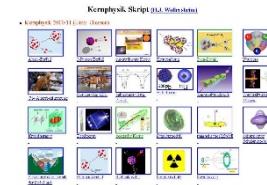


Outline: Helium burning

Lecturer: Hans-Jürgen Wollersheim

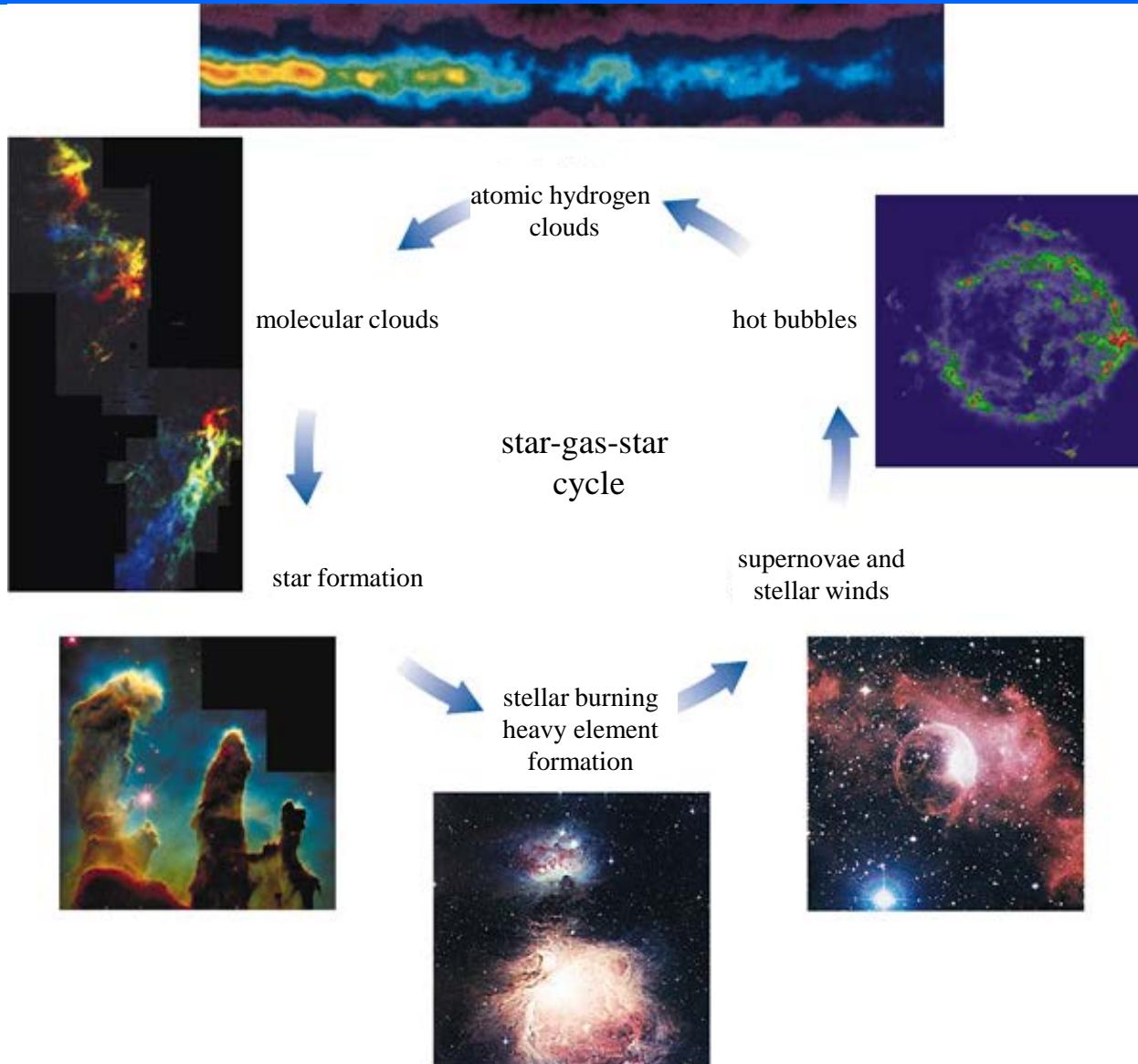
e-mail: h.j.wollersheim@gsi.de

web-page: <https://web-docs.gsi.de/~wolle/> and click on



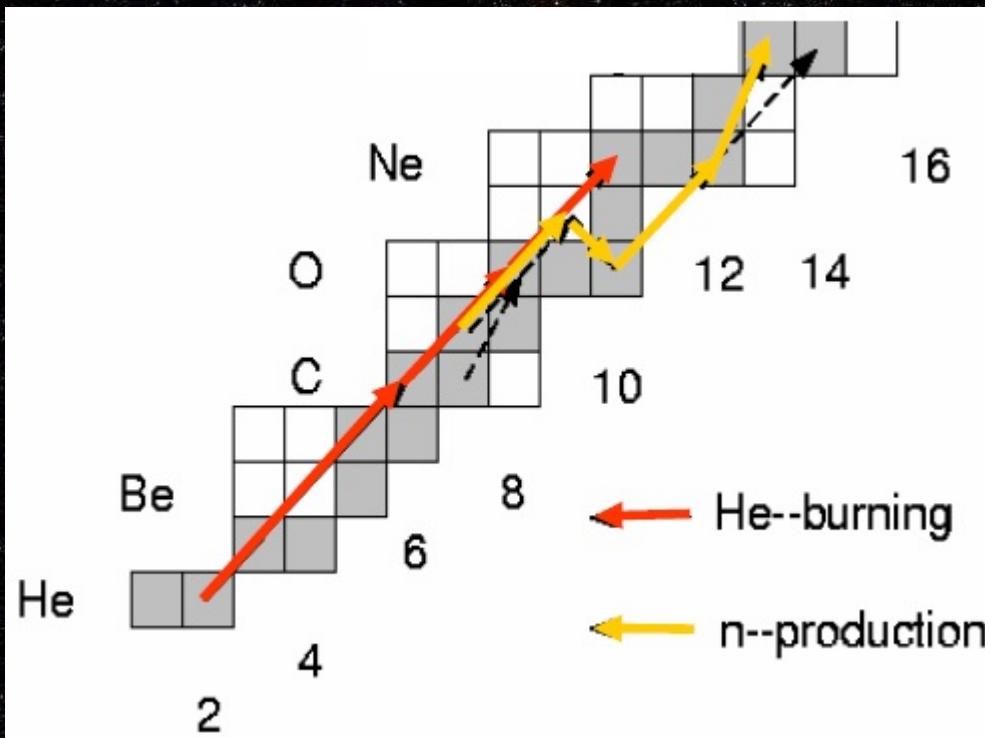
1. critical reactions in He-burning
2. the 3α reaction as two step process
3. $^{12}\text{C}(\alpha,\gamma)^{16}\text{O}$ reaction

The star-gas-star cycle

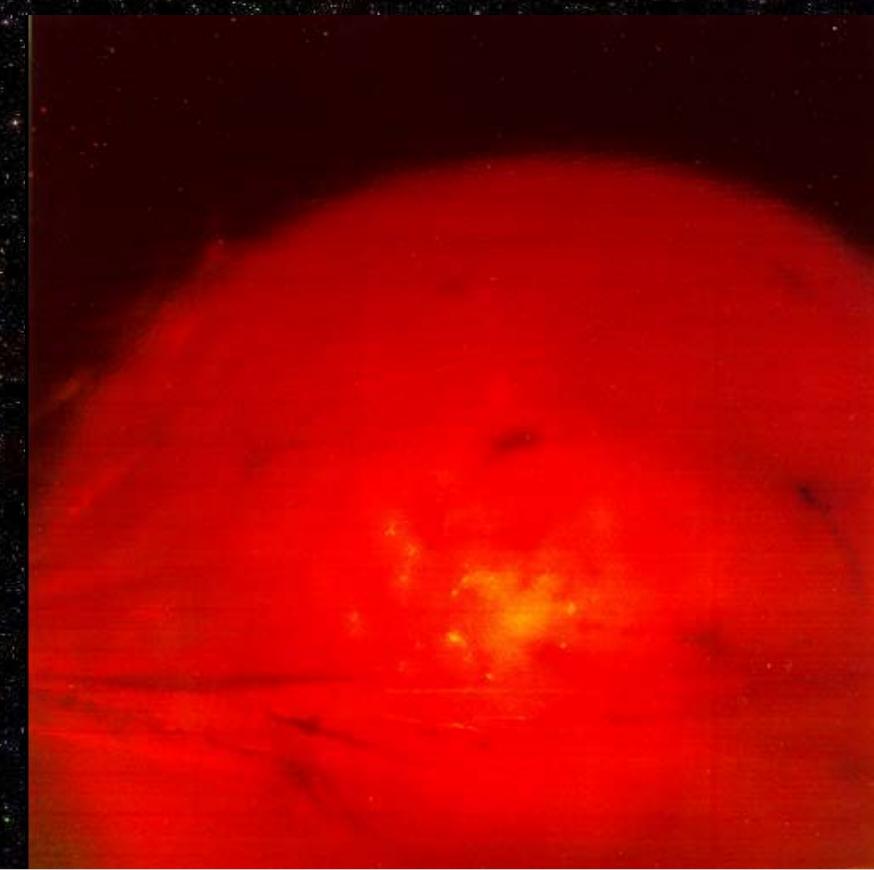


He-burning in massive stars

He-burning is ignited on the ^4He and ^{14}N ashes
of the preceding hydrogen burning phase!

