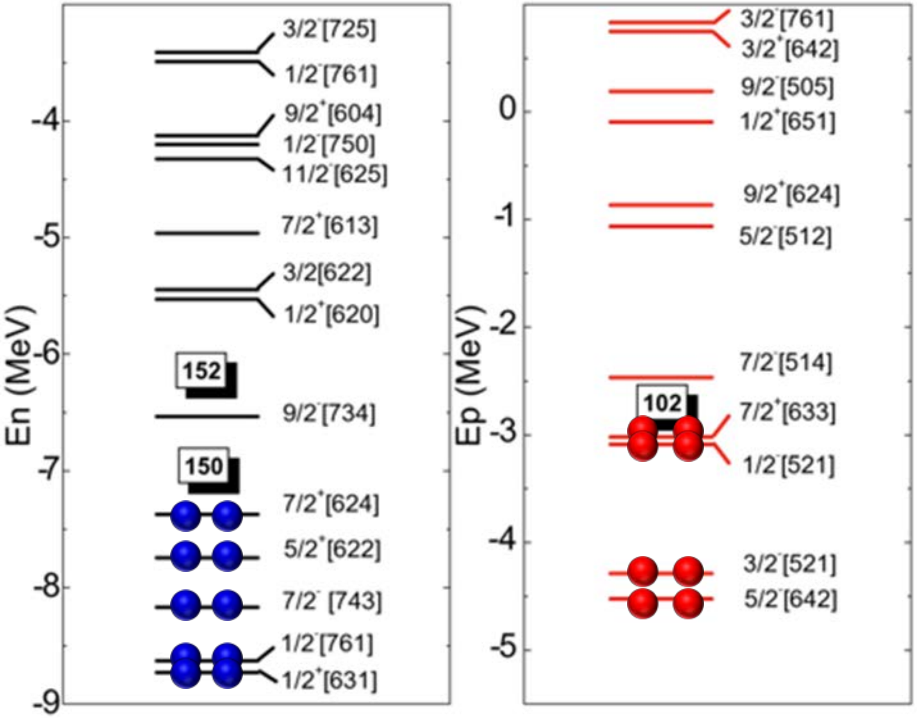
Why K-isomers occur?

- deformed nuclei
- breaking of particle pair at Fermi surface
- selection rule for electromagnetic transitions
  $\lambda \geq \Delta K$ is not fulfilled
- excitation energy of quasi-particle:
  $E = \sqrt{(\varepsilon - \lambda)^2 + \Delta^2}$



What we can learn?

- information about Nilsson level energy gaps
- influence on stability of superheavy elements
- pairing interaction



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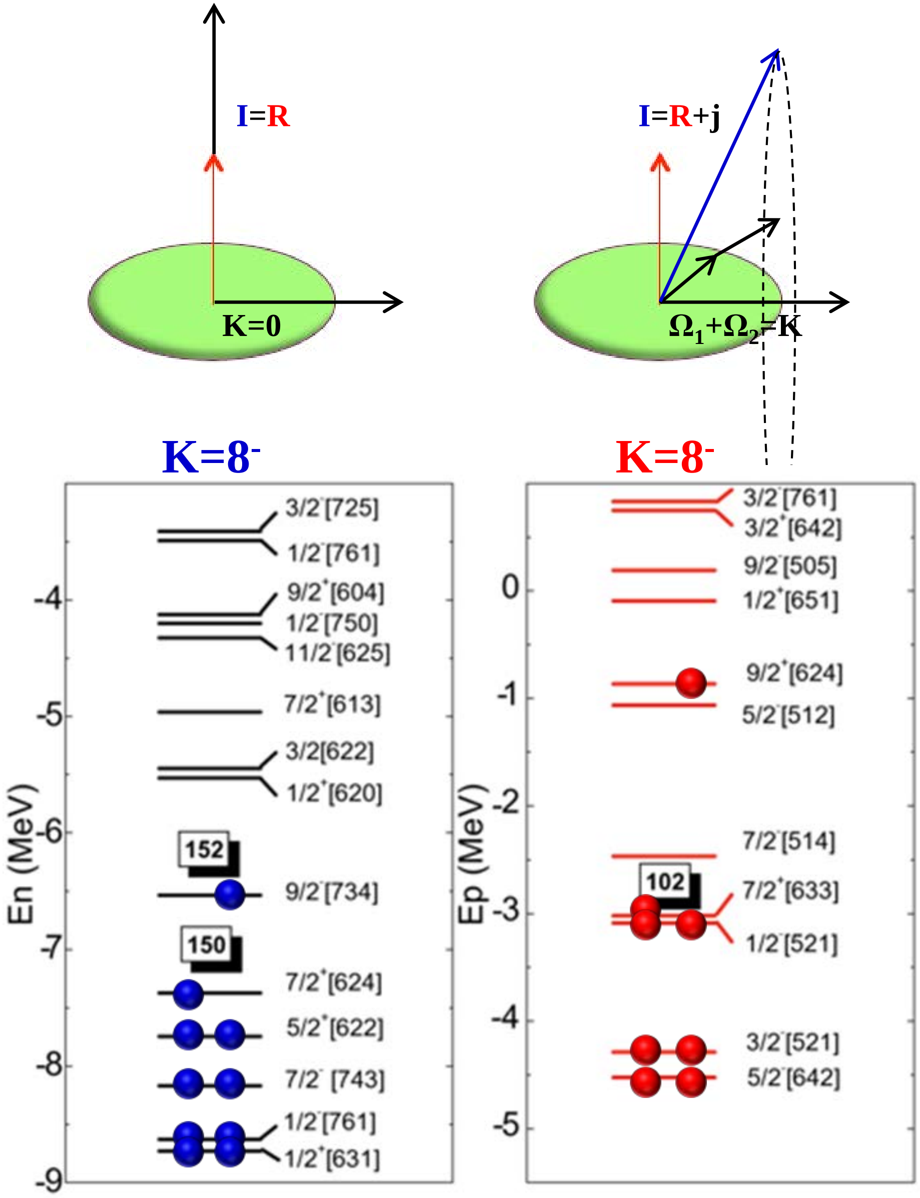
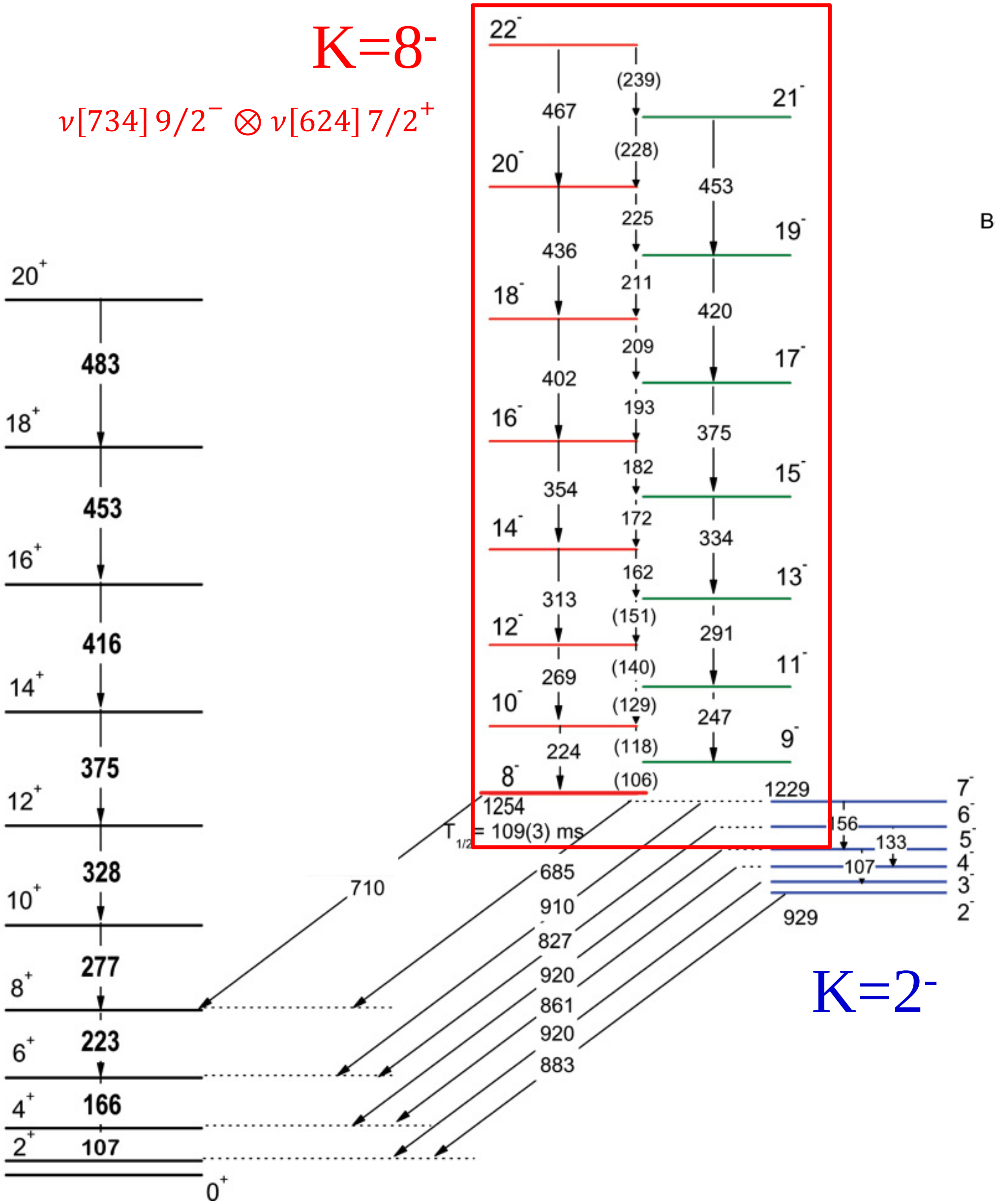


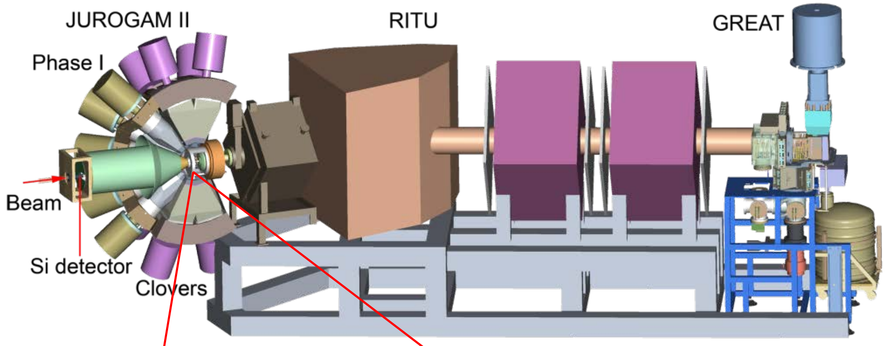
K-isomers in ^{252}No

K=8$^-$

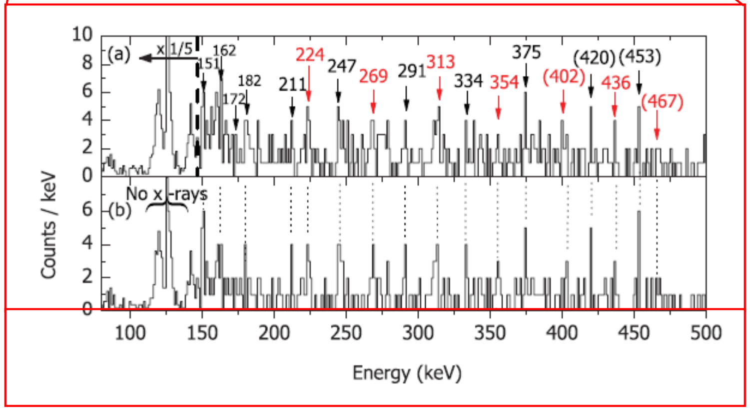
$\nu[734]\,9/2^- \otimes \nu[624]\,7/2^+$



K=2$^-$



K=8$^-$



B. Sulignano st al; Phys. Rev. C 86, 044318 (2012)

Neutron single particle states in ^{252}No

$7/2^+[624]$ Nilsson configuration $i_{11/2}$

$9/2^-[734]$ Nilsson configuration $j_{15/2}$

$$g(j) = \begin{cases} \dfrac{2\ell \cdot g_\ell + g_s}{2\ell + 1} & for\ j = \ell + 1/2 \\ \dfrac{2(\ell + 1) \cdot g_\ell - g_s}{2\ell + 1} & for\ j = \ell - 1/2 \end{cases}$$

proton: $g_\ell = 1$ $g_s = 5.59$
neutron: $g_\ell = 0$ $g_s = -3.83$

$g(i_{11/2}) = +0.295$ $g(j_{15/2}) = -0.255$

$$g(j_1 \times j_1; J) = \frac{1}{2} \cdot (g_1 + g_2) + \frac{j_1(j_1+1) - j_2(j_2+1)}{2J(J+1)} \cdot (g_1 - g_2)$$

$g(i_{11/2}\ x\ j_{15/2};\ 8) = 0.09$

$$B(M1; I \to I-1) = \frac{3}{4\pi}[g_K - g_R]^2 \cdot K^2 |\langle IK10|(I-1)K\rangle|^2$$

$$B(E2; I \to I-2) = \frac{5}{16\pi} Q_0^2 |\langle IK20|(I-2)K\rangle|^2$$

$g_R = Z/A = 0.40$

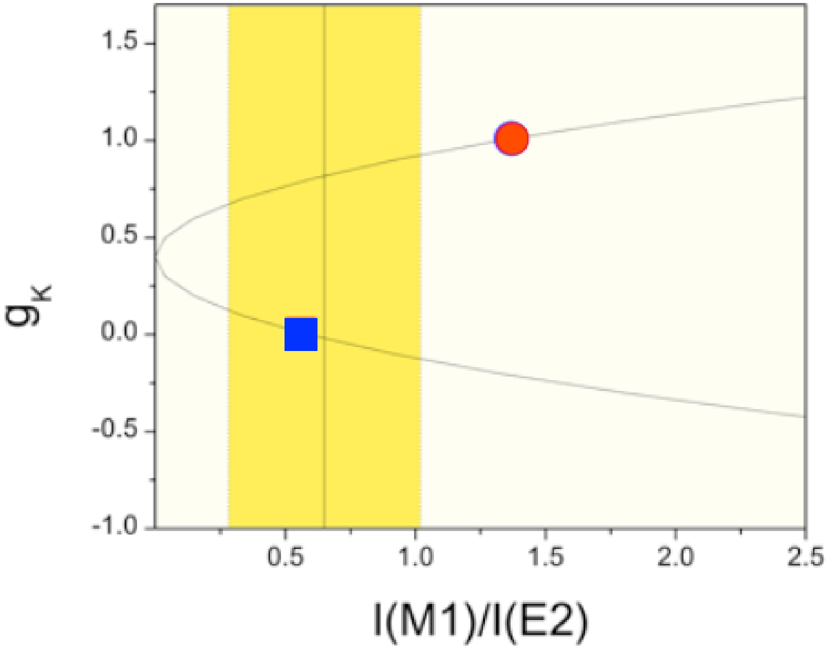
g_k νν(th.)
0.01

g_k ππ(th.)
1.001

Branching ratio

$$R_{exp} = I_\gamma(J \to J-1) / I_\gamma(J \to J-2)$$

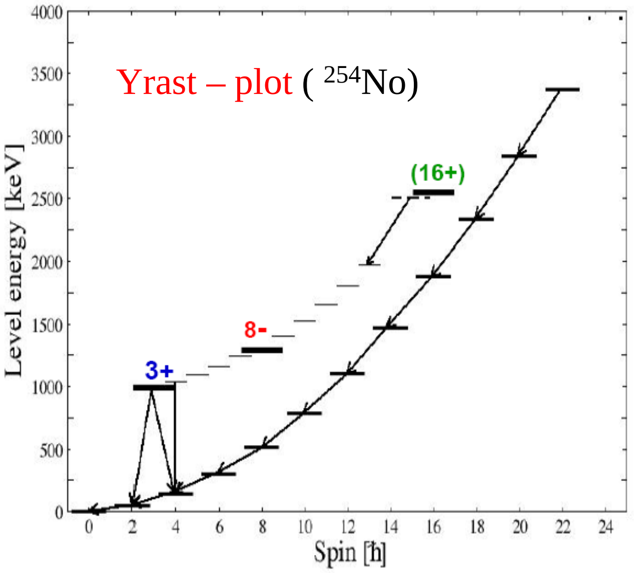
$$R_{th} = \frac{T(M1)}{T(E2)} = \frac{1.76[E_\gamma(M1)]^3 B(M1) s^{-1}}{1.22[E_\gamma(E2)]^5 B(E2) s^{-1}}; \propto \frac{(g_K - g_R)^2}{Q_0^2}$$



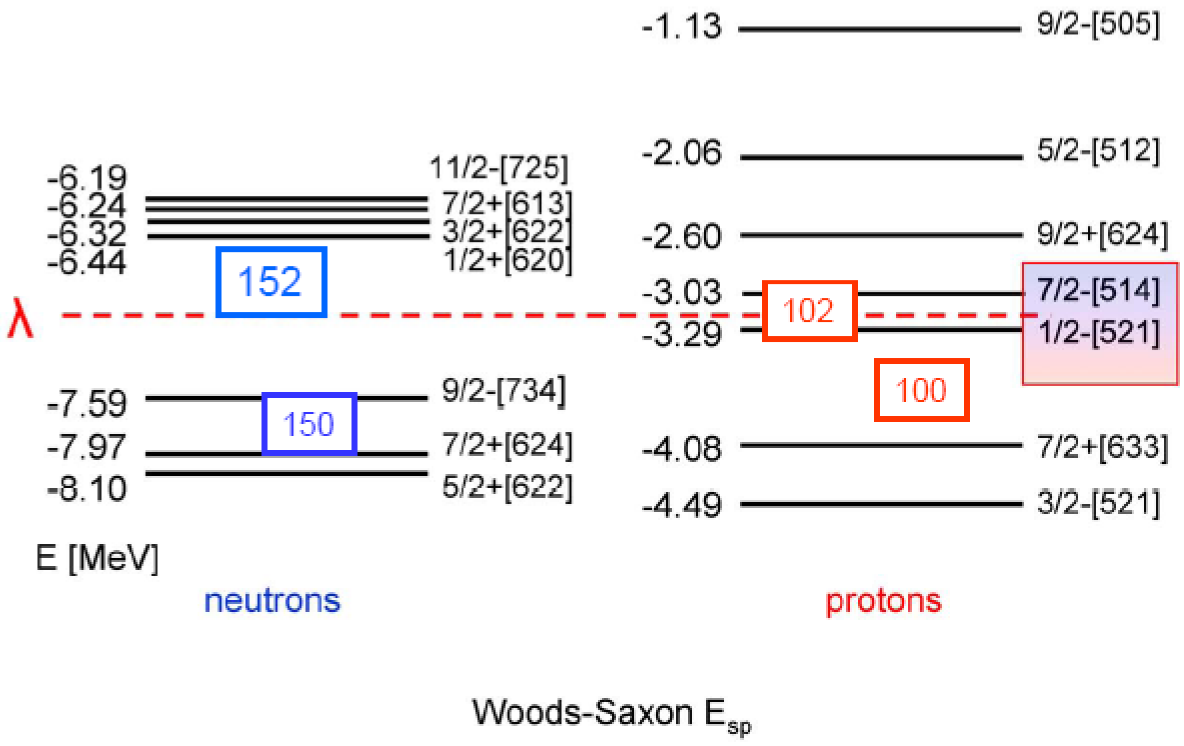
The isomer K=8$^-$ is based on 2-qn excitation

K-isomeric states

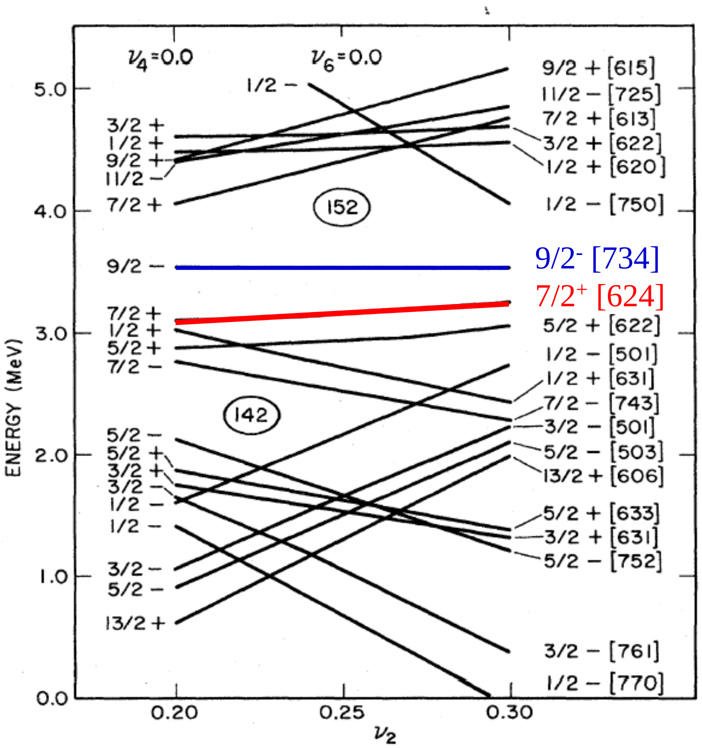
Exp. results: excitation of isomeric states