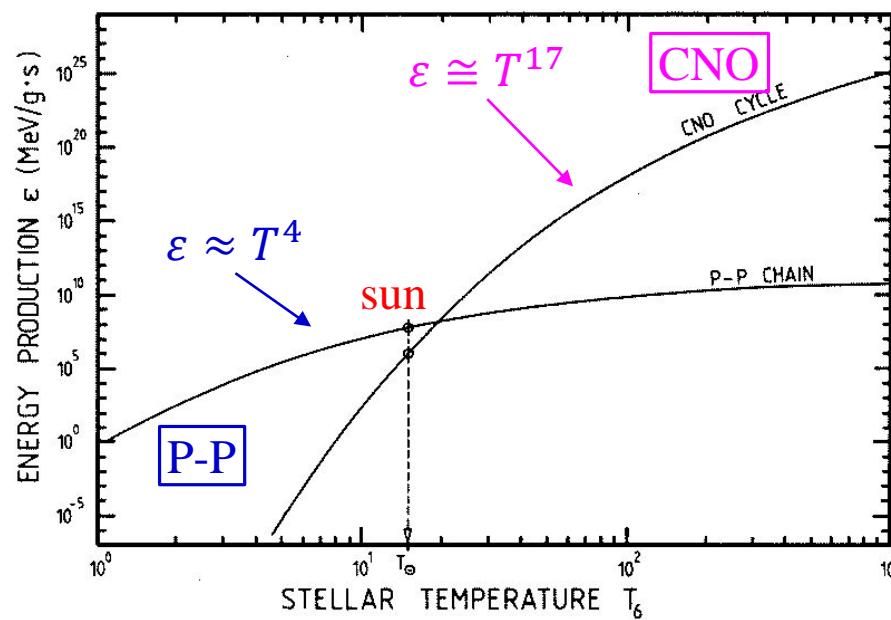
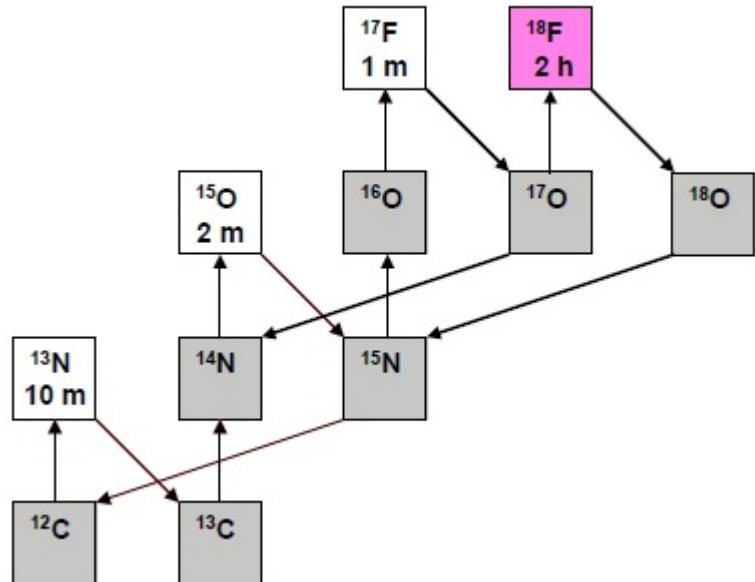


Most important reaction
– triple alpha process –
 $3\alpha \rightarrow 12C + 7.96 \text{ MeV}$

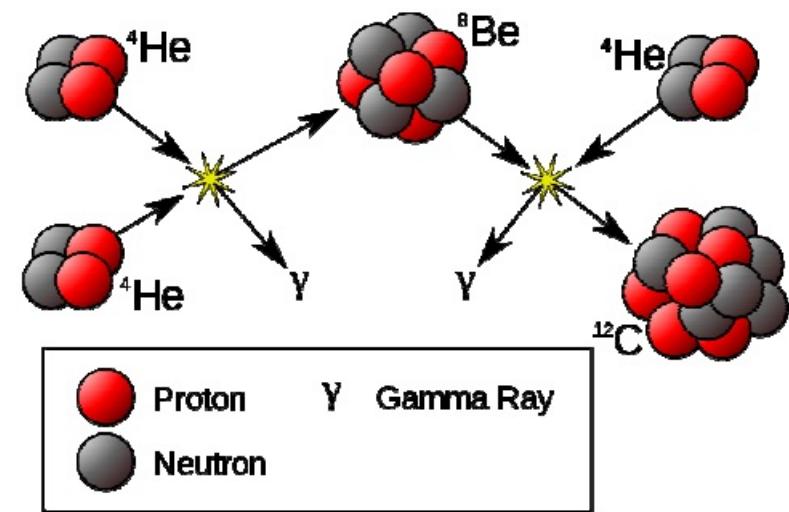
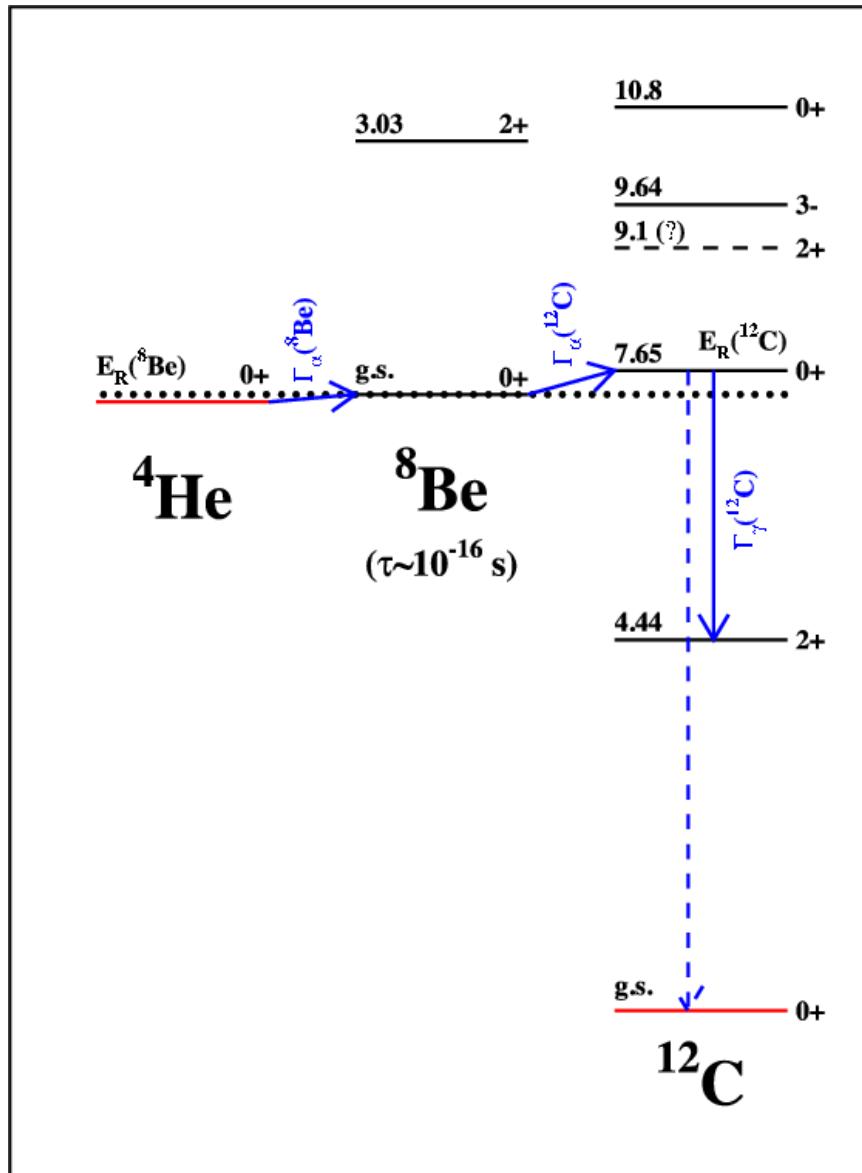


CNO-cycle

- Dominant source of energy generation in stars heavier than the sun
- What is CNO contribution to energy production in the sun? few %?
- CNO abundances in sun certain
- Stellar photospheric metallicity disagrees with helioseismology



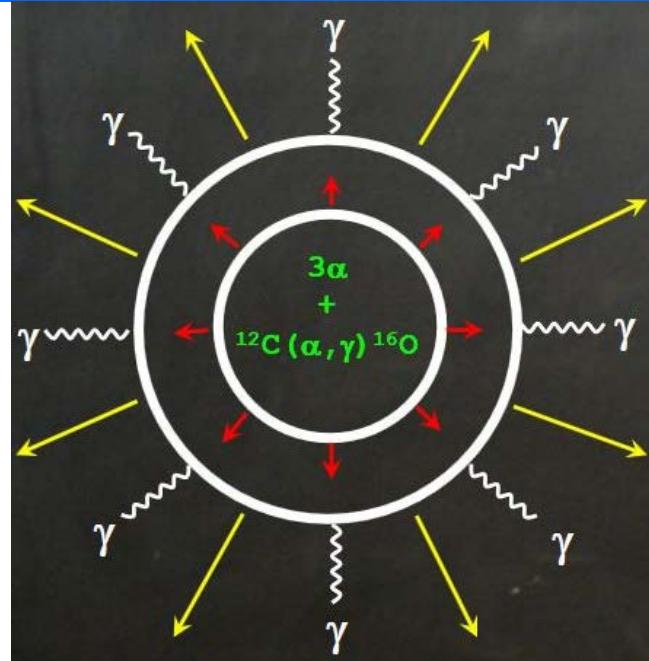
Critical reactions in He-burning



Resonance in Gamow window
- C is made !

He burning

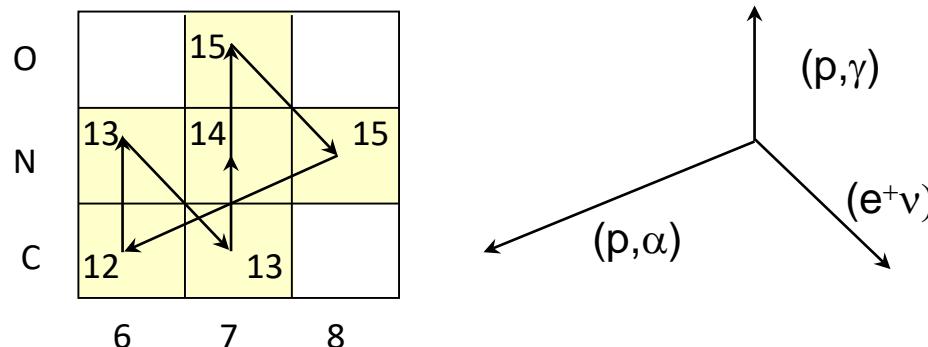
- Typical conditions:
 - Temperature: (1-2) 10^8 K
 - Density: a few $10^2 - 10^4$ g/cm³
- Net reaction: ${}^4\text{He} (2\alpha, \gamma) {}^{12}\text{C}$
- fuel: helium
- main products: carbon, oxygen
- ${}^4\text{He} + {}^4\text{He} \leftrightarrow {}^8\text{Be} + \gamma$
 ${}^8\text{Be} + {}^4\text{He} \leftrightarrow {}^{12}\text{C} + \gamma$
- and ${}^{12}\text{C} + {}^4\text{He} \rightarrow {}^{16}\text{O} + \gamma$
- **difficulty: lifetime of ${}^8\text{Be} \sim 10^{-16}$ s**
→ Hoyle state (resonance in ${}^{12}\text{C}$ at E=7.68 MeV)
- Other products: ${}^{21,22}\text{Ne}$, ${}^{25,26}\text{Mg}$, ${}^{36}\text{S}$, ${}^{37}\text{Cl}$, ${}^{40}\text{K}$, ${}^{40}\text{Ar}$
- ${}^{14}\text{N} (\alpha, \gamma) {}^{18}\text{F}$ (e^+, ν) ${}^{18}\text{O} (\alpha, \gamma) {}^{22}\text{Ne}$ (α, n) ${}^{25}\text{Mg}$



$$kT \sim 8.6 \times 10^{-8} T[\text{K}] \text{ keV}$$

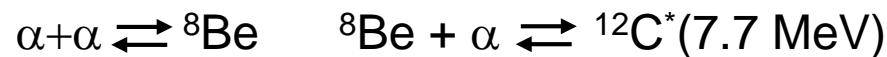
CNO-cycle

CNO cycle



Cycle limited by β -decay of ^{13}N ($t \sim 10$ min) and ^{15}O ($t \sim 2$ min)

CNO isotopes act as catalysts



nucleosynthesis

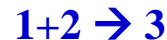
energy production

changes in stellar conditions \Rightarrow changes in energy production and nucleosynthesis

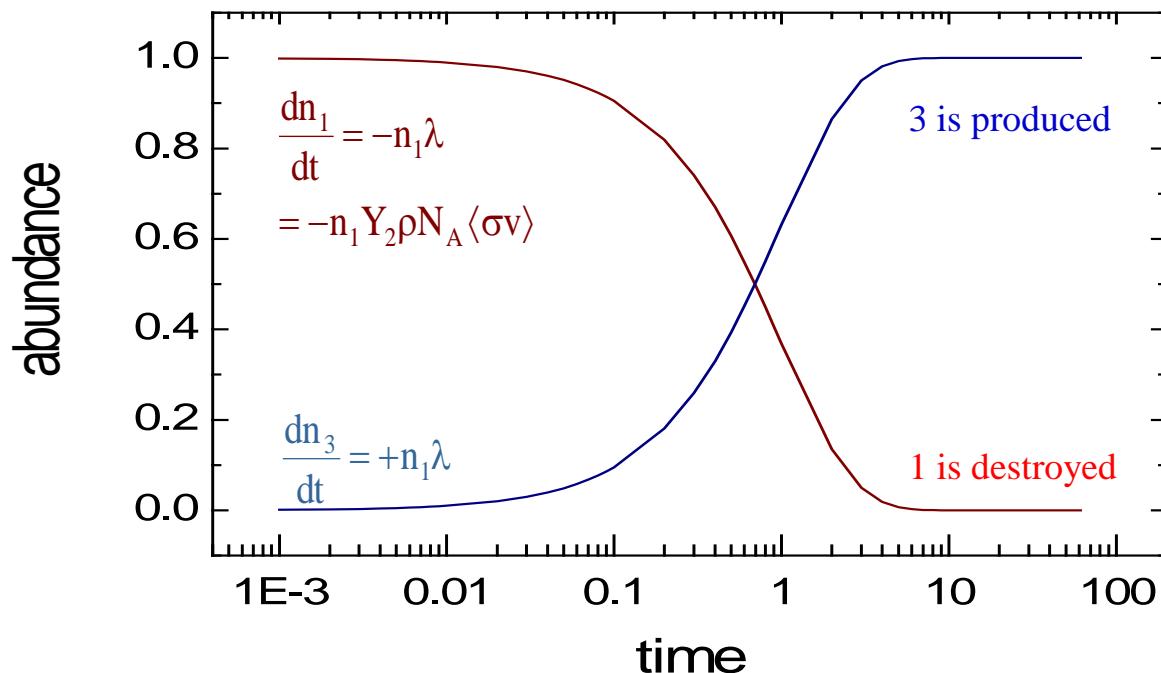
need to know REACTION RATE at all temperatures to determine ENERGY PRODUCTION

Abundance change and lifetimes

consider reaction



where **1** is destroyed through capture of **2** and **3** is produced

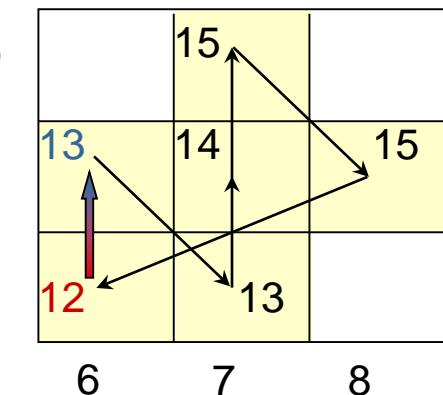


define:

lifetime of 1 against destruction by with 2:

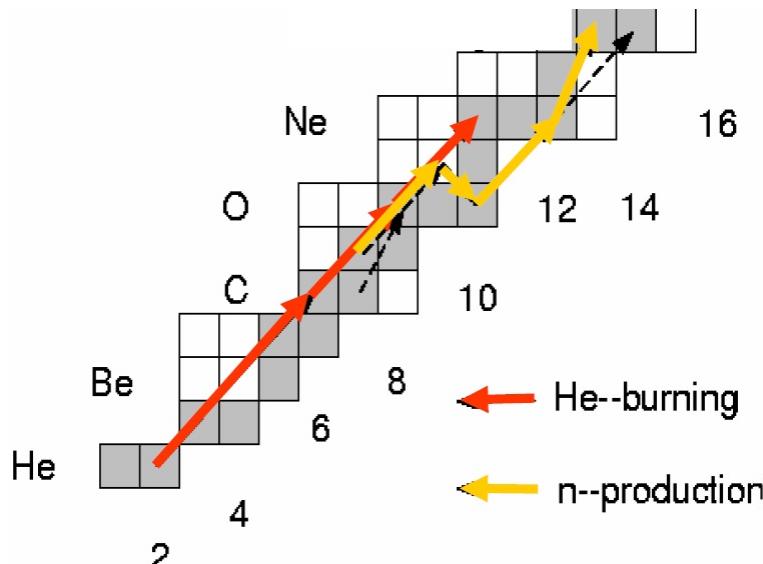
$$\tau = \frac{1}{\lambda} = \frac{1}{Y_1 \rho N_A \langle \sigma v \rangle}$$