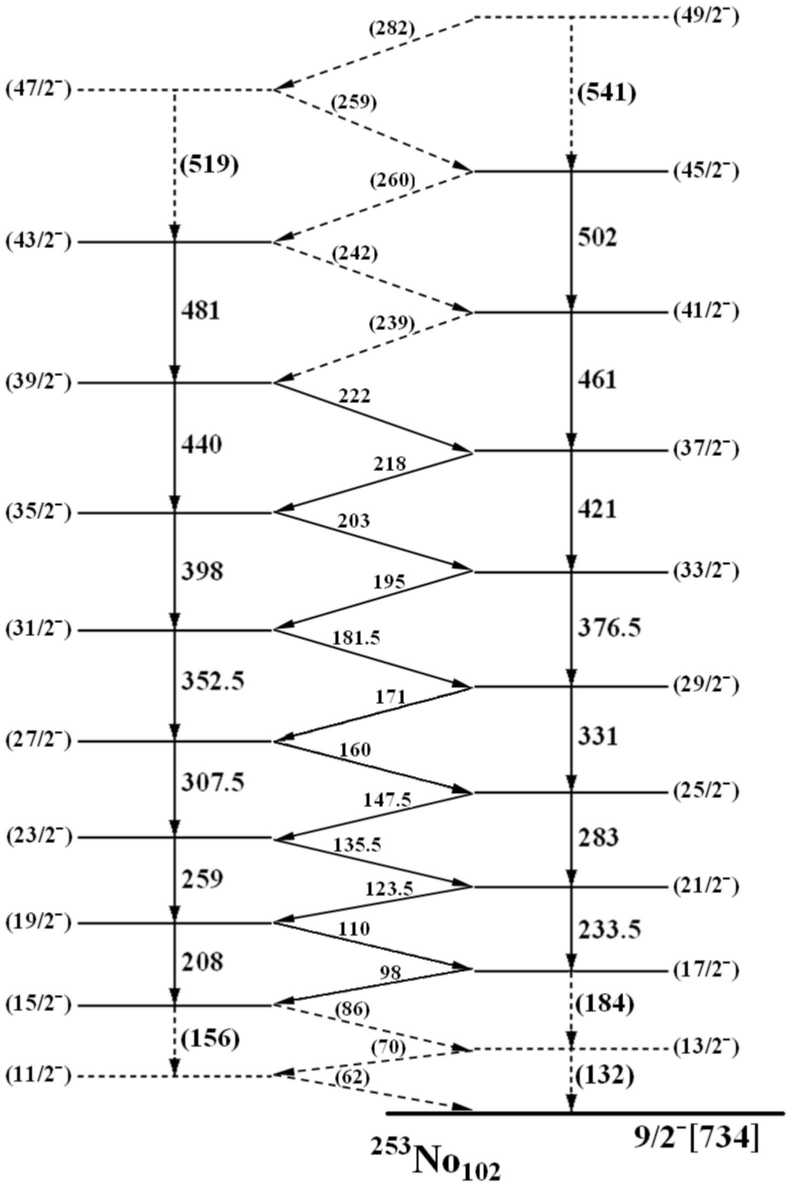
^{254}No with Z=102 and N=152 – protons will be excited easily
^{250}Fm with Z=100 and N=150 – neutrons will be excited easily



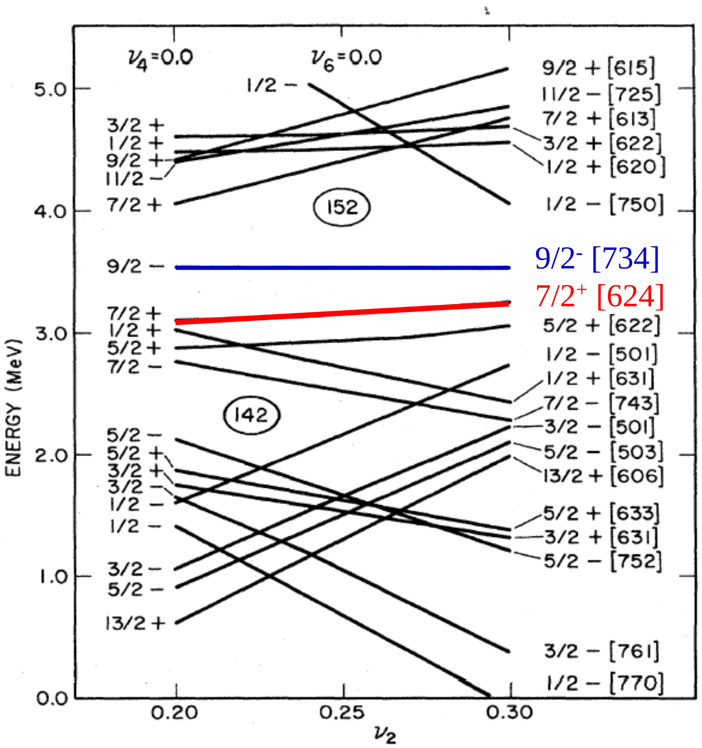
Level scheme of ^{253}No (151 neutrons)



R. Chasman et al.; Rev. Mod. Phys. 49 (1977) 833

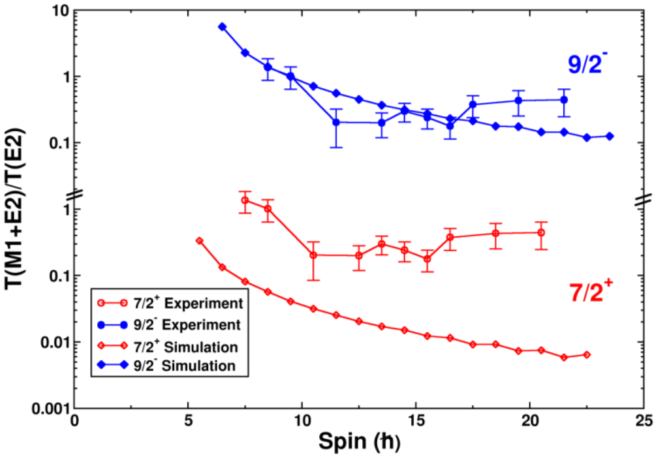


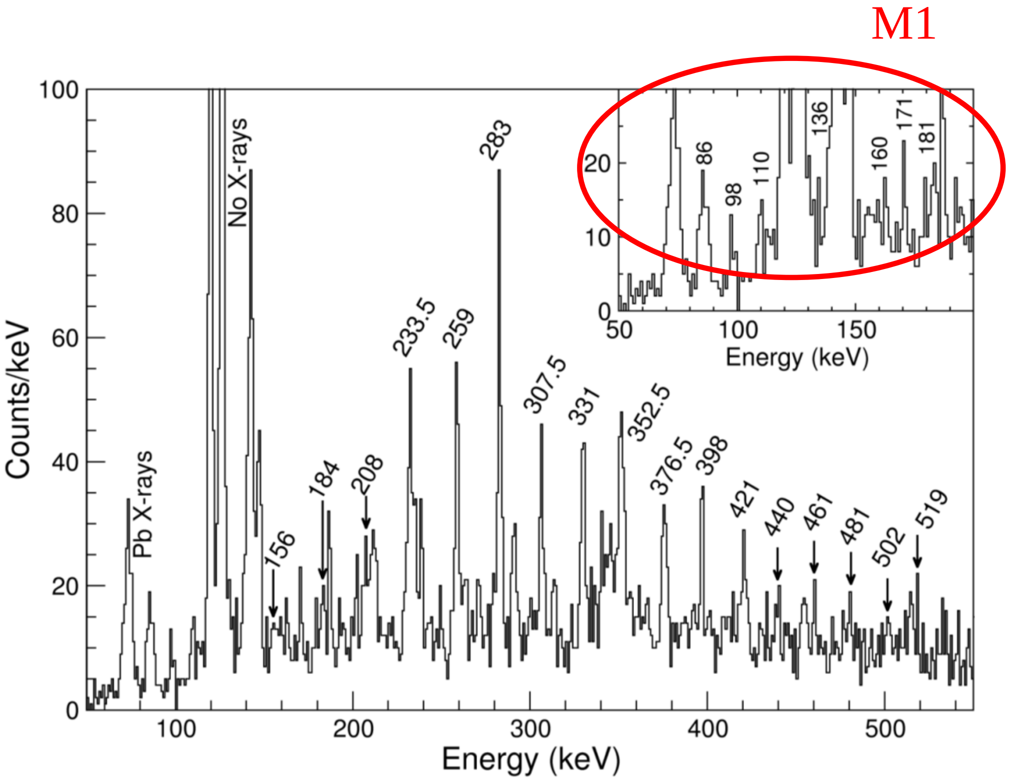
Level scheme of ^{253}No (151 neutrons)



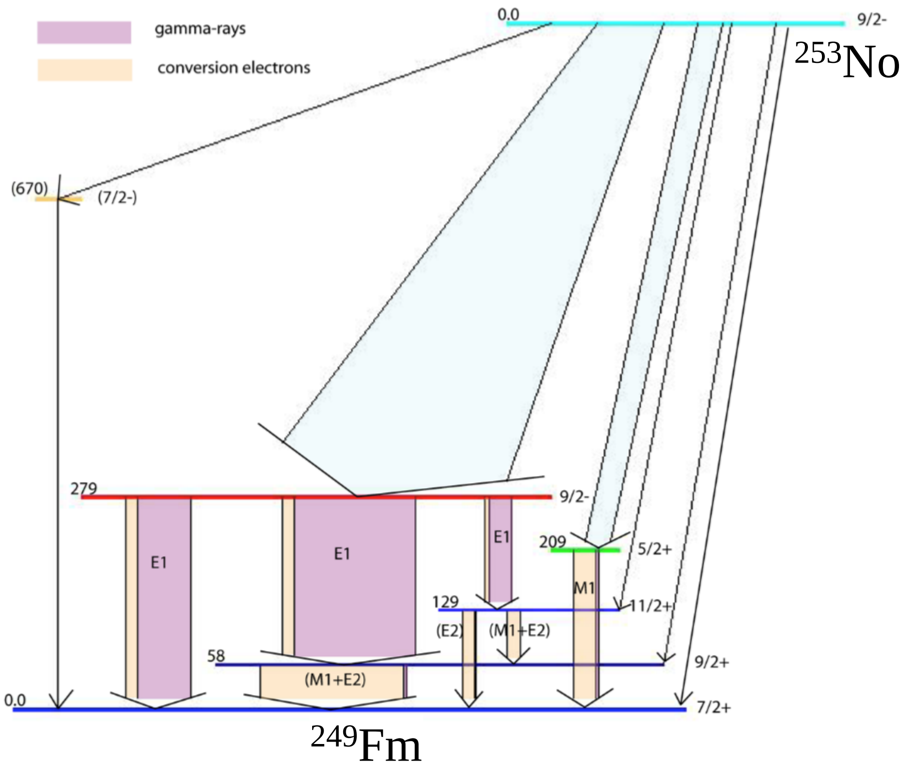
- ground state 9/2$^-$
- excited state 7/2$^+$