example

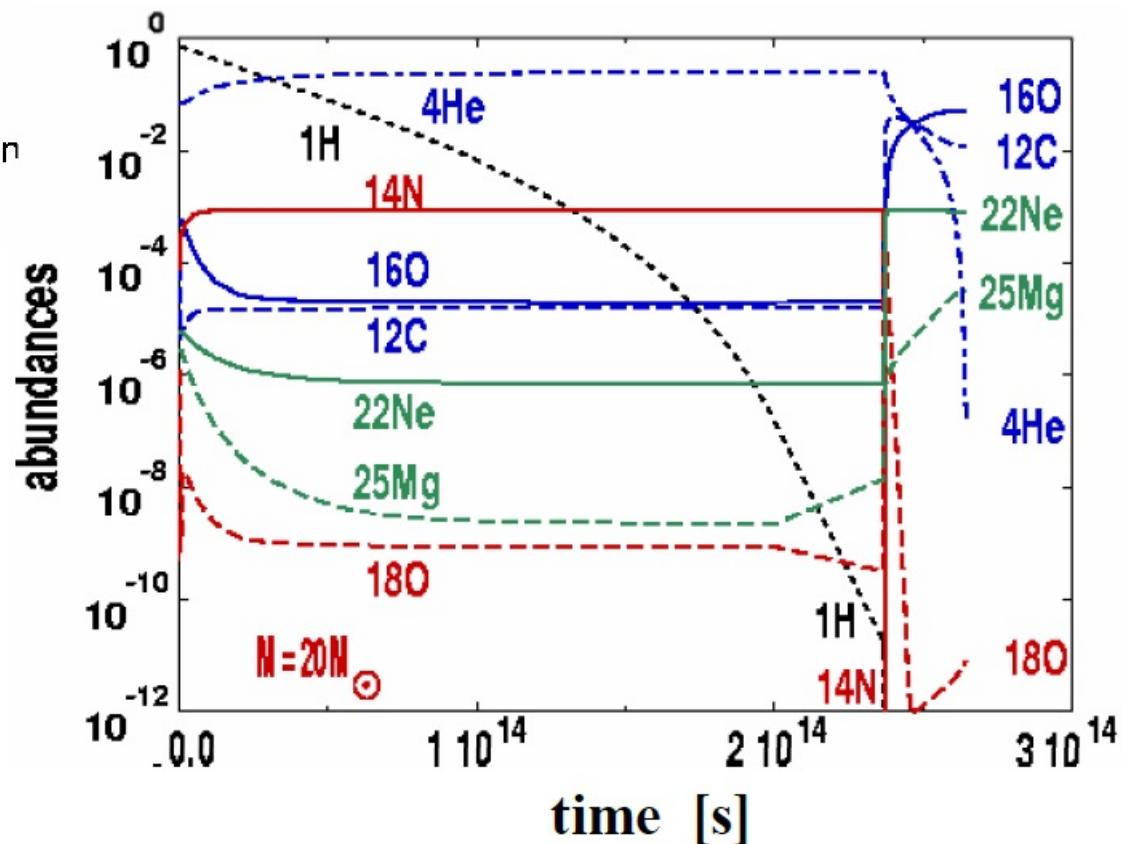


need to know REACTION RATE at all temperatures to determine NUCLEOSYNTHESIS

Abundance evolution in stellar core

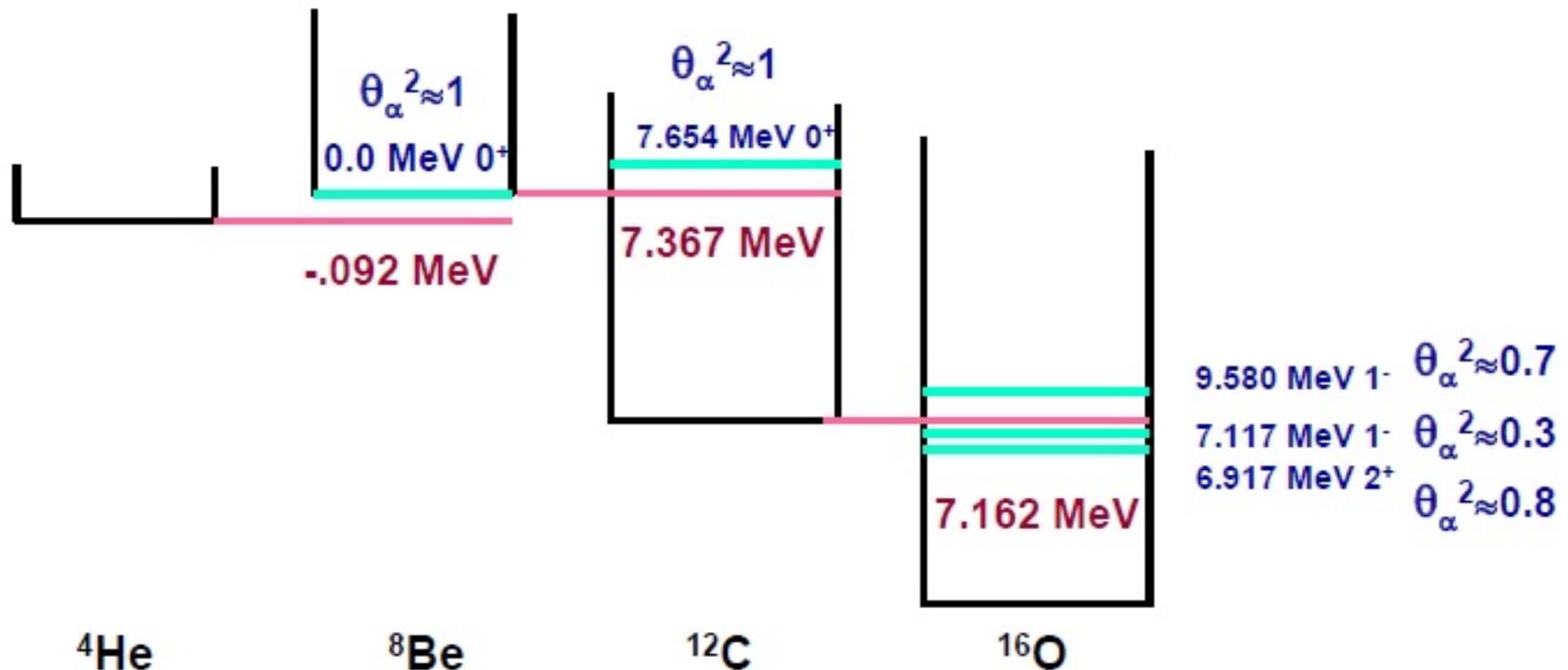


Decline of ${}^4\text{He}$ (time scale)
 increase in ${}^{12}\text{C}$, ${}^{16}\text{O}$
 \rightarrow equilibrium ${}^{12}\text{C}/{}^{16}\text{O}$
 rapid decline in ${}^{14}\text{N}$

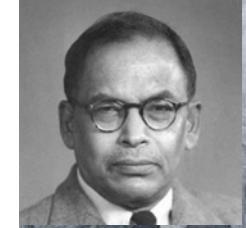


The case of 3α and $^{12}\text{C}(\alpha,\gamma)^{16}\text{O}$

Reaction rates determined by
 α cluster state configurations
providing strong resonances!

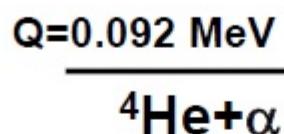


The ($\alpha\alpha\alpha$) reaction as two step process



Meghnad Saha

first step!



$$T_{1/2}({}^8\text{Be}) = 9.7 \cdot 10^{-17} \text{ s}$$

$$\Gamma_\alpha = 6.8 \text{ eV}$$

pure α cluster configuration

Example for ${}^8\text{Be}$ equilibrium abundance:

Case of typical He-burning: $T=0.1\text{GK} \Rightarrow T_9=0.1$; $\rho=10^5 \text{ g/cm}^3$

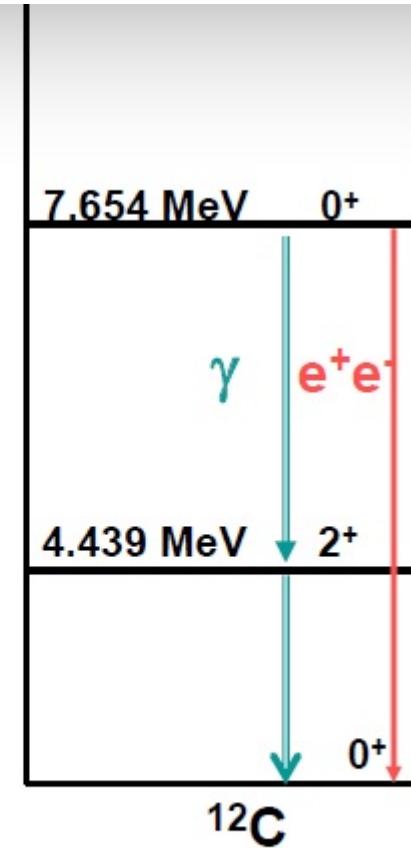
Resonant capture on ${}^8\text{Be}$

The Hoyle resonance!

$$\begin{array}{c} E_R = 0.287 \text{ MeV} \\ Q = 7.367 \text{ MeV} \\ {}^8\text{Be} + \alpha \end{array}$$

$$N_A \langle \sigma v \rangle = 1.54 \cdot 10^{11} \cdot \omega \gamma \cdot \left(\frac{1}{\mu \cdot T_9} \right)^{3/2} \cdot e^{-\left(\frac{11.605 - E_R}{T_9} \right)}$$

$$\omega \gamma = (2J+1) \cdot \frac{\Gamma_{in} \cdot \Gamma_{out}}{\Gamma_{tot}}$$



Decay by sequential E2 γ transitions
or internal $e^+ e^-$ pair conversion

The resonance strength

$$\omega\gamma = \frac{\Gamma_\alpha \cdot (\Gamma_\gamma + \Gamma_{e^+e^-})}{\Gamma_\alpha + \Gamma_\gamma + \Gamma_{e^+e^-}}$$