- rotational band observed at Gammasphere & JUROGAM





Alpha decay of ^{253}No

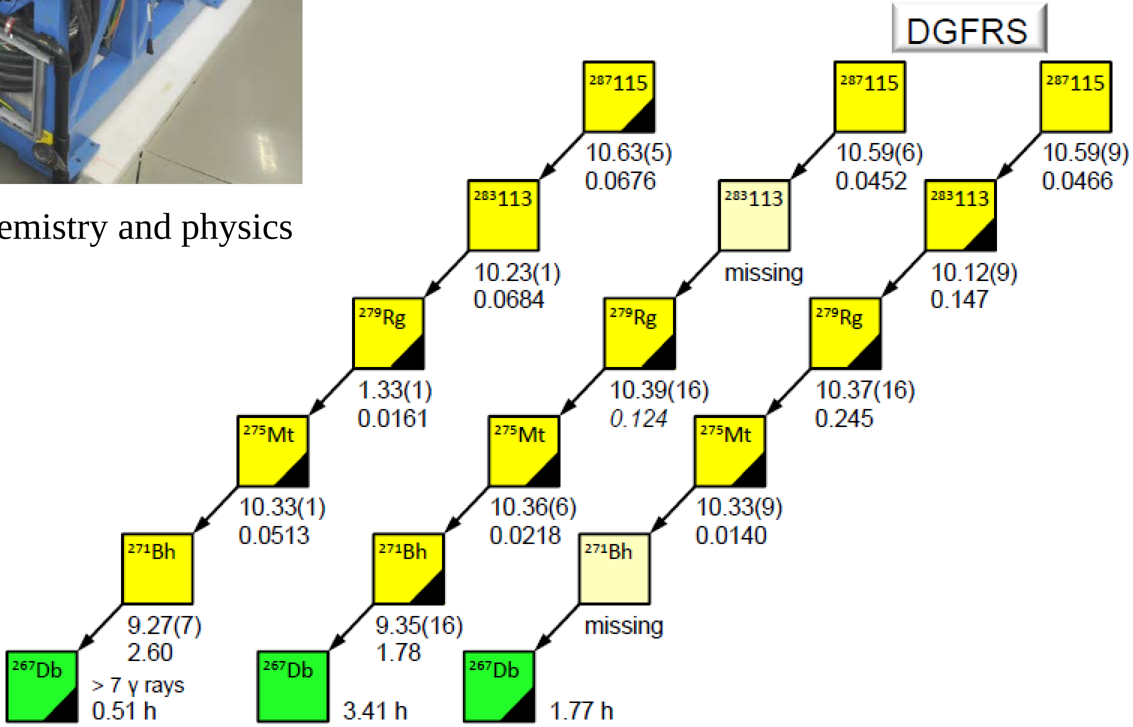


In general, α-decay in even-even parents lead to the ground state of the daughter nucleus so that the emitted particle carries away as much energy as possible and as little angular momentum as possible.
α-decays of ***odd-A*** heavy nuclei populate predominantly low-lying excited states that ***match the spin of the parent*** so that the orbital angular momentum of the α-particle can be zero.

Spectroscopy of element 115

Recoil separator for superheavy element chemistry and physics

1 chain (out of 30) is compatible with
2 chains (out of 37) associated with the
4n channel $^{287}_{115}Mc$ by *Oganessian et al.*



Mc = Moscovium

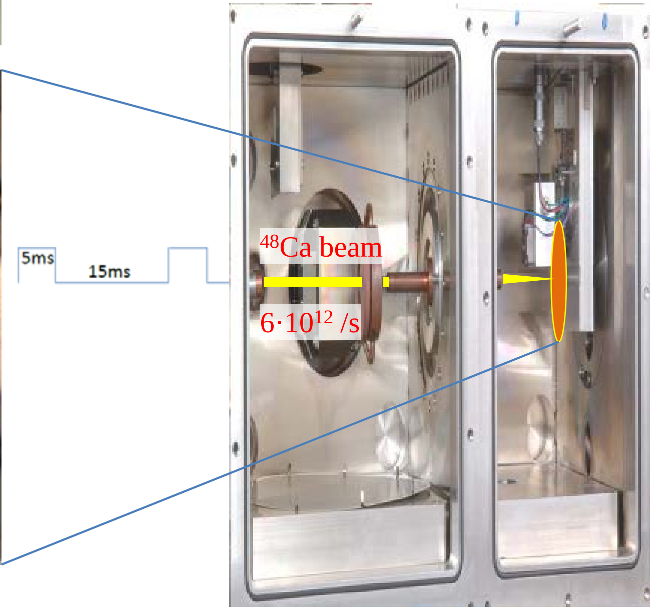
Dubna **G**as **F**illed **R**ecoil **S**eparator

Spectroscopy of element 115

^{243}Am target wheel

0.83 mg/cm^2
20 mg, 5·10^{19}
> 150 MBq α,β,γ

target chamber side view



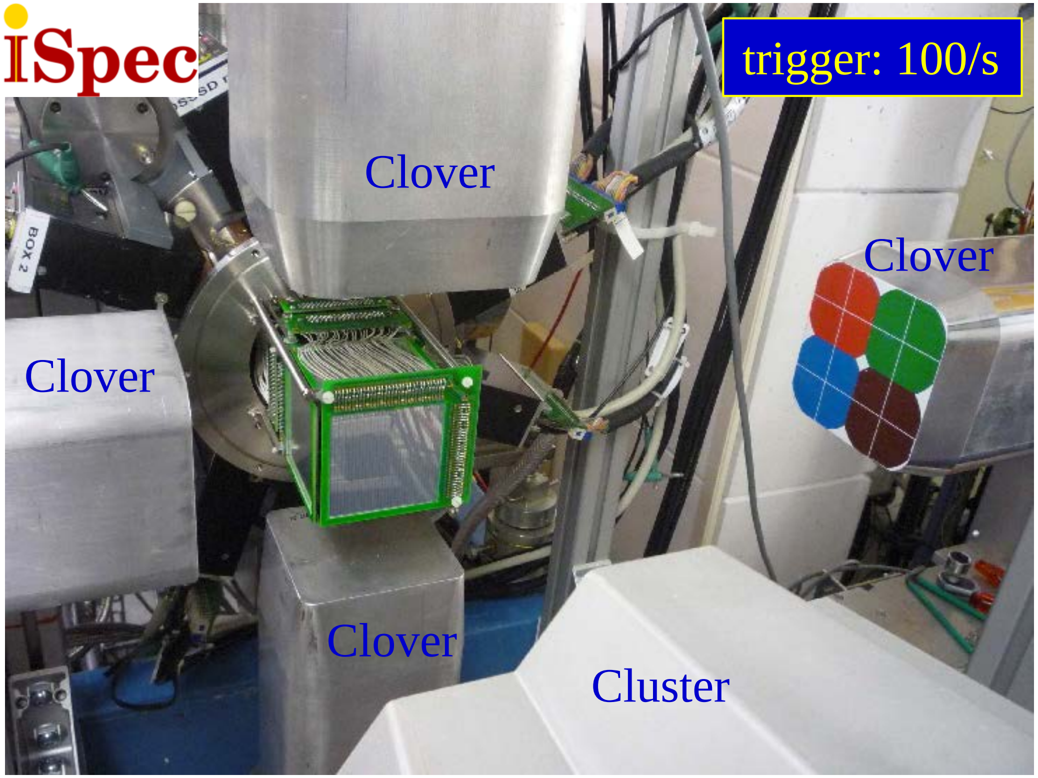
Spectroscopy of element 115

TASISpec

Highly efficient multi-coincidence spectroscopy set-up for TASCA's very compact focal plane image

1 Implantation DSSSD (1024 pixels)
4 box-DSSSDs (1024 pixels)
→ ~ 80% α-detection efficiency

4 Ge Clover (4·4 crystals)
1 Ge Cluster (7 crystals)
→ ~ 40% γ-detection eff. at 150 keV



L.-L. Andersson et al.; NIM A622, 164 (2010)
L. G. Sarmiento et al.; NIM A667, 26 (2011)

TAsca Small Image mode SPECtroscopy

TASISpec

Highly efficient multi-coincidence spectroscopy set-up for TASCA's very compact focal plane image

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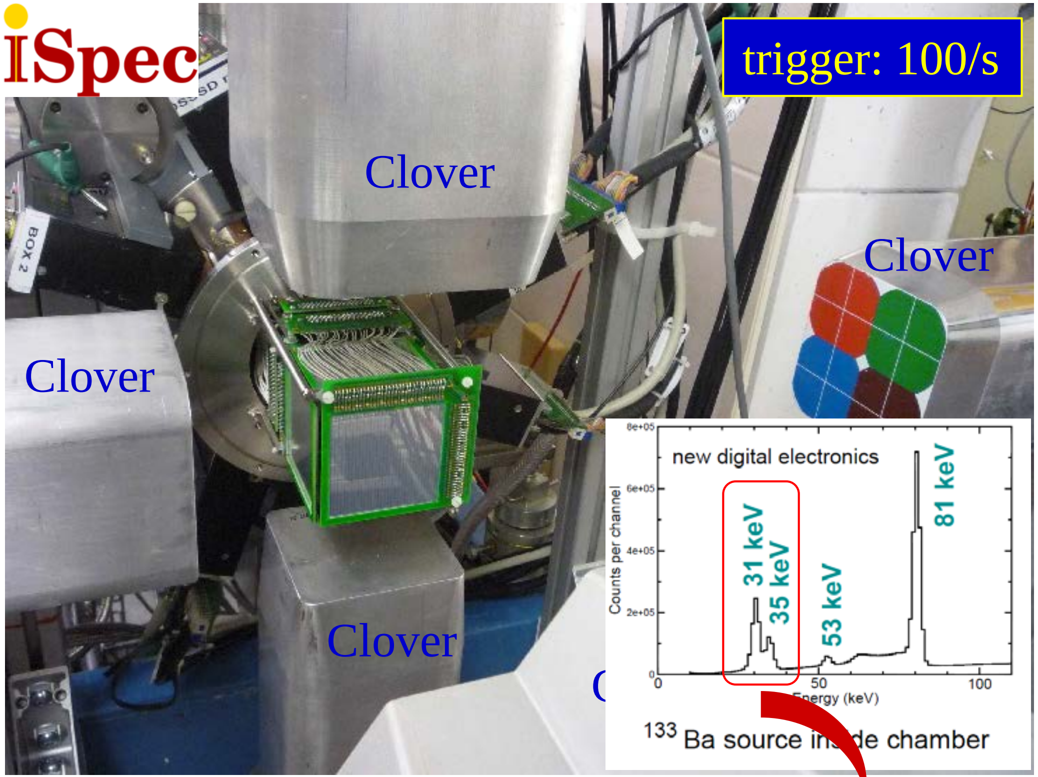
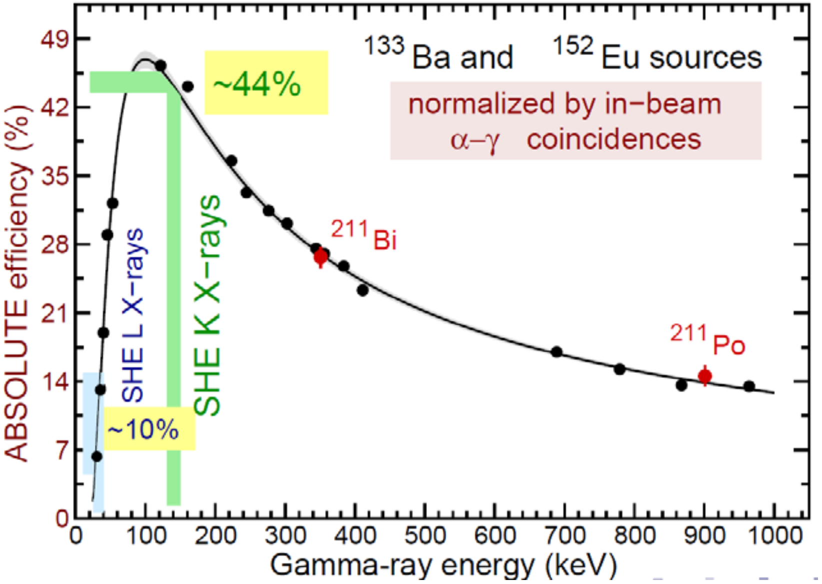
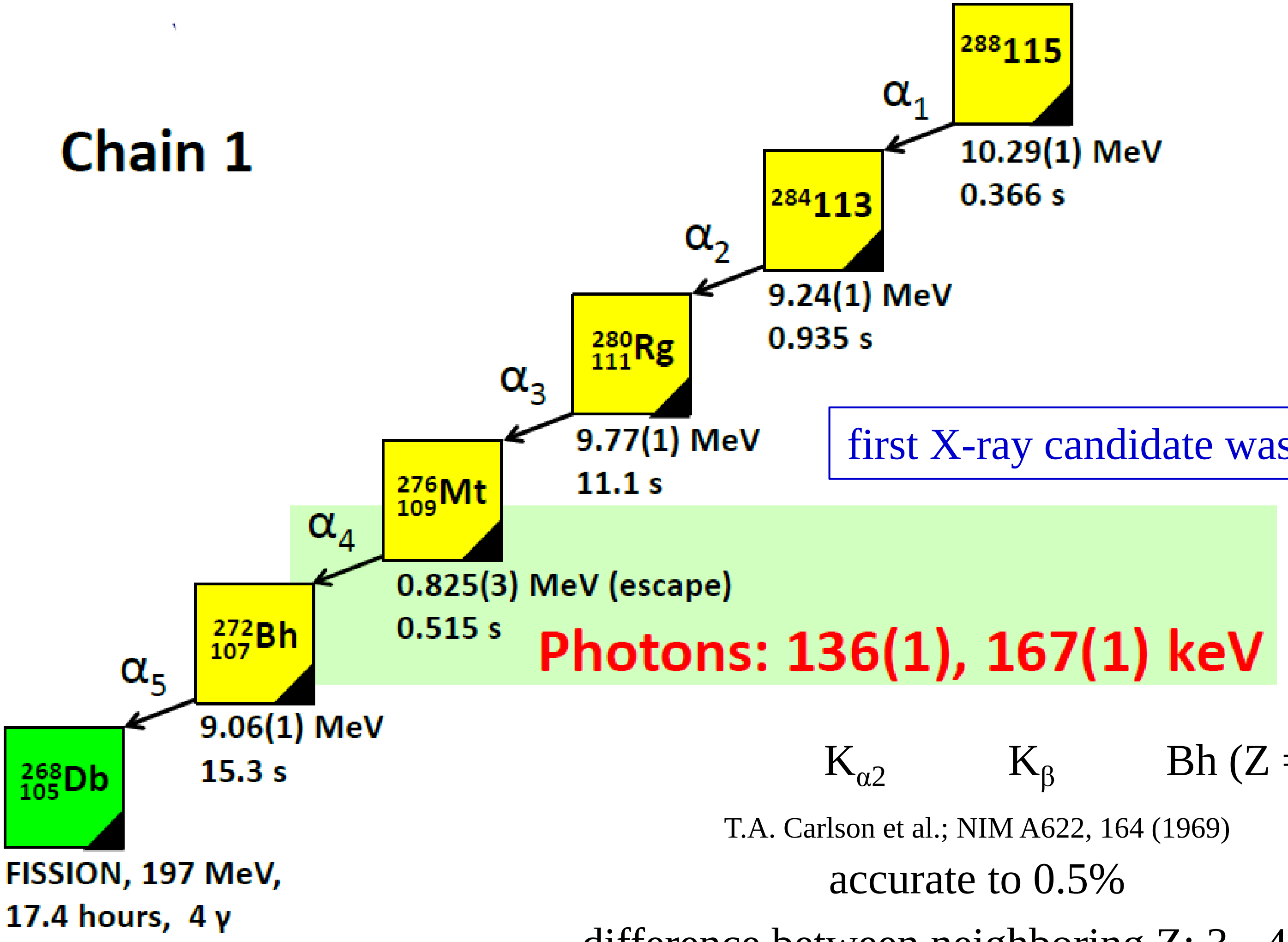


Ba$_{56}$ K X-rays ≡ Nh$_{113}$ L X-rays !!!



Characteristic X-rays



Chain 1

$^{288}115$

α_1 10.29(1) MeV
0.366 s

$^{284}113$

α_2 9.24(1) MeV
0.935 s

$^{280}_{111}\mathrm{Rg}$

α_3 9.77(1) MeV
11.1 s

$^{276}_{109}\mathrm{Mt}$

α_4 0.825(3) MeV (escape)
0.515 s

Photons: 136(1), 167(1) keV

$^{272}_{107}\mathrm{Bh}$

α_5 9.06(1) MeV
15.3 s

$^{268}_{105}\mathrm{Db}$

FISSION, 197 MeV,
17.4 hours, 4 γ

first X-ray candidate was seen!

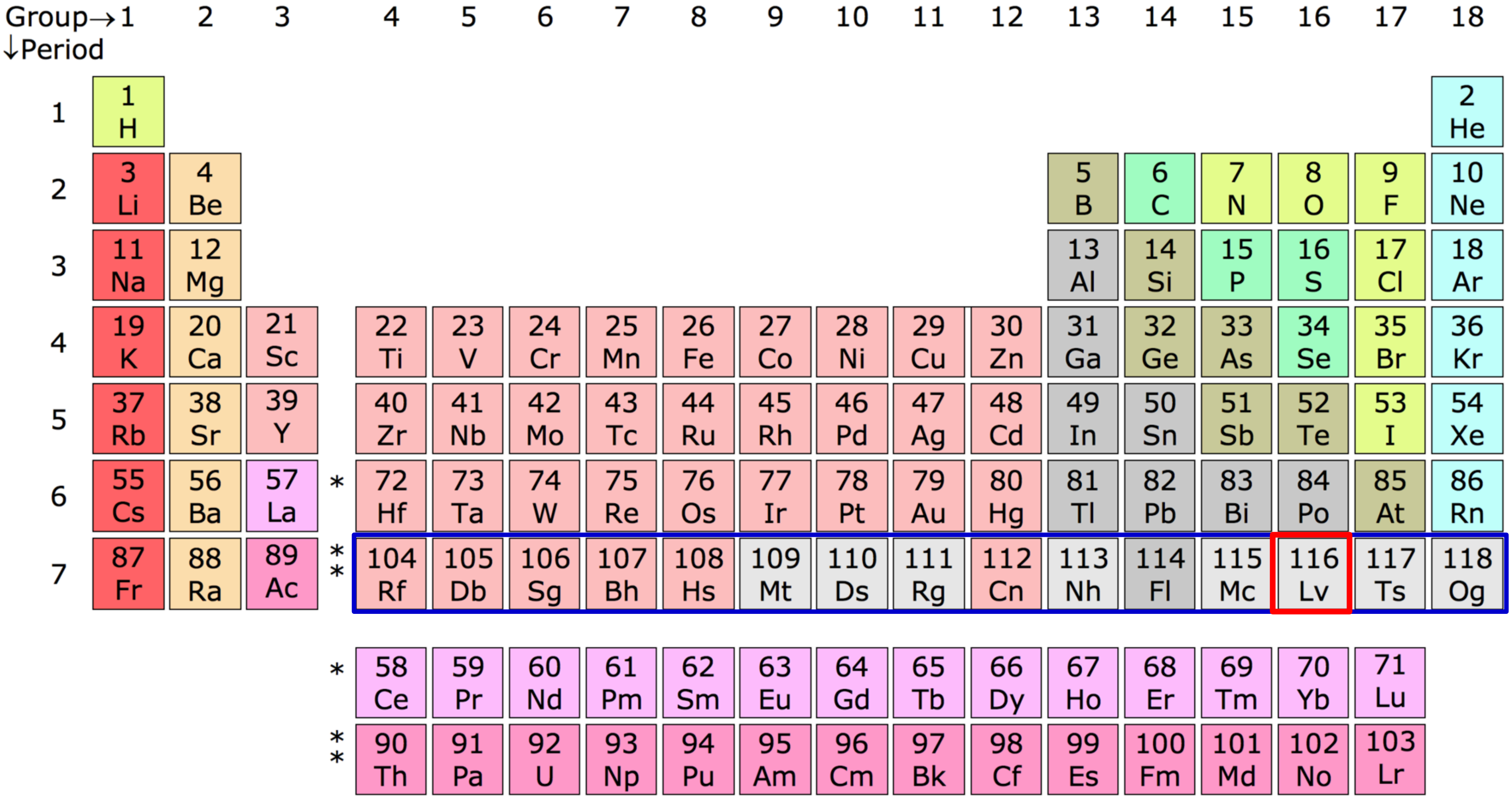
$K_{\alpha 2}$ K_{β} Bh (Z = 107)