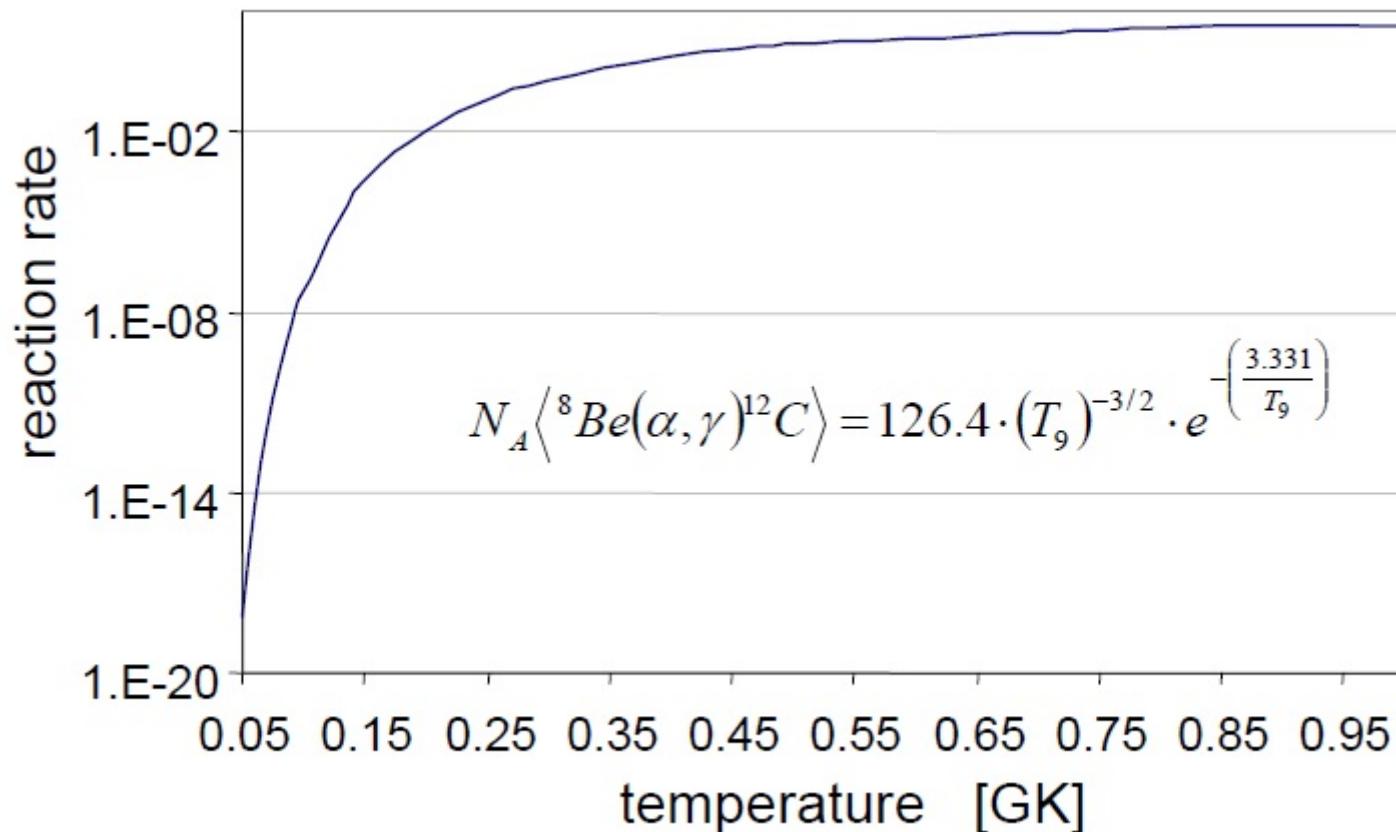
$$\Gamma_\alpha = 8.09 \pm 1.08 \text{ eV}$$

$$\Gamma_\gamma = 3.58 \pm 0.5 \text{ meV} \quad \frac{\Gamma_{rad}}{\Gamma_{tot}} = 4.12 \cdot 10^{-4}$$

$$\Gamma_{e^+e^-} = 60.6 \pm 3.9 \text{ } \mu\text{eV}$$

$$\boxed{\omega\gamma = 3.58 \cdot 10^{-9} \text{ MeV} \quad \pm 12\%}$$

The ${}^8\text{Be} + \alpha$ reaction rate



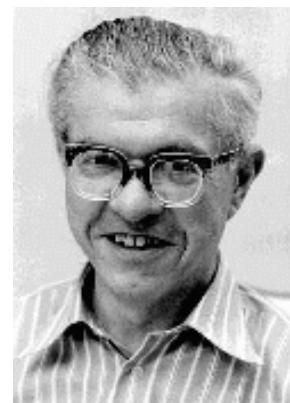
How did they do the experiment?

- Used a deuterium beam on a ^{11}B target to produce ^{12}B via a (d,p) reaction.
- ^{12}B β -decays within 20 ms into the second excited state in ^{12}C
- This state then immediately decays under alpha emission into ^8Be
- Which immediately decays into 2 alpha particles

So they saw after the delay of the β -decay 3 alpha particles coming from their target after a few ms of irradiation

This proved that the state can also be formed by the 3 alpha process ...

→ removed the major roadblock for the theory that elements are made in stars
→ Nobel Prize in Physics 1983 for Willy Fowler (alone !)



The total $\langle \alpha\alpha\alpha \rangle$ rate

$$r_{\alpha\alpha\alpha} = N_{^8Be} \cdot \rho \cdot \frac{X_\alpha}{A_\alpha} \cdot N_A \langle ^8Be(\alpha, \gamma)^{12}C \rangle$$

Step 1

$$N(^8Be) = 6 \cdot 10^{-35} \cdot N_\alpha^2 \cdot T_9^{-3/2} \cdot e^{\left(-\frac{1.068}{T_9}\right)}$$

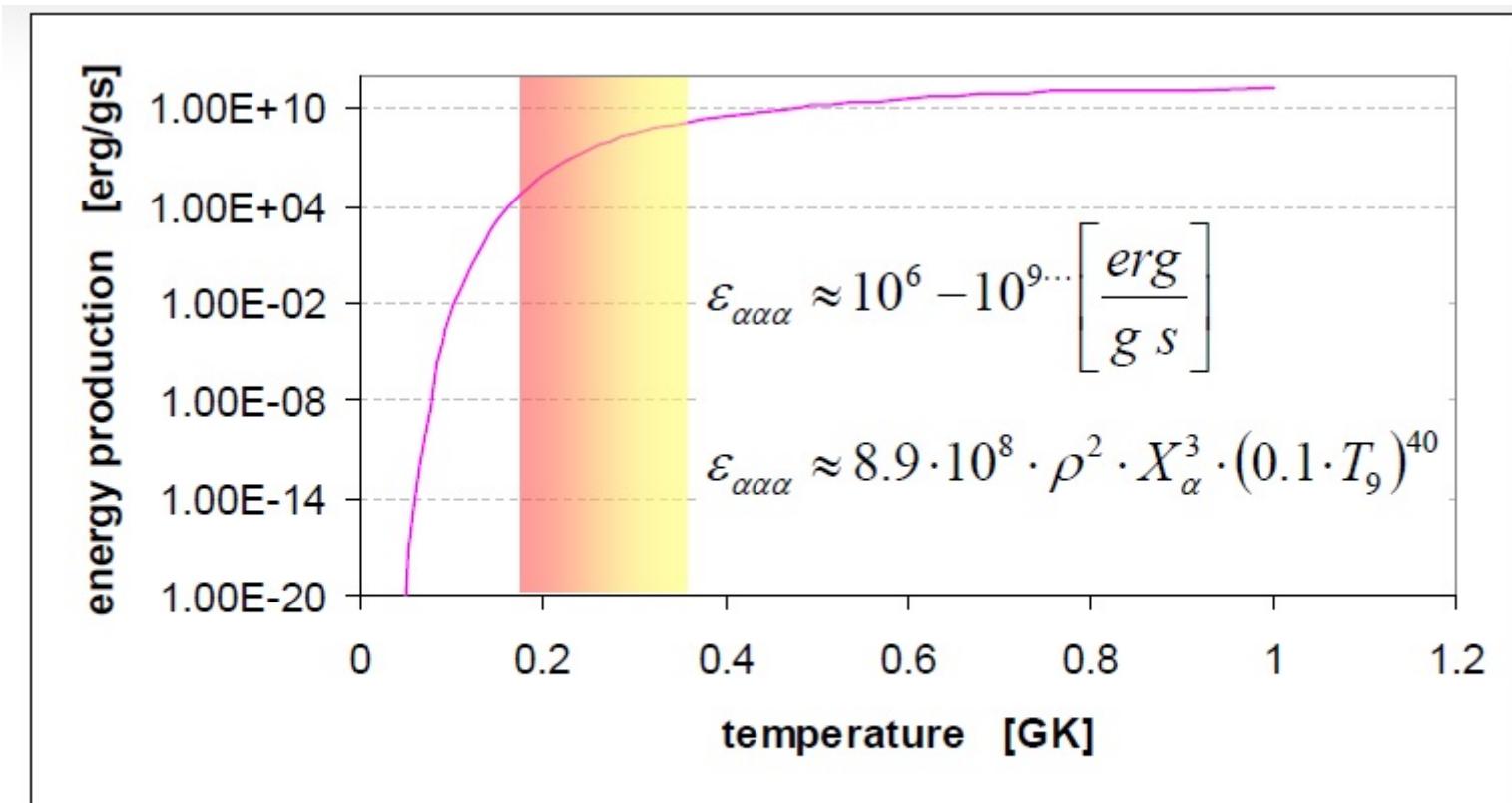
Step 2

$$N_A \langle ^8Be(\alpha, \gamma)^{12}C \rangle = 126.4 \cdot (T_9)^{-3/2} \cdot e^{-\left(\frac{3.331}{T_9}\right)}$$

$$r_{\alpha\alpha\alpha} = \frac{1.26 \cdot 10^{-56}}{1 + \delta_{\alpha\alpha}} \cdot N_\alpha^3 \cdot T_9^{-3} \cdot e^{\left(-\frac{11.605 \cdot (0.092 + 0.278)}{T_9}\right)}$$

$$r_{\alpha\alpha\alpha} = 1.38 \cdot 10^{15} \cdot \rho^3 \cdot \left(\frac{X_\alpha}{4}\right)^3 \cdot T_9^{-3} \cdot e^{\left(-\frac{4.294}{T_9}\right)} \quad [cm^{-3}s^{-1}]$$