T.A. Carlson et al.; NIM A622, 164 (1969)

accurate to 0.5%

difference between neighboring Z: 3 - 4 keV

Chemistry of superheavy elements



Group→ ↓Period	1	2	3		4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
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				**	90 Th	91 Pa	92 U	93 Np	94 Pu	95 Am	96 Cm	97 Bk	98 Cf	99 Es	100 Fm	101 Md	102 No	103 Lr	

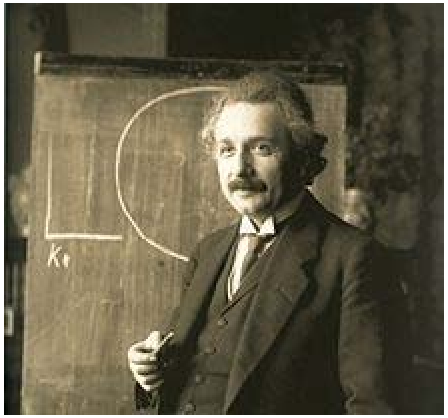
- Are the new elements in the same period?
- Does e.g. **Lv** show the same chemical properties as **O**, **S**, **Se**, **Te** and Po?

Lv = Livermorium

Chemistry of superheavy elements

relativistic effect: important for large Z

$$E = mc^2$$



- High atomic number: *strong Coulomb attraction causes electrons to move faster*.
- Causes relativistic mass increase $m = m_0(1-\beta^2)^{-1/2}$, with β = v/c; and as a consequence, *contraction of spherical orbitals* (ns, $np_{1/2}$)
- The s and $p_{1/2}$ atomic orbitals contract relativistically.
- The shrinking of the inner shells results in an increased screening of the nuclear charge, and this gives rise to an expansion of the $p_{3/2}$ and of higher angular momentum orbitals.
- *Strong spin-orbit splitting*

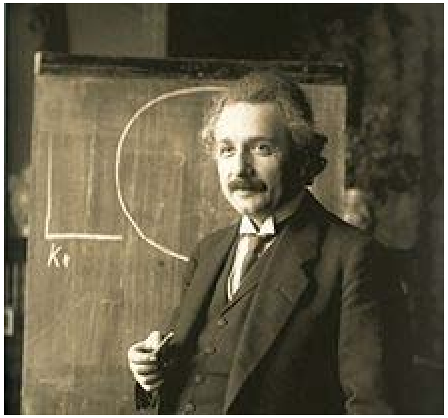
Bohr model: $E = -(2\pi^2 e^4/n^2 h^2)\cdot m\cdot Z^2$ $r = Ze^2/m\cdot v^2$ $v = (2\pi e^2/n\cdot h)\cdot Z$

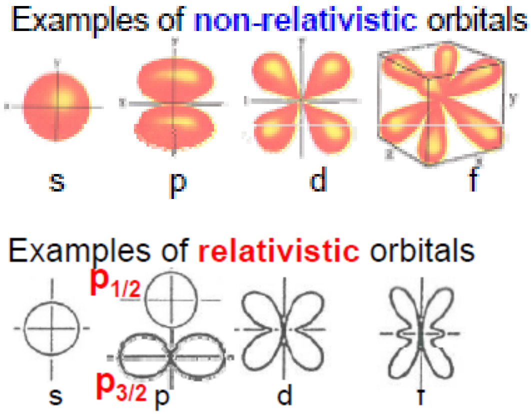
for hydrogen, m/m_0 = 1.000027, for element Z=114 , m/m_0 = 1.79, for element Z=118, m/m_0 = 1.95

Chemistry of superheavy elements

relativistic effect: important for large Z

$$E = mc^2$$

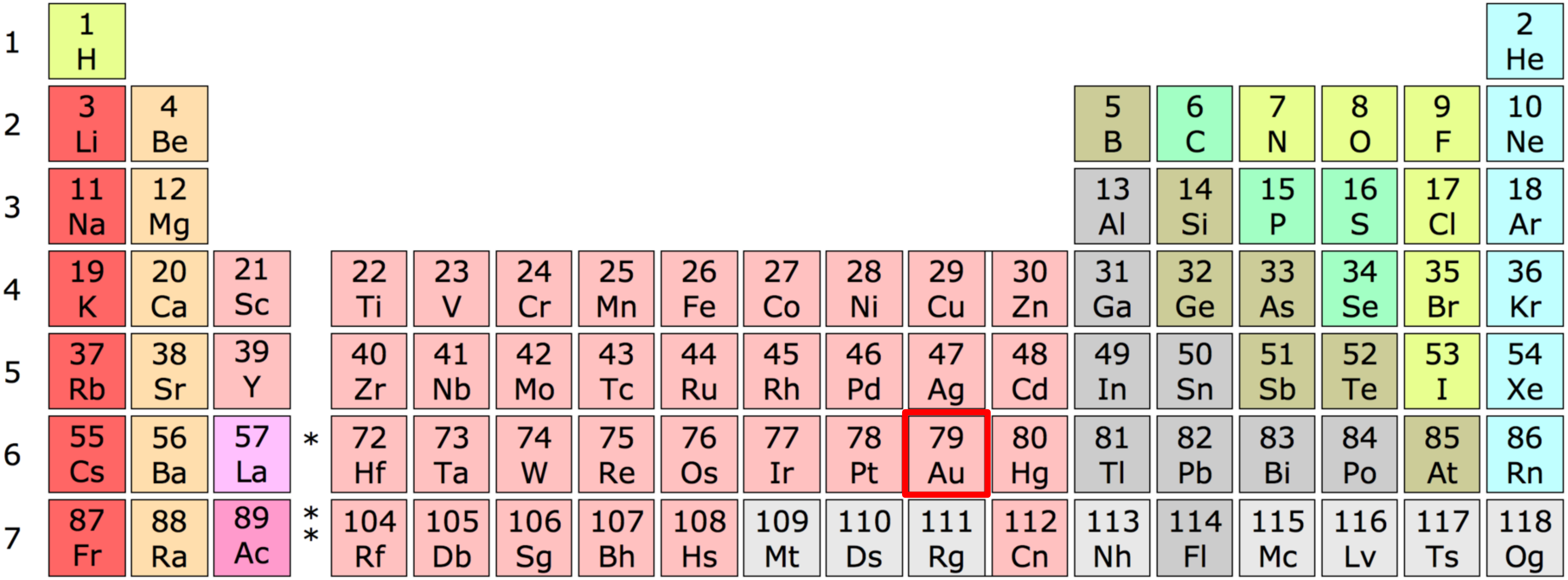




Solution of the Dirac equation (relativistic quantum mechanics) for a hydrogen-like atom:

$$E_{1s} = mc^2\sqrt{1-(Z\alpha)^2} \sim mc^2 \cdot \left(1 - \frac{(Z\alpha)^2}{2} \underbrace{- \frac{(Z\alpha)^4}{8} + \cdots}_{\text{relativistic effect}}\right)$$

Famous example of relativistic effects: the color of gold



1	1 H																		2 He
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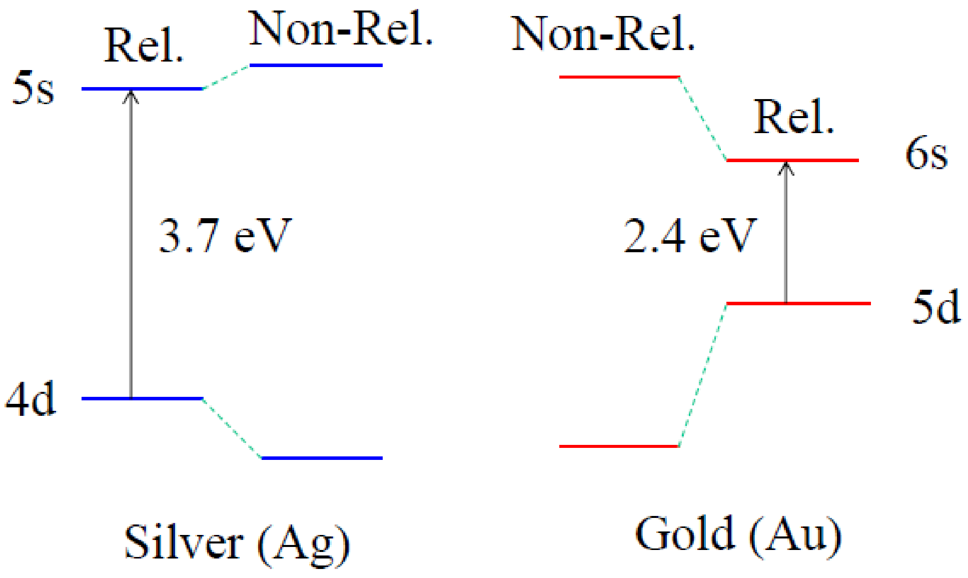
7	87 Fr	88 Ra	89 Ac	**	104 Rf	105 Db	106 Sg	107 Bh	108 Hs	109 Mt	110 Ds	111 Rg	112 Cn	113 Nh	114 Fl	115 Mc	116 Lv	117 Ts	118 Og

Gold looked like silver if there was no relativistic effect!

Famous example of relativistic effects: the color of gold



Rel.

Non-Rel.

5s

3.7 eV

4d

Silver (Ag)

Non-Rel.

Rel.

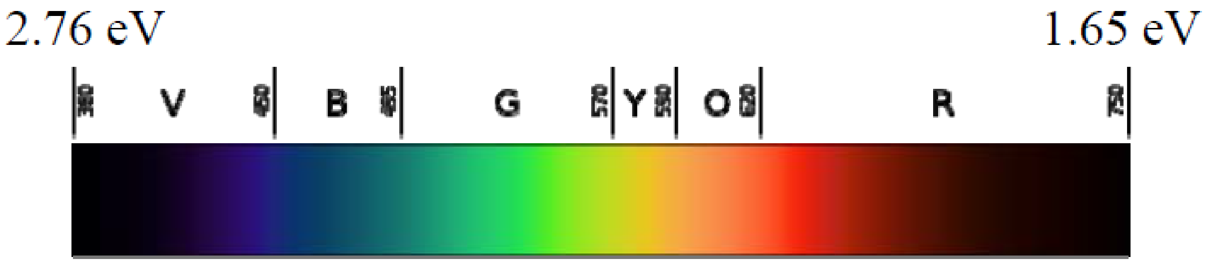
6s

2.4 eV

5d

Gold (Au)

cf. visible spectrum



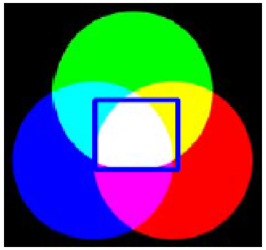
2.76 eV

1.65 eV

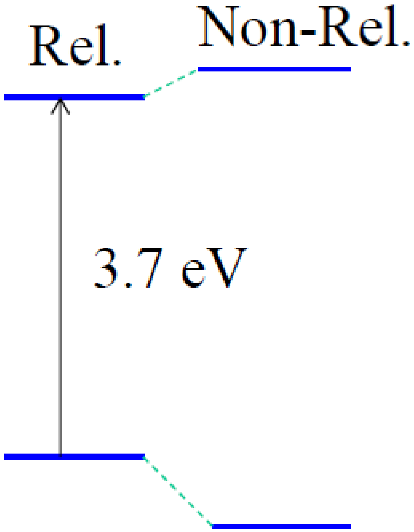
↑ 3.7 eV

↑ 2.4 eV

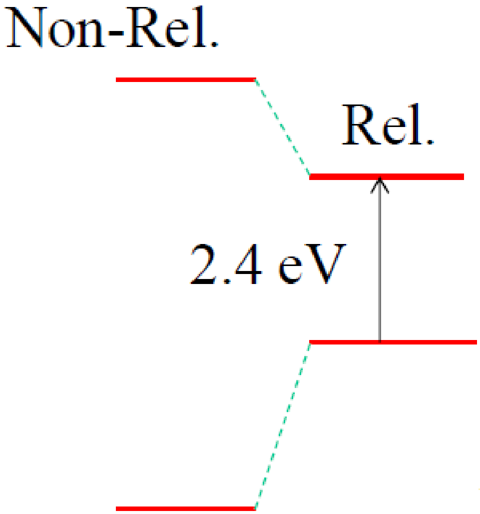
Famous example of relativistic effects: the color of gold



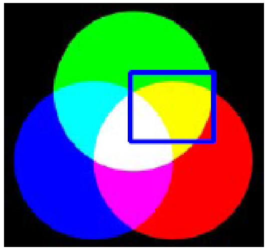
no color
absorbed



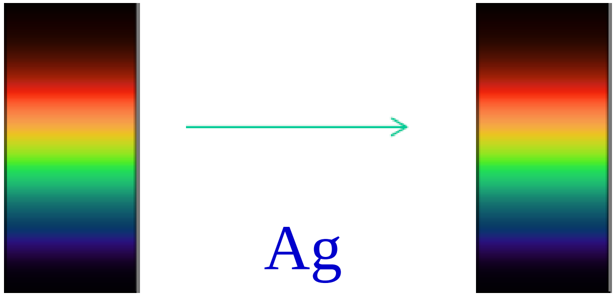
Silver (Ag)

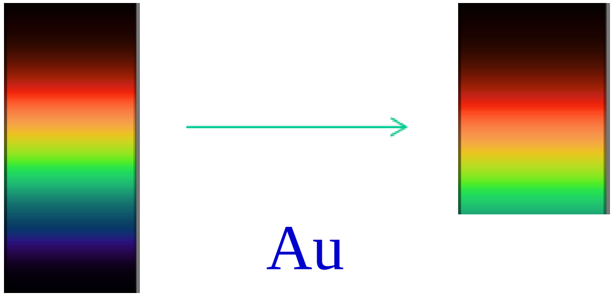


Gold (Au)

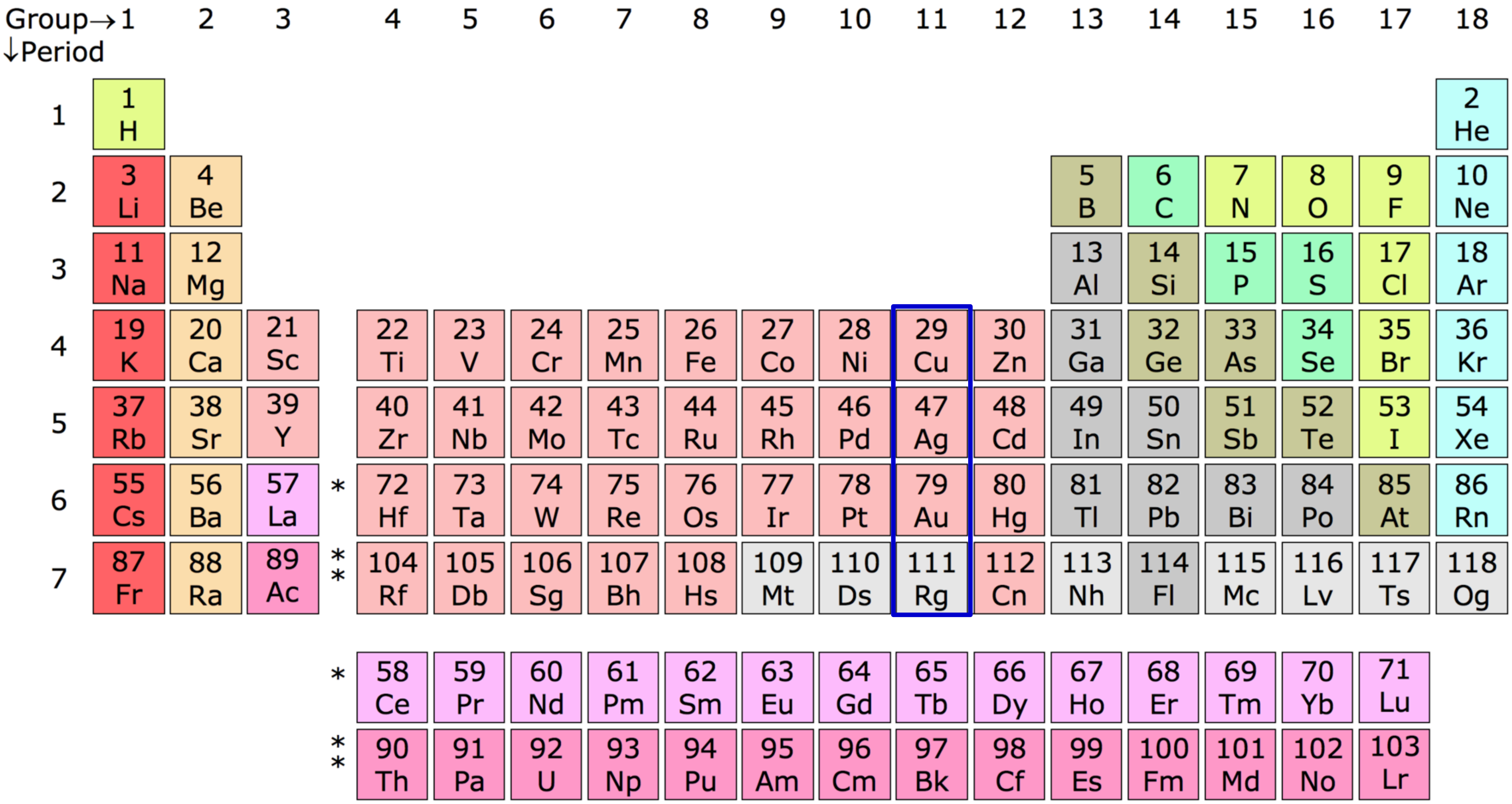


blue: absorbed





Chemistry of superheavy elements



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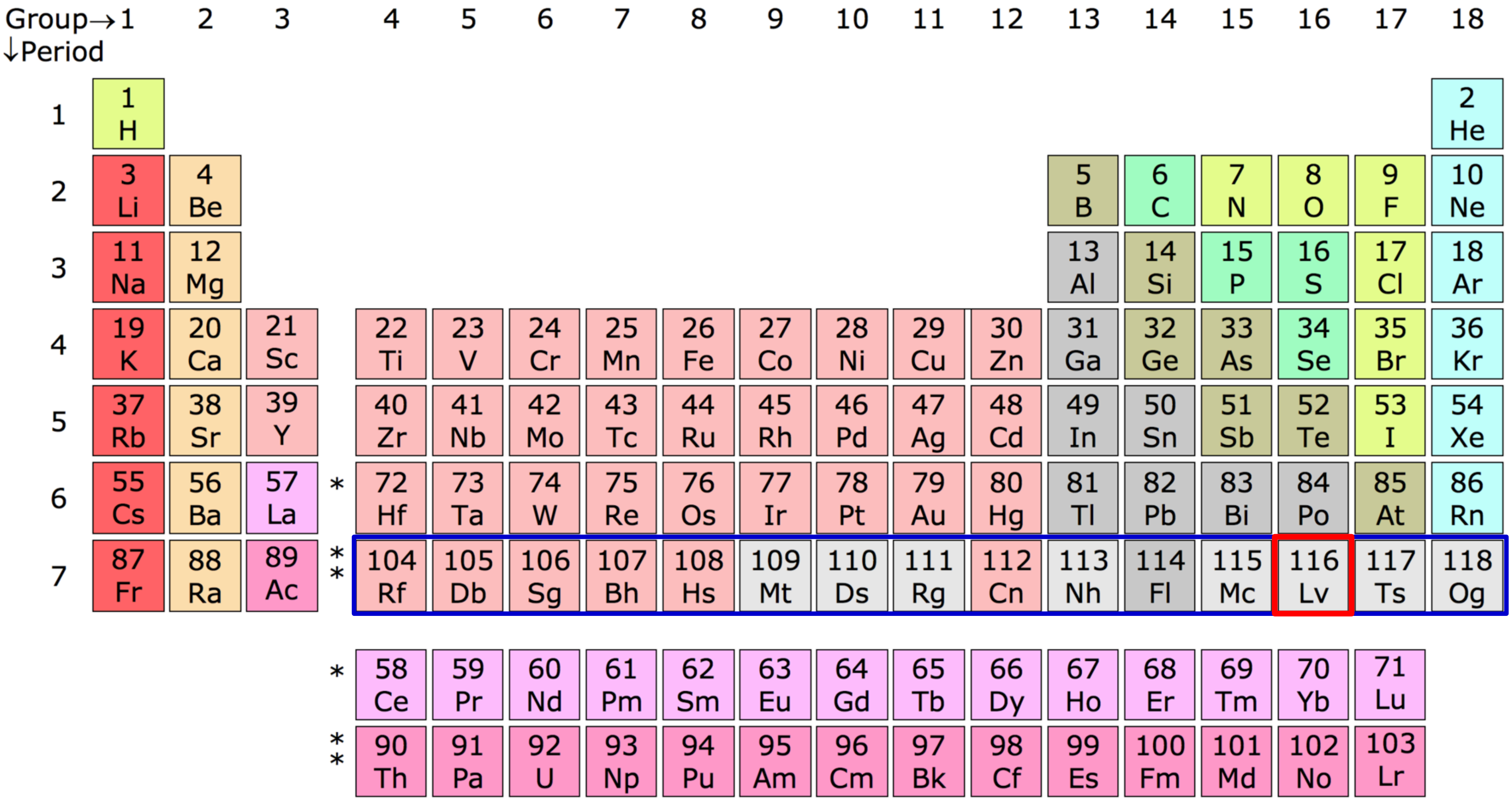
				**	90 Th	91 Pa	92 U	93 Np	94 Pu	95 Am	96 Cm	97 Bk	98 Cf	99 Es	100 Fm	101 Md	102 No	103 Lr	