Example: $\rho=10^5$ g/cm³



T-dependent main energy source for stellar He-burning

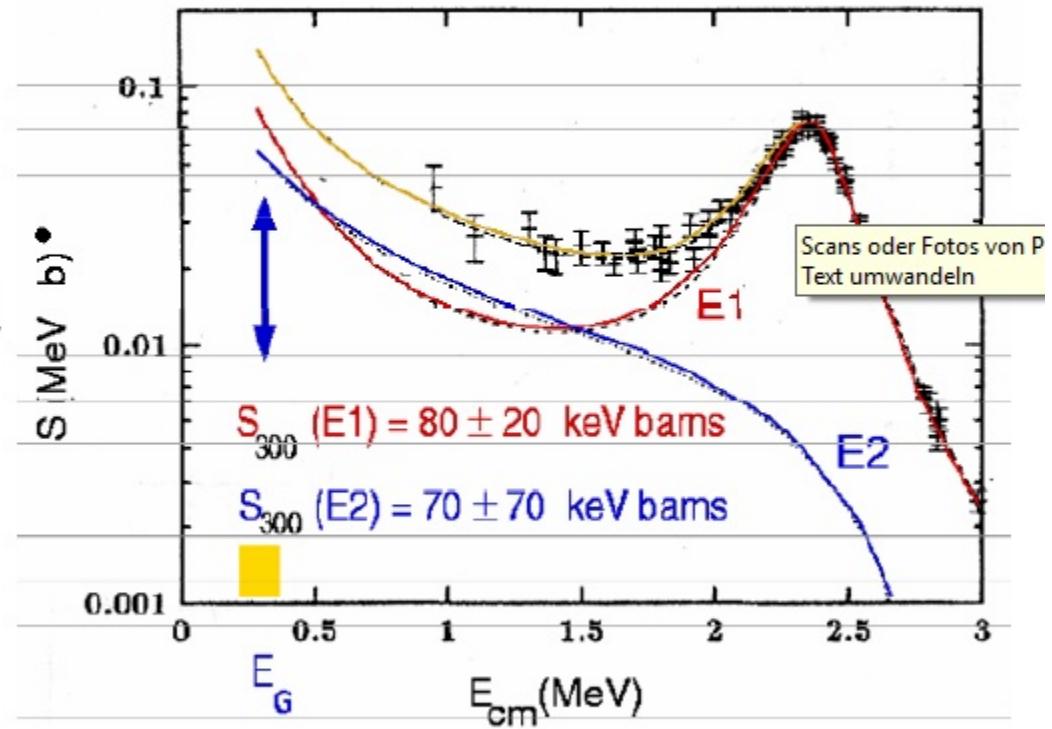
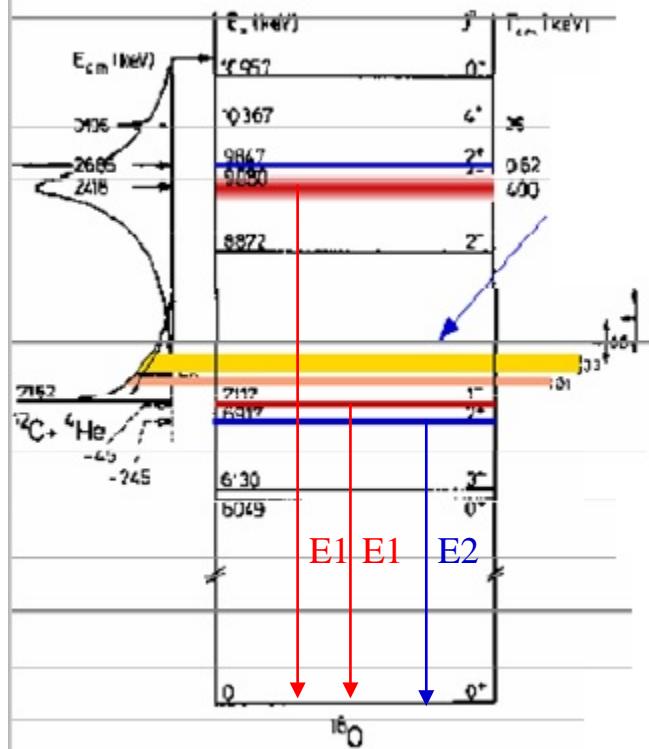
C-burning

- Typical conditions:
 - Temperature: (6-8) 10^8 K
 - Density: 10^5 g/cm³
- Net reaction: $^{12}\text{C} + ^{12}\text{C}$
- fuel: carbon
- main products: neon, magnesium, oxygen
- $^{12}\text{C} + ^{12}\text{C} \rightarrow \alpha + ^{20}\text{Ne}$ (Q = 4.62 MeV)
- $^{12}\text{C} + ^{12}\text{C} \rightarrow p + ^{23}\text{Na}$ (Q = 2.24 MeV)
- other reactions: $^{23}\text{Na} + p \rightarrow \alpha + ^{20}\text{Ne}$
 $^{20}\text{Ne} + \alpha \rightarrow ^{24}\text{Mg}$
- $^{12}\text{C} + \alpha \rightarrow ^{16}\text{O} + \gamma$
conversion of ^4He into ^{12}C and ^{16}O

$$kT \sim 8.6 \times 10^{-8} T[\text{K}] \text{ keV}$$

$^{12}\text{C}(\alpha, \gamma)^{16}\text{O}$, the Holy Grail

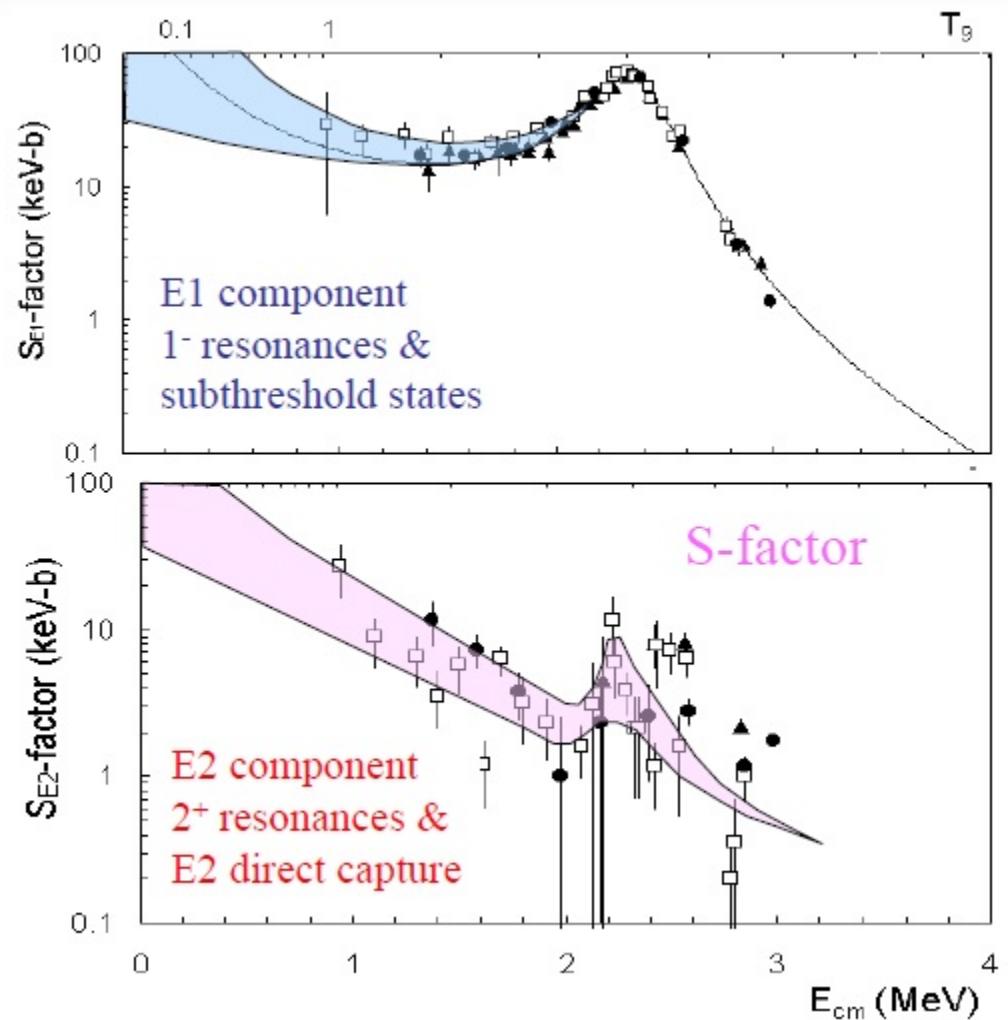
Level and Interference Structure between 1^- levels (E1) and 2^+ states (E2).



Uncertainty in low energy extrapolation

Reaction contributions in $^{12}\text{C}(\alpha,\gamma)^{16}\text{O}$

Difficulty in the reliability of low energy extrapolation



$^{12}\text{C}(\alpha,\gamma)^{16}\text{O}$ reaction rate

$$N_A \langle \sigma v \rangle = 6.9 \cdot 10^8 \cdot T_9^{-2/3} \cdot S_{eff} [MeV - b] \cdot e^{-\frac{32.11}{T_9^{1/3}}} \left[\frac{cm^3}{s} \right]$$

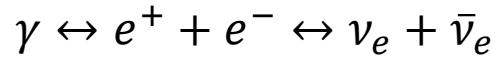
$$S_{eff} \approx 0.17 [MeV - b]$$

$$N_A \langle \sigma v \rangle \approx 1.2 \cdot 10^8 \cdot T_9^{-2/3} \cdot e^{-\frac{32.11}{T_9^{1/3}}} \left[\frac{cm^3}{s} \right]$$