- Gold is smaller than Silver
- Roentgenium is of the same size as Copper

Chemistry of superheavy elements



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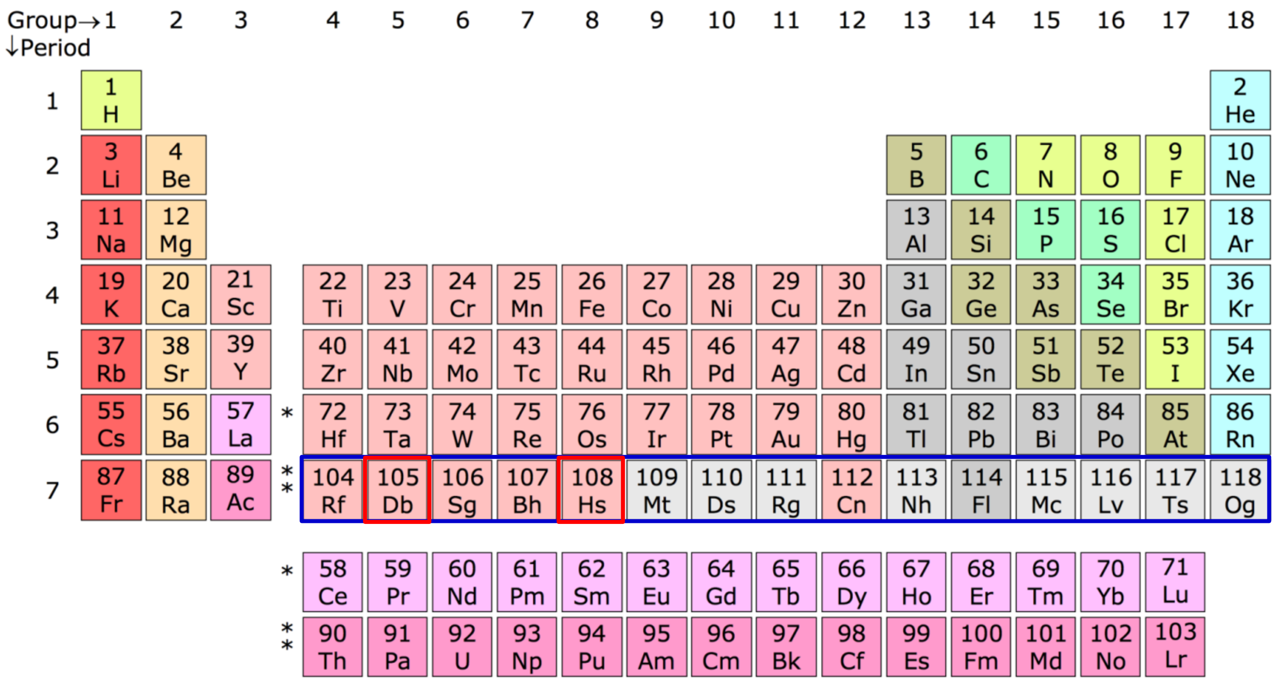
How do the relativistic effects alter the periodic table for SHE?

→ a big open question

Chemistry of superheavy elements



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5	37 Rb	38 Sr	39 Y	40 Zr	41 Nb	42 Mo	43 Tc	44 Ru	45 Rh	46 Pd	47 Ag	48 Cd	49 In	50 Sn	51 Sb	52 Te	53 I	54 Xe
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- Dubnium does not behave like Tantalum (1993)
- Sg (1997) Bh (2000) and Hs (2001) confirmed relativistic calculations predicting the expected behavior in the periodic table.
- Hassium, for instance, forms a gaseous oxide similar to Osmium.

The atomic orbitals $s_{1/2}$ and $p_{1/2}$ are contracted relativistically. The shrinking of the inner shells results in an increased *screening of the nuclear charge*, and this gives rise to an expansion of the $p_{3/2}$ and of higher angular momentum orbitals. Another relativistic effect is a *change in the spin-orbit coupling*. Both can produce drastic rearrangements of orbital levels. That is what is predicted to happen for Cn. Recent calculations indicate that the 7s orbital should be shifted below the $6d_{5/2}$ orbital due to relativistic effects. It is the large relativistic stabilization of its valence 7s orbital, combined with its closed shell electron configuration, that has led to the prediction that element 112 is chemically inert (not very reactive).

Chemistry of the superheavy elements with Z > 118 is believed to show relativistic effects that are so large that comparison with lighter elements or nonrelativistic results is meaningless.

The limits of the periodic table

- Can this go forever? NO!!!

The relativistic Dirac equation gives the ground state energy as

$$E = \frac{m_0 \cdot c^2}{\sqrt{1 + \dfrac{Z^2 \cdot \alpha^2}{n - (j + 1/2) + \sqrt{(j + 1/2)^2 - Z^2 \cdot \alpha^2}}}}$$

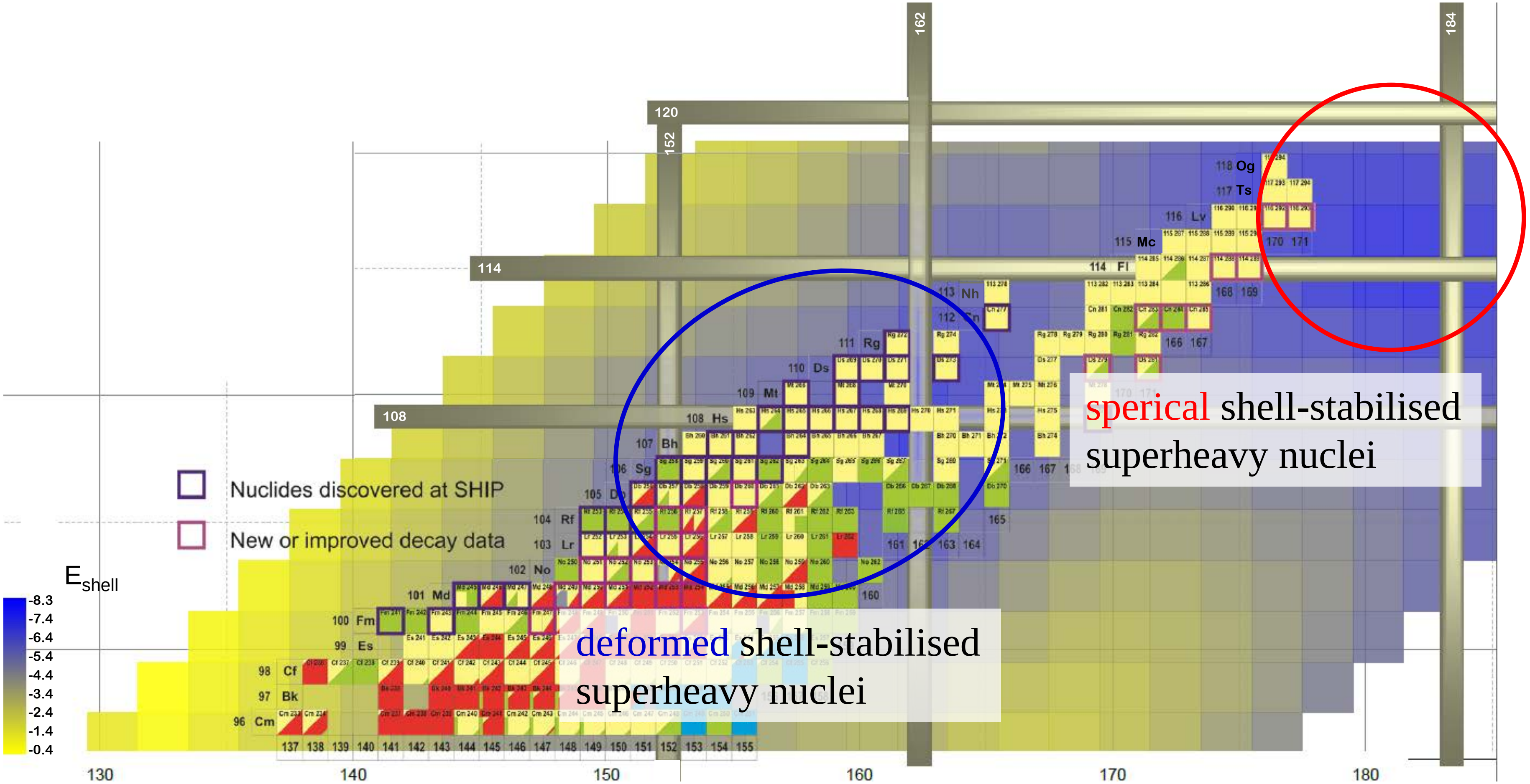
where m_0 is the rest mass of the electron. For $Z > 137$, the wave function of the Dirac ground state is oscillatory, rather than bound.

More accurate calculations taking into account the effects of the finite size of the nucleus indicate that the binding energy first exceeds $2mc^2$ for $Z > 173$.

- The end of chemistry
- Does the periodic table has limits? Yes!!!

- At some point (Z ~ 122) all the electron energy levels of adjacent elements are similar so that there are no differences in their chemical behavior.

Chart of nuclides: the domain of heavy and superheavy elements



sperical shell-stabilised superheavy nuclei

deformed shell-stabilised superheavy nuclei

Calc.: A. Sobiczewski