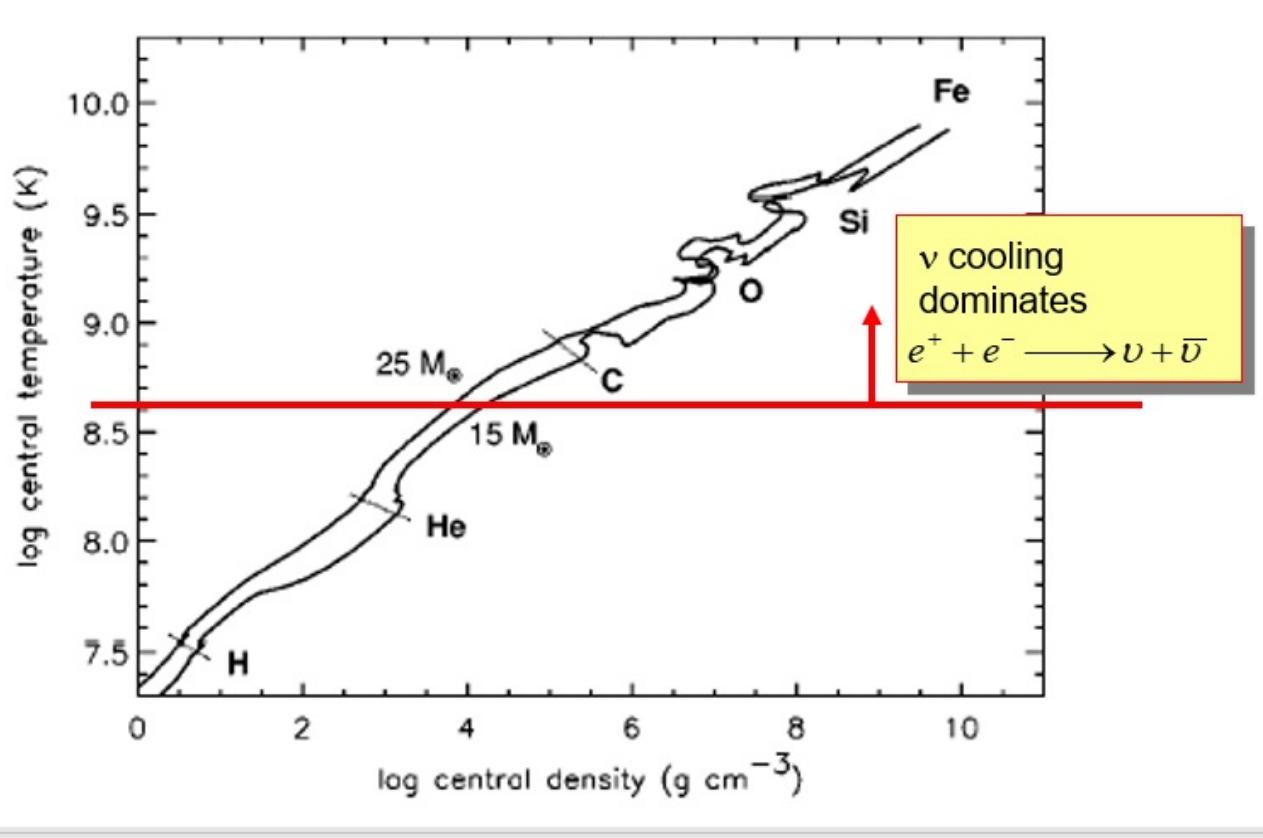
Only very crude estimate!
E-T dependency needs to be considered!

The role of neutrino-losses

- At temperatures above $\sim 10^9$ K: pair-production



- Luminosity of **photons** and **neutrinos**

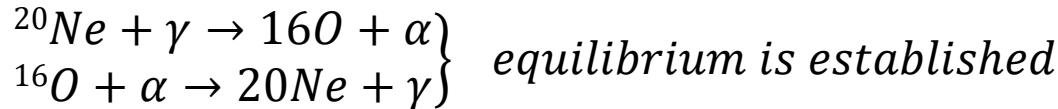


Ne-burning

- Typical conditions:
 - Temperature: (1-2) 10^9 K
 - Density: 10^6 g/cm³
- Reactions: $^{20}\text{Ne} + ^{20}\text{Ne} \rightarrow ^{16}\text{O} + ^{24}\text{Mg} + 4.59 \text{ MeV}$
- fuel: neon
- main products: oxygen, silicon
- $^{20}\text{Ne} (\gamma, \alpha) ^{16}\text{O}$
- other reactions:
 - $^{20}\text{Ne} (\alpha, \gamma) ^{24}\text{Mg} (\alpha, \gamma) ^{28}\text{Si} (\alpha, \gamma) ^{32}\text{S}$
 - $^{21}\text{Ne} (\alpha, n) ^{24}\text{Mg} (n, \gamma) ^{25}\text{Mg} (\alpha, n) ^{28}\text{Si}$
 - $^{23}\text{Na} (\alpha, p) ^{25}\text{Mg} (\alpha, n) ^{28}\text{Si}$
 - $^{25}\text{Mg} (p, \gamma) ^{25}\text{Al}$
 - $^{23}\text{Na} (p, \alpha) ^{20}\text{Ne}$

Why would neon burn before oxygen?

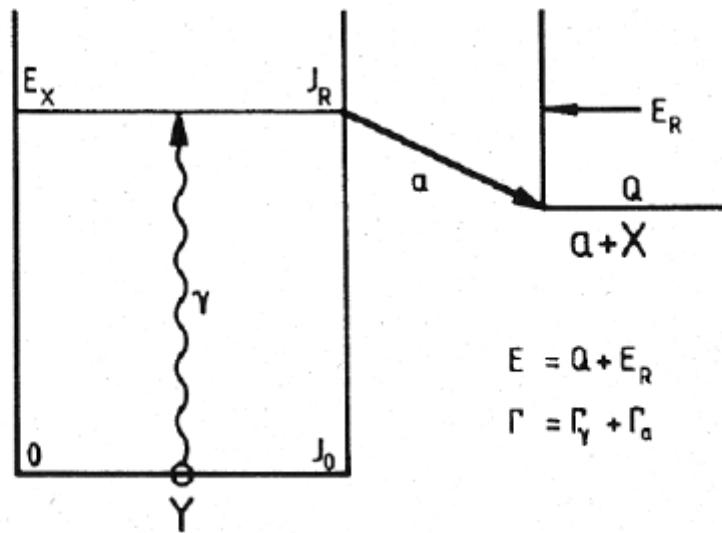
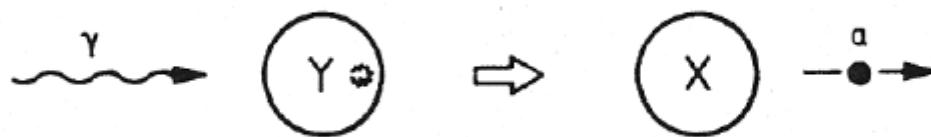
Temperatures are sufficiently high to initiate **photodisintegration** of ^{20}Ne



$$kT \sim 8.6 \times 10^{-8} T[\text{K}] \text{ keV}$$

Photodisintegration

PHOTODISINTEGRATION $\gamma(\gamma, \alpha)X$



O-burning

- Typical conditions:
 - Temperature: $(1.5\text{-}2.2) \times 10^9 \text{ K}$
 - Density: 10^7 g/cm^3
- Reactions:
 - fuel: oxygen
 - main products: silicon, sulfur (90%)
 - $^{16}\text{O} + ^{16}\text{O} \rightarrow ^{32}\text{S}^* \rightarrow p + ^{31}\text{P}$ (**56%**, $Q = 7.676 \text{ MeV}$)
 - $^{16}\text{O} + ^{16}\text{O} \rightarrow ^{32}\text{S}^* \rightarrow \alpha + ^{28}\text{Si}$ (**34%**, $Q = 9.593 \text{ MeV}$)
 - $^{16}\text{O} + ^{16}\text{O} \rightarrow ^{32}\text{S}^* \rightarrow n + ^{31}\text{S}$ (**5%**, $Q = 1.459 \text{ MeV}$)
- other reactions:
 - $^{31}\text{P} (p, \alpha) ^{28}\text{Si}$
 - $^{33}\text{S} (e^-, v) ^{33}\text{P}$
 - $^{35}\text{Cl} (e^-, v) ^{35}\text{P}$ $^{25}\text{Mg} (p, \gamma) ^{25}\text{Al}$
 - $^{23}\text{Na} (p, \alpha) ^{20}\text{Ne}$

$$kT \sim 8.6 \times 10^{-8} T[\text{K}] \text{ keV}$$

Si-burning

- Typical conditions:
 - Temperature: (3-4) 10^9 K
 - Density: 10^9 g/cm³
- Net reaction: $^{28}\text{Si} + ^{28}\text{Si}$
- fuel: silicon
- main products: Fe-group elements ($A = 50\text{-}60$ nuclei)
- other reactions: $^{28}\text{Si} + \gamma \rightarrow p + ^{27}\text{Al}$
 $^{28}\text{Si} + \gamma \rightarrow \alpha + ^{24}\text{Mg}$
 $^{28}\text{Si} + \gamma \rightarrow n + ^{27}\text{Si}$
- Balance between forward and reverse reactions
for increasing number of processes: $a + b \leftrightarrow c + d$
→ Nuclear Statistical Equilibrium (NSE)

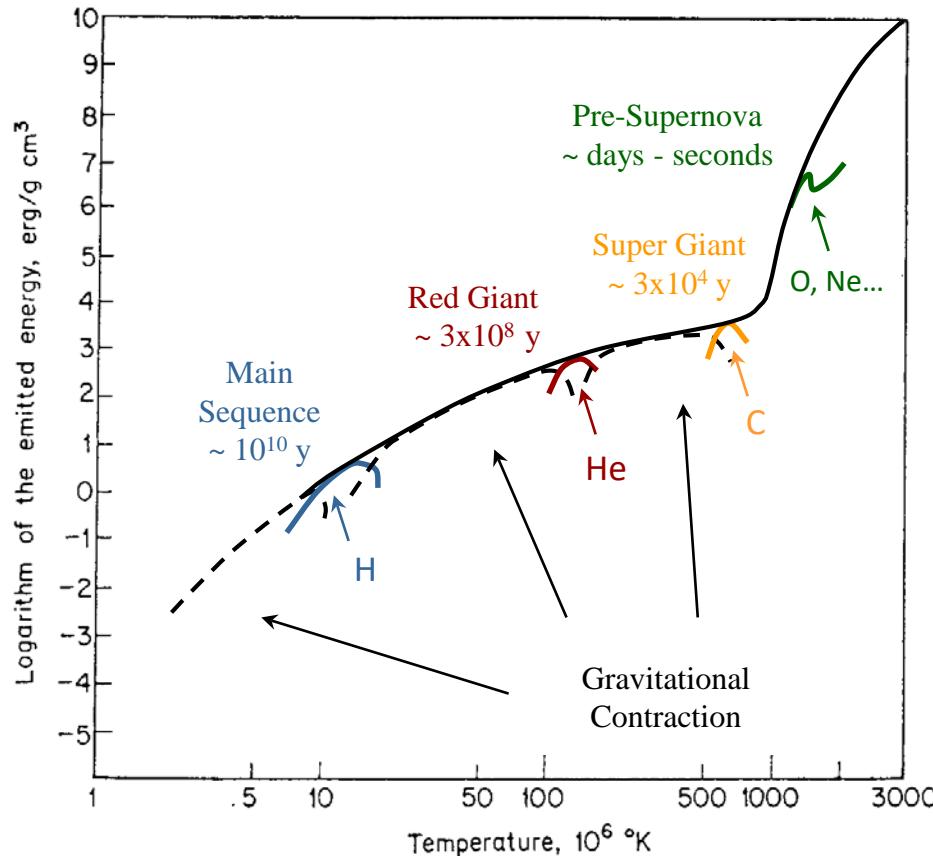
$$kT \sim 8.6 \times 10^{-8} T[\text{K}] \text{ keV}$$

Stellar evolution

Main parameters governing evolution: initial mass & initial chemical composition

Example: evolution stages of a $25 M_{\odot}$ star

Quiescent burning



Energy generation rate

$$\varepsilon \sim T^n$$

$n \sim 4$ (H-burning)

$n \sim 30$ (C-burning)



innermost regions only contribute to nuclear burning

e.g. $1/10 M_{\odot}$ for H-burning less for subsequent stages



H-burning \equiv MAIN SEQUENCE longest stage of star's lifetime

Summary stellar burning

TABLE 8.1 Evolutionary Stages of a $25 M_{\odot}$ Star^a

Stage	Time Scale	Temperature (T_9)	Density (g cm $^{-3}$)
Hydrogen burning	7×10^6 y	0.06	5
Helium burning	5×10^5 y	0.23	7×10^2
Carbon burning	600 y	0.93	2×10^5
Neon burning	1 y	1.7	4×10^6
Oxygen burning	6 months	2.3	1×10^7
Silicon burning	1 d	4.1	3×10^7
Core collapse	seconds	8.1	3×10^9
Core bounce	milliseconds	34.8	$\simeq 3 \times 10^{14}$
Explosive burning	0.1–10 s	1.2–7.0	Varies

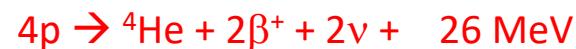
Why do timescales get smaller ?

Note: Kelvin-Helmholtz timescale for red supergiant ~10,000 years,
so for massive stars, no surface temperature - luminosity change for C-burning and beyond

... and nucleosynthesis

nucleosynthesis energy

HYDROGEN BURNING (1st equilibrium)



HELIUM BURNING (2nd equilibrium)



¹²C/¹⁶O BURNING

... ¹²C ashes = Ne, Na, Mg
... ¹⁶O ashes = Al, ... Si

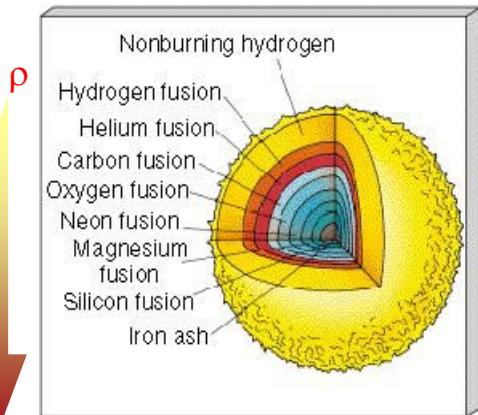
²⁸Si MELTING

major ash = ⁵⁶Fe
... A = 40-65

further reactions endothermic

gravitational collapse

The Stellar Onion



SUPERNOVA EXPLOSION

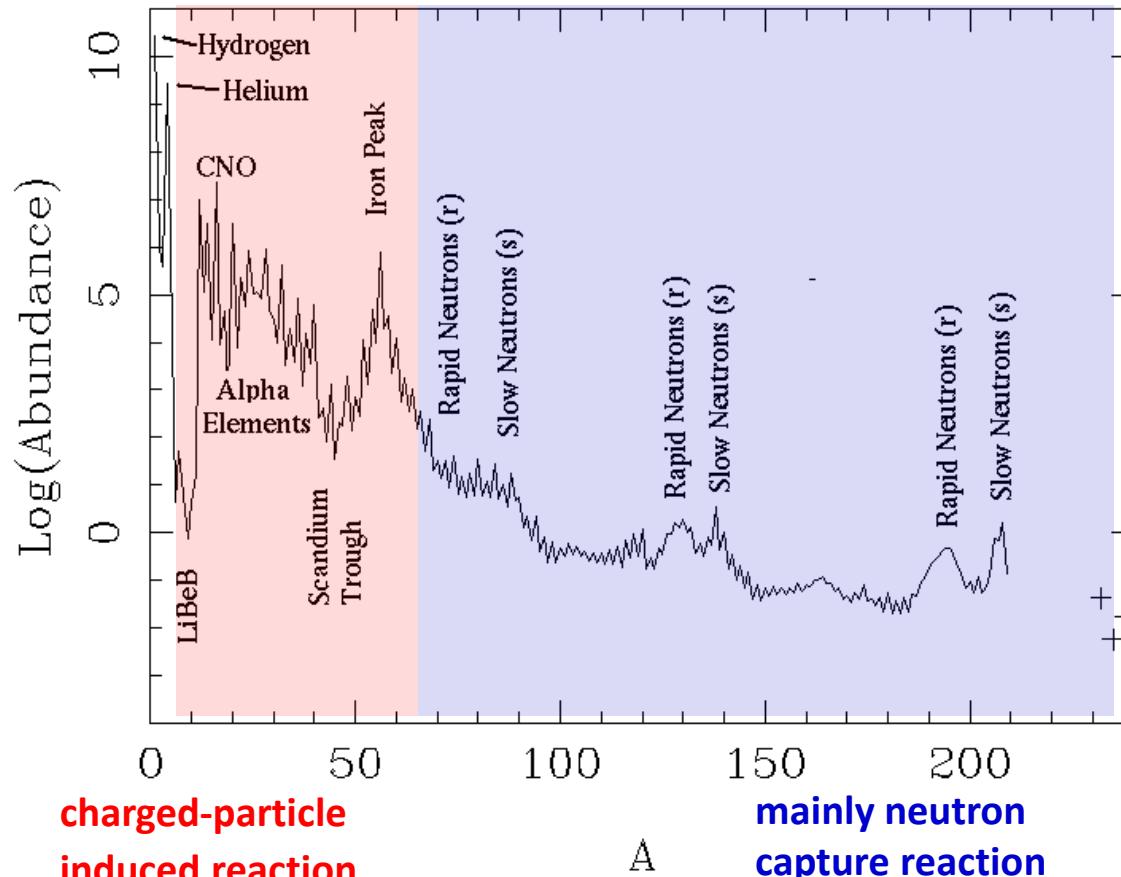


neutron star

black hole

Nuclear processes in stars

Standard Abundance Distribution (SAD) vs. A



**charged-particle
induced reaction**

during quiescent stages
of stellar evolution

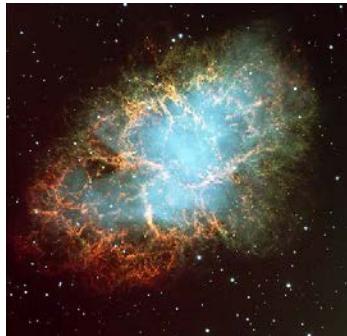
**mainly neutron
capture reaction**

mainly during explosive
stages of stellar evolution

involve mainly **STABLE NUCLEI**

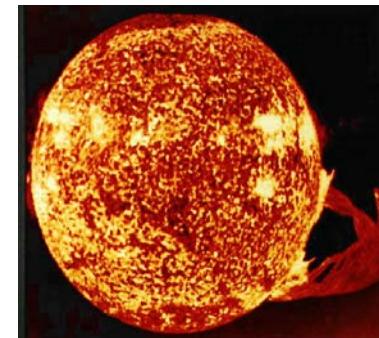
involve mainly **UNSTABLE NUCLEI**

Interstellar gas



BIRTH
gravitational
contraction

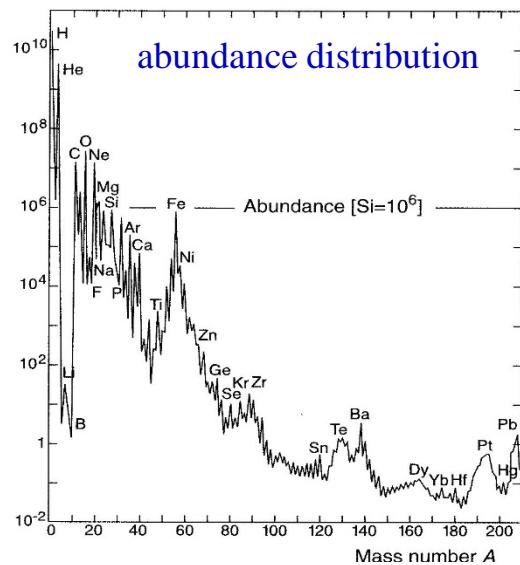
Stars



explosion
ejection
DEATH

mixing of
interstellar gas

thermonuclear
reactions



- energy production
- stability against collapse
- synthesis of “metals